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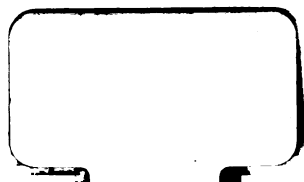
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AIR

AND

ITS RELATIONS TO LIFE

BEING WITH SOME ADDITIONS

THE SUBSTANCE OF A COURSE OF LECTURES DELIVERED
IN THE SUMMER OF 1874 AT THE ROYAL
INSTITUTION OF GREAT BRITAIN

BY

WALTER NOEL HARTLEY, F.C.S.

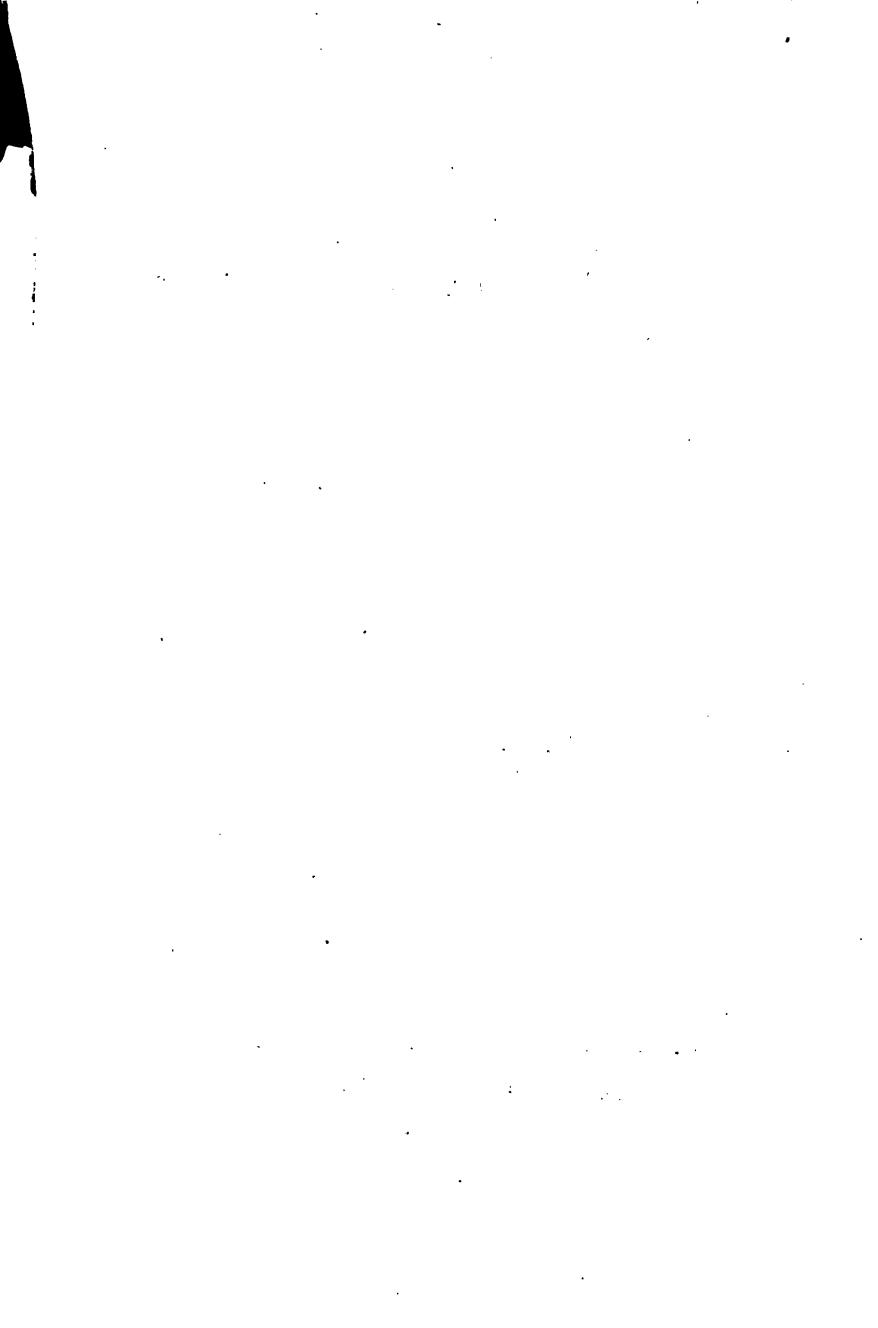
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1875



PREFACE.

THE following pages present an endeavour to give, in a light and popular manner, some information concerning that particular form of matter called Air, which is so essential to man that it comes to each individual with life, and leaves him not till death. Besides the narration of facts, an account of how these facts were obtained offers an insight into the particular mode of reasoning employed in scientific research, and endows the statements with that weight and interest necessary to leave a distinct impression upon the mind. As far as is consistent with clearness of expression, the use of scientific terms has been avoided. In one or two cases important practical details have been more fully treated than usual. Generally speaking, the original sources of information have been consulted, and to some

readers it may be of interest to know that these consisted of the papers of Andrews and Tait, of Graham, Ray Lankester, and others, in the 'Philosophical Transactions,' and in the 'Proceedings of the Royal Society'; likewise the contributions of Laplace, Dumas, and Boussingault, Regnault, Lewy, Schoenbein and Pasteur, in the 'Annales de Chimie et de Physique' and the 'Comptes Rendus' of the French Academy; of Bunsen, Brunner, and Pettenkofer, in the 'Annalen der Chemie,' &c. &c. I am further indebted to Dr. Roscoe's article 'Air,' in 'Watts' Dictionary'; Dr. Angus Smith's 'Life of Dalton'; also his very important work on 'Air and Rain'; 'Practical Hygiene,' by Dr. Parkes; 'Hétérogénie,' by M. Pouchet; De Bary's 'Morphologie und Physiologie der Pilze'; Cohn's 'Beiträge zur Biologie der Pflanzen'; 'Fungi,' by the Rev. M. J. Berkeley, in Hooker's 'English Flora'; and Dr. M. C. Cooke's recently published work; also Pettenkofer's 'Popular Lectures,' translated by Dr. Augustus Hess; besides many other writings, to enumerate which would furnish too long a list. Those who are acquainted with

Prof. Bloxam's work on 'Chemistry' will recognise some of the experimental illustrations and wood-cuts employed. I have to acknowledge that my constant intercourse and friendship with him have given me advantages of which I have availed myself. New illustrations, through the liberality of Messrs. Longmans, have been engraved by Mr. Collings from my own drawings, but for some of those in Chapter IV. my thanks are due to the Council of the Royal Society. The favour accorded to the short course of lectures which I had the honour of delivering at the Royal Institution of Great Britain, in the summer of 1874, and the subsequent enquiries as to their issue in a printed form, has been the chief inducement to their rearrangement and publication. Furthermore, having been struck with the very general want of knowledge in this country, even among scientific men, of the work accomplished by that celebrated French chemist, M. Pasteur, I thought that to make his labours more widely known would greatly promote scientific truth and accuracy.

For the sake of accuracy it may as well be

stated here, with reference to a sentence on p. I, that although the rotation of the earth is said in general language to be accomplished once in 24 hours, it is really completed in 23h. 56m. 4s., and that the time of its actual revolution round the sun, the duration of a sidereal year, is 365d. 6h. 9m. 9⁶s.

WEST DULWICH: *September* 1875.

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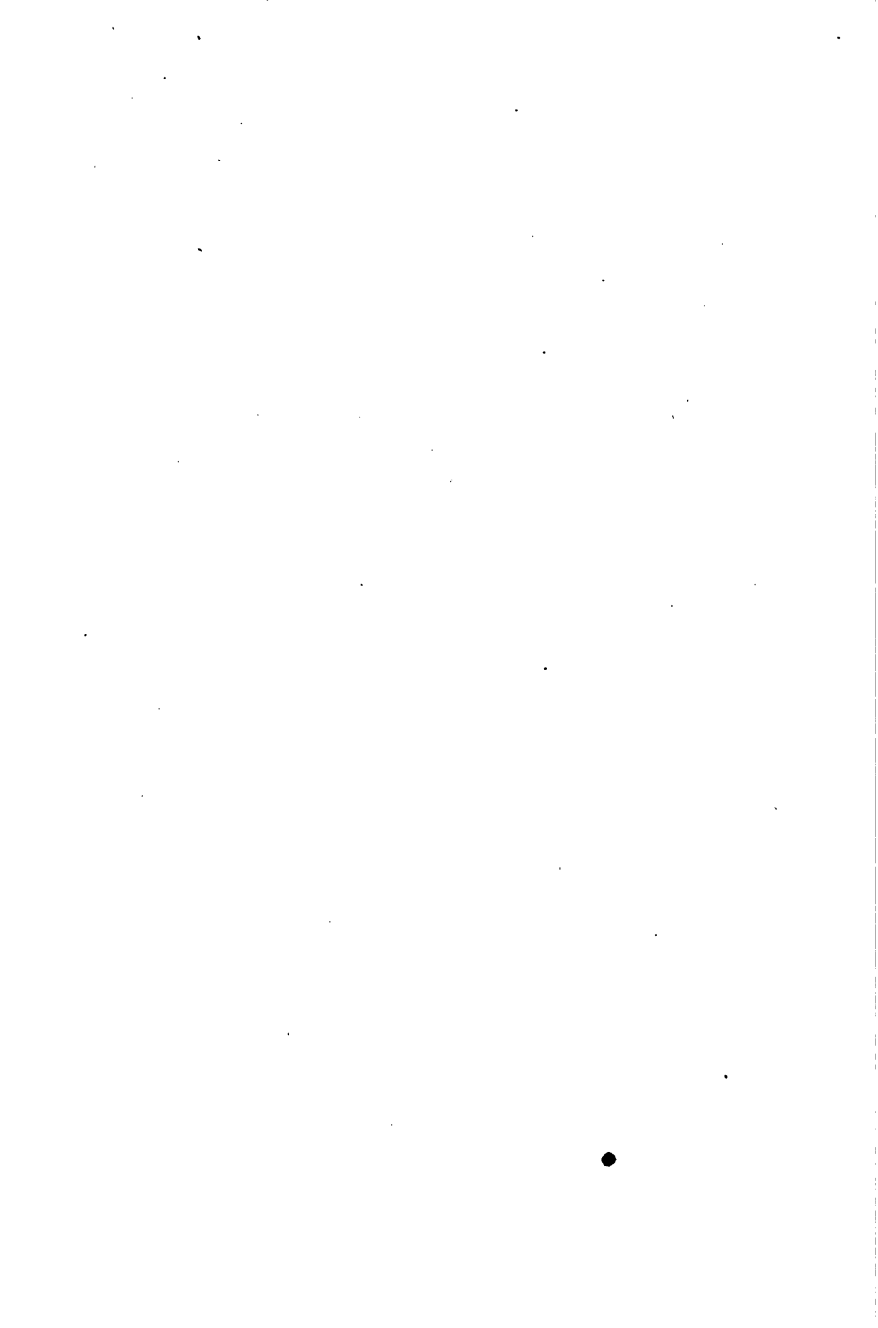
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A I R

AND

ITS RELATIONS TO LIFE.

CHAPTER I.

Proofs of a Material Medium surrounding the Globe—Galileo's Theory, and Torricelli's Proof of its Weight, in the Year 1640—Priestley's Discovery of Oxygen, and Lavoisier's Discovery of the Nature of Air, in 1774—The Composition of Air and the Properties of its chief Constituents demonstrated.—Dumas and Bous-singault's exact Experiments in 1841—Regnault, Bunsen, and Lewy's Determination of the Oxygen—Air from different Parts. of the World—Conclusions as to its Composition—Later Re-searches of Dr. Angus Smith—Reasons for considering Air a Mixture, and not a Compound, of Oxygen and Nitrogen—Tessié du Mothay's Process for extracting Oxygen from the Air.

ASTRONOMERS tell us that the Earth has the figure of an oblate spheroid, that it revolves on its axis once in 24 hours, and traverses its orbit in $365\frac{1}{4}$ days; also that its shortest diameter measures 7927·4 miles. There is no occasion to question this statement, yet it may be suggested that this latter measurement is true only so far as the visible

portion of the earth is concerned. A correction amounting to the addition of *at least* 90 miles to this number must be made on account of the surrounding invisible gaseous envelope or shell, which, by its action on the sun's light, has been computed to be not less than 45, and is probably 200, miles in thickness. It moves in the same orbit and on the same axis, and is in reality a most essential part of the globe, and more especially so to all living creatures. This portion is called the atmosphere. We are cognisant that we inhale this atmosphere, we are aware that it affords a certain resistance to the sails of a boat, that something preventing the entrance of water fills a glass or bottle plunged mouth-downwards below its surface, that this something has no definite shape, is without colour, taste, or odour, and is altogether unlike the other forms of matter around us.

We know that air is essential to life, yet it is more by being so taught than by intuition that we have this knowledge, for, as Dr. Angus Smith remarks, 'When we are children air is to us nothing ;' 'a vessel of air is a vessel with nothing in it.' Air enters our bodies by the lungs ; a portion is absorbed by the blood, and courses through the system ; we cease to exist without it, and as living creatures cannot rise above it ; it is at one and the same time a part of the earth, and a part of nearly

all living things, animal and vegetable, upon the earth. It constitutes by far the greatest portion of the atmosphere ; so much so, indeed, that the words air and atmosphere may be looked upon in most cases as having the same meaning. A history, therefore, of so important a matter as air cannot be otherwise than somewhat interesting to everyone.

That the atmosphere was a material substance was believed by Aristotle, but he did not succeed in establishing the truth of his convictions. The earliest experiment which failed to yield the desired proof may easily be repeated by anyone, and may be discussed as follows :—

All material substances possess weight and occupy space ; if, therefore, we take a bladder of air and weigh it, then extract the air and weigh it again, it might at first sight appear that the second weighing would be less than the first by the weight of the air abstracted ; but under ordinary conditions of *weighing in air* it is not so, because the experiment as thus performed, as far as the object in view is considered, gives us nothing more than the simple weight of the bladder itself. To make this more apparent, the experiment may be modified by substituting water for air. Fill a bladder with water, sink it in a vessel of water, and, by means of a string, attach it to a

pair of scales ; weigh it ; now, without removing the bladder from below the surface of the liquid, squeeze it until it is as nearly as possible empty, tie up the mouth, and weigh again ; the first and second weighing will be identical. The conditions of the two experiments are analogous. In the former case a bladder first full of air and then empty is weighed in air, in the latter a bladder of water, first full and then empty, is weighed in water, the conclusion in each case being the same. The reason why no satisfactory proof is thus obtained is accounted for by the following explanation. When a substance is weighed during immersion in any gaseous or liquid medium, the weight observed is smaller than the true weight by the weight of the surrounding medium displaced by the substance. A cubic inch of gold, for example, if weighed in air gives not its true weight, but a quantity, which, if added to the weight of a cubic inch of air, is equal to its true weight. Or the same thing may be stated thus : if a cubic inch of gold be weighed in water, it gives its true weight minus the weight of a cubic inch of water. A bladderful of air weighed in air therefore yields a quantity which is minus the weight of the air occupied by the distended bladder, and these two weights are practically identical.

The apparatus with which the true weight of

air was discovered consists of a glass globe of about a quart capacity, fitted with an accurate stopcock, the air being entirely withdrawn from the flask by means of an air-pump and the vessel suspended and accurately counterpoised on a balance. It is evident that if the tap be now opened the air will rush in, and this quantity of air being sufficiently heavy to turn a good balance, the beam descends on that side sustaining the flask.

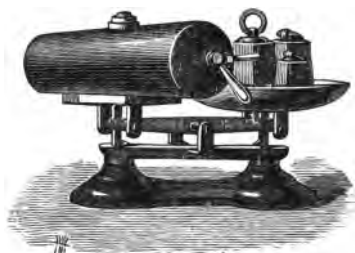


Fig. 1. Method of showing the Weight of Air.

One decisive experiment to show that air has weight may, as it is a novel one, be mentioned here. An iron bottle, into which $7\frac{1}{2}$ cubic feet of air have been pumped, is exactly balanced on a pair of scales; by allowing the air to escape, and, weighing again, it may be found to lose more than 7 ounces, and weights to this amount must be placed on the cylinder to restore equilibrium (see fig. 1). As the air flows into an empty gas bag

it is shown to occupy space. What further proof is necessary? Air is evidently a material substance. At a temperature of 60° F., the height of the barometer being 30 inches, 13 cubic feet of air would weigh one pound. An experiment done on so large a scale and with such rough weighing apparatus as the above may be considered to yield a very fair indication of this fact.

Galileo gave rise to a great discovery when he first explained to the pump-makers in Florence that the weight of a column of the atmosphere was not sufficiently heavy to balance a similar column of water more than 32 feet in height, hence the reason why they had failed to raise water in pumps to a greater height than 32 feet. Torricelli, his pupil, argued that the atmosphere which would support a column of water 32 feet would support a similar column of mercury only 30 inches in height, because mercury is 14 times heavier than water. This was verified by the following experiment, and so the first barometer was constructed:—A tube about 33 inches in length is filled to the top with mercury; being closed with the finger, it is inverted, the closed end dipped below mercury contained in a dish, and the finger removed; the mercury is then found to sink a little more or less than 3 inches. Now the weight of a column of mercury 30 inches in height and with a surface of

1 square inch would be 15 pounds ; the weight, therefore, of a column of the whole 45 miles of atmosphere an inch square would be 15 pounds, because it exactly balances 15 pounds of mercury. Hence the barometric pressure is said to be 15 pounds on the square inch.

The necessary consequences of the weight of the air which surrounds all terrestrial things are curious. For instance, the temperature at which water boils is 100° Centigrade, or 212° Fahrenheit, at the sea-level. If we decrease the pressure on the surface of the water we lower the boiling point, and this may be done by placing it under an air-pump and extracting some of the air surrounding it, when it can be very readily made to boil at the temperature of a warm summer day. Or we may lessen the pressure on the surface by ascending a high mountain, and so diminish the amount of atmosphere which rests its weight on the liquid. This Professor Tyndall has done, and he states that the boiling point of water on Mont Blanc is 184.95° F., or 27° below the boiling point at the sea-level. This difference is due to the removal of the pressure caused by a layer of air 16,000 feet in height. A variation of 1° F. indicates an ascent of 596 feet. It follows from this that if we increase the pressure the boiling point will be raised ; and it has been found by Regnault that at twice the

8 INCREASED AND DIMINISHED PRESSURE.

atmospheric pressure, or 30 pounds upon the square inch of surface, water boils at $249\cdot5^{\circ}$ F.—a rise in temperature of $37\cdot5^{\circ}$ F. Were it not for the weight of the atmosphere the ocean would evaporate, all living things, vegetables, and animals could no longer exist, and the earth would roll on its course a dull lifeless rock, clad here and there with ice, and surrounded by aqueous vapour.

The superincumbent mass of the atmosphere, which exerts a pressure expressed by the term 15 pounds on the square inch, but the extent of which maybe made more apparent to you by the statement that it amounts to 14 or 15 tons on the body of a man of average size, has its counterpart in the enormous pressure of water at great depths of the ocean. The dredging operations on board H.M.S. 'Challenger' have brought up fishes from great depths, distended and deformed, with their swimming bladders protruding from their mouths, because on coming to the surface the relief of the great pressure under which they were accustomed to live caused the air within them to be greatly expanded. Thermometers let down, which had been submitted to a pressure of 4 tons, were broken simply by the weight of the superincumbent water being greater than this. The iron cylinders filled with compressed gases, such as may be used in the foregoing experiment, usually sustain a pressure of 30

atmospheres, or $30 \times 15 = 450$ pounds on the square inch.

It is easily seen, then, that our condition on the earth is that of creatures living at the bottom of a vast ocean which floods the entire globe.

This gaseous ocean has its tides which ebb and flow from exactly the same causes as those of the aqueous ocean. Laplace was the first to draw attention to this in his '*Mécanique Céleste*.' These atmospheric tides are due to three causes—first, the direct action of the sun and moon upon the atmosphere; second, the periodical elevation and depression of the ocean; third, the attraction of the atmosphere by the ocean, the form of which changes periodically. These three causes are all derived from the attractive force of the sun and the moon, consequently the atmospheric and oceanic tides occur at the same periods. The atmospheric like the oceanic flow is due to the combined action of two partial tides, one caused by the sun, the other by the moon. We have evidence of the tidal flow of the atmosphere in the diurnal variations of the barometer. Extremely accurate observations, made every day for six consecutive years, showed that the mercury is always highest at nine o'clock in the morning and lowest at three in the afternoon; and at Paris the difference between the two readings amounts to eight-

tenths of a millimètre. Thus far we have had to deal with such of the physical properties of the atmosphere as are common to many other material substances, but in regarding it from a chemical point of view we must submit it to a closer examination, learn with what special properties it is endowed which distinguish it from all other kinds of matter, and likewise what are its relations to

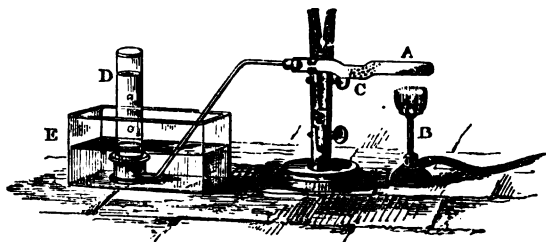


Fig. 2. Preparation of Oxygen from Red Precipitate or Oxide of Mercury.

- A Oxide of mercury in a glass tube.
- B Gas-burner.
- C Metallic mercury.
- D Cylinder of water inverted in the pneumatic trough E to collect the gas.

other substances, and more particularly to the various forms of living things.

Rutherford in 1772, in Edinburgh, discovered an inactive gas as a constituent of the air, but the discovery by Priestley in August 1774, a century since, that when he heated red precipitate (see fig. 2) he obtained metallic mercury and a gas or air

of remarkably active properties, which caused a glowing taper to be rekindled and combustible bodies to burn with increased brilliancy, paved the way to the discovery of the nature of air by Lavoisier still later in the same memorable year 1774. The experimental proof employed by the French chemist consisted in the calcination or burning of mercury at a moderate temperature

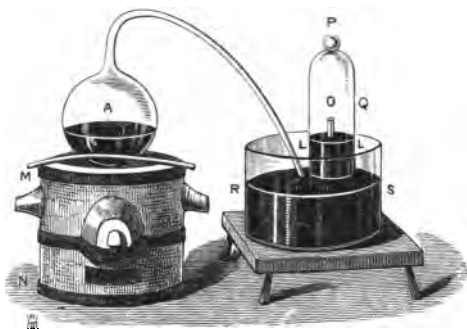


Fig. 3. Lavoisier's Experiment.

and its revivification by increased heat with the simultaneous production of oxygen gas. A translation of Lavoisier's own words affords the best description of his experiment :—

‘ I took a flask of about 36 cubic inches capacity, the neck of which was very long, with an internal diameter of half an inch. I bent it in the manner seen in fig. 3, so that it could be placed on the furnace M N, whilst the extremity O of its neck

fitted under the bell-glass P Q, placed in a bath of mercury R S. I put into this flask 4 ounces of very pure mercury ; then by sucking with a syphon introduced into the bell-glass P Q, I raised the mercury to L L. I carefully marked the height of this with a band of paper, and observed exactly the barometer and thermometer. These preparations made, I lighted the fire in the furnace M N, and sustained it for very nearly twelve days in such a way that the mercury was heated almost to its boiling point. Nothing remarkable occurred during the first day. The mercury, although not boiling, was in a continual state of evaporation ; it coated the inside of the vessel with globules, at first very minute, which increased until, having obtained a certain size, they fell back into the vessel and reunited with the mercury. On the second day I began to see little red particles swimming on the surface, which for four or five days continued to increase in number and size, after which they ceased to grow and remained in absolutely the same state. At the end of twelve days, seeing that the calcination of the mercury (oxidation of the mercury) made no further progress, I withdrew the fire and left the vessels to cool. The volume of air contained in the flask, its neck, and the bell-glass, reduced to a pressure of 28 inches and a temperature of 19° Centigrade, was before the

operation about fifty cubic inches.¹ When the operation was concluded the volume, under the same conditions of temperature and pressure, was found to be no more than forty-two or forty-three inches; there had consequently been a diminution in volume of about one-sixth. On the other hand, having collected carefully the red particles which were formed, and separated them as much as possible from the mercury with which they were mixed, their weight was found to be forty-five grains.

'The air which remained after this operation, and which had been reduced to five-sixths of its original volume by the calcination of mercury, was no longer fit for respiration or combustion; animals placed in it perished in a few moments, and lights were extinguished immediately, as if they had been plunged under water. On the other hand, I took the forty-five grains of red matter formed during the operation, and I introduced them into a little glass retort which was fitted to an apparatus meant to receive the liquid and gaseous products that might separate. Having lighted the fire in the furnace I observed that as the red substance was heated its colour increased in intensity. When the

¹ As the space occupied by a gas, that is to say, its bulk or volume, varies with the temperature and the barometric pressure, these details are necessary for exact measurement.

retort became nearly red-hot, the red substance began gradually to diminish in quantity, and after some minutes it had disappeared altogether; at the same time there were collected in the receiver $41\frac{1}{2}$ grains of mercury, and there passed into the bell-glass 7 to 8 cubic inches of an elastic fluid much more fit than atmospheric air to support combustion and the respiration of animals. Having passed a portion of this gas into a tube of glass an inch in diameter, a taper was plunged into it and diffused a dazzling brilliancy; the carbon, instead of burning quietly as in ordinary air, burnt with a flame and a kind of decrepitation like that of phosphorus, and an intensity of light that the eyes could scarcely support.

‘On reflecting on this experiment, it is seen that mercury during calcination (oxidation) absorbs the wholesome and respirable portion of the air; that the portion of the air that remains is a kind of choke-damp, incapable of supporting combustion and respiration. Atmospheric air, then, is composed of two elastic fluids of a different and, so to speak, opposite nature.

‘A proof of this important truth is, that in recombining the two elastic fluids which have thus been obtained separately, that is to say, the 42 inches of choke-damp or irrespirable air, and the 8 cubic inches of respirable air, air is re-formed in

every particular like that of the atmosphere, and which is equally fit for combustion, the calcination of metals, or the respiration of animals.'

Lavoisier adds that the proportion of respirable gas found by his experiment is probably a little too small, because he could not get the entire quantity of the constituent oxygen to combine with the mercury.

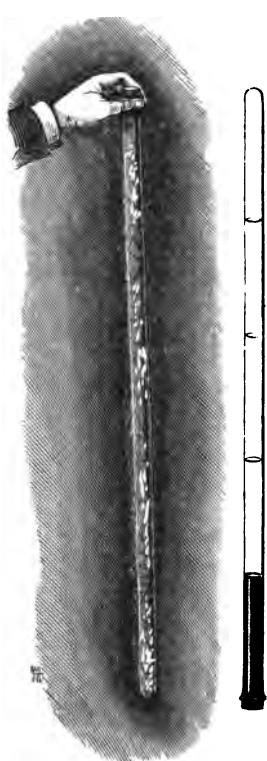
The conclusions of Lavoisier have been verified and corrected by later experiments more easily performed, but the great neatness of his demonstration, together with the light it throws on the action of air in the phenomena of combustion, casts a glory over the whole investigation.

That air consists for the most part of 4 volumes of nitrogen and 1 volume of oxygen we show nowadays in various manners, for instance, by suspending a stick of phosphorus upon a wire stand in a measured volume of air confined over water in a glass cylinder. The glass cylinder is divided into five equal parts by measuring water into it, and each space is marked on the outside by black lines.

After a few hours the phosphorus, having combined with all the oxygen to form phosphorous acid, a substance which is absorbed by the water, leaves the nitrogen occupying only four of the spaces marked on the jar.

We may illustrate the same fact in a much

more striking way, by the following new method (see fig. 4):—A good glass tube closed at one end, having an internal diameter of 1 inch and a length



of 45 inches, is fitted with an india-rubber bung. The tube should then be divided into five equal parts from end to end by small india-rubber rings stretched over it. A perfectly dry piece of phosphorus is placed within the tube about 3 inches from the india-rubber stopper; this is warmed until well melted, and then jerked down to the other end of the tube. In its progress it takes fire and so uses up all the oxygen. The tube soon becomes sufficiently cool for the end to be dipped beneath coloured water, when on removal of the cork the liquid rushes in to supply the place of the missing

Fig. 4. Experiment to show the Composition of the Air.

oxygen. On raising or depressing the tube so as to bring the liquid within and without to the same

level, inserting the cork again beneath the water and removing, it is seen that the liquid within occupies one-fifth the entire capacity of the tube, which if the tube be of equal bore throughout will be as nearly as possible 9 inches. See fig. 4.

It having thus been demonstrated that there are two bodies, nitrogen and oxygen, present in the air, let us examine their properties. First the nitrogen. Taking a bottle full of this substance it is seen to be a colourless gas, for the bottle has the appearance which we generally term empty. Removing the stopper there is no smell detected, and placing a taper within it is found not to be an empty bottle, for the taper is extinguished; another is lighted, and this is extinguished; the gas is therefore neither combustible nor a supporter of combustion. On placing some water in the bottle and shaking it up, the mouth being closed by the hand, we do not find the water altered at all; it is therefore not soluble to any great extent. All positive properties are 'conspicuous by their absence.' It is colourless, odourless, tasteless, incombustible, irrespirable, insoluble, and a non-supporter of combustion. When, however, it is combined, on the one hand, with hydrogen to form ammonia, or with oxygen and hydrogen to form nitric acid; we have in either case a new substance produced which is remarkable for its most strongly pronounced

properties, and of the greatest chemical activity ; but the strangest fact is, that these two substances, ammonia and nitric acid, have exactly opposite properties, but on mixing they combine together and neutralise the distinctive characters of each other.

With regard to the other constituent, oxygen, all the positive character of air is due to it. In



Fig. 5. Charcoal burning in Oxygen.



Fig. 6. Phosphorus burning in Oxygen.

other words, the properties of air are the properties of oxygen much enfeebled. We therefore find that all substances which burn in air will burn with greater brilliancy in oxygen, while some, such as iron and zinc, which are incombustible in air are readily consumed in oxygen.

Experiments arranged to show the burning in

oxygen of charcoal, phosphorus, and also a steel spring, are shown in the accompanying illustrations (see figs. 5, 6 and 7). In the first case it is difficult to get the charcoal to take fire at all in air, but having at last kindled it so as to get a minute point to glow, on placing it in the oxygen it bursts into a brilliant glittering shower of sparks. With phosphorus the scarcely luminous and, ghost-like



Fig. 7. A Steel Spring burning in Oxygen.

flame becomes a surprisingly brilliant blaze, and even a steel watch-spring is consumed like a piece of charcoal. The opposite characters and the different properties of the two gases are thus forcibly demonstrated.

The question would arise in the mind of every chemist after repeating the experiments of Lavoisier—What is the constitution of the atmosphere: are

these two gaseous constituents simply mixed together, or are they chemically combined? Are the oxygen and nitrogen combined in the proportion of 1 volume of the former to 4 of the latter, so that they produce a third distinct substance called air, just as sulphur and quicksilver can be combined together to form a third substance, vermilion; or are the two gases merely mixed? While many chemists, amongst whom we find the names of Dœbereiner,¹ Prout, and Thomson, considered, not without good reason, the atmosphere to be a compound of 20 volumes of oxygen and 80 volumes of nitrogen (that is 1 to 4), Dalton, on the other hand, supposed it to be a mixture differing in composition at different altitudes according to the relative weights of the two gases, but this conviction was founded less on experiment than calculation. At first it appeared probable that the unequal development of animal life on various parts of the globe would effect considerable alteration in the air of different localities, and this idea was supported not only by the unhealthiness of the air of certain districts, such as crowded towns and marshy lands, but also by the earliest quantitative analyses, which showed differences in the composition of air of different localities amounting to 10 per cent.

¹ The professor under whom Goethe studied chemistry at Goettingen.

Babinet, indeed, went so far as to calculate from the air of Paris the probable composition of the air at various heights up to an elevation of 10,000 mètres, or more than 32,500 feet. Dalton, Gay-Lussac, Davy, and Boussingault, however, by making use of improved methods of analysis, found the oxygen to vary but slightly. Gay-Lussac and Thénard, in a memorable balloon ascent at an altitude of 7,000 mètres, obtained a sample of air which gave the same numbers on analysis, namely 21.6 per cent., as a sample taken at the surface of the earth in Paris at the same time. This proved that altitude does not affect the relative proportion of the two constituent gases in air. This was afterwards confirmed by Brunner, who analysed air collected on the summit of the Faulhorn.

But it was necessary to have more accurate methods of analysis than those hitherto employed.

The experiments of Dumas and Boussingault in November 1841 finally established with great exactitude the true composition of air. Their apparatus (see fig. 8) consisted of a tube C C, provided with stopcocks, and containing metallic copper. The whole being weighed previously to being heated to redness in a furnace was connected at the one end with a series of tubes A, containing potash to absorb carbonic acid, with another set B, containing

sulphuric acid to absorb ammonia and water, and at the other end a large glass globe N, provided with a stopcock exhausted of air and accurately weighed. On slowly opening the taps the pressure of the external air caused it to flow through the series of tubes, thus becoming perfectly purified, and gain admission to the heated metallic copper ; here the oxygen was retained by the copper, forming a black solid metallic oxide, while pure nitrogen filled the globe. All the stopcocks were then

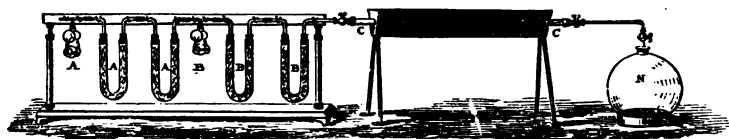


Fig. 8. Apparatus for the exact Analysis of Air.

closed, and the globe was detached and weighed : the difference between the previous and this second weighing expressed the weight of the nitrogen which had entered the globe. The tube was then separated and weighed a second time, then exhausted at the air-pump and weighed a third time : the difference between the second and third weighings gave the weight of the remnant of nitrogen in the tube, that portion not capable of passing into the globe. The difference between the first and second weighings of the tube expressed the weight of oxygen held in combination with the copper.

For example, the figures representing the result of the analysis are here given :—

	Weight of	In Grains
Globe (N) with nitrogen (at the conclusion).	.	3076
Exhausted globe (at the commencement)	.	3000
Nitrogen in the globe	.	76
Tube (C) with residual nitrogen (at the conclusion)	.	2574
Exhausted tube (at the conclusion)	.	2573
Nitrogen remaining in the tube	.	1
Nitrogen in the globe	.	76
Nitrogen remaining in the tube	.	1
Total nitrogen in the air analysed	.	77
Exhausted tube (C) with oxidised copper (at the conclusion)	.	2573
„ „ metallic copper (at the commencement)	.	2550
Oxygen in the air analysed	.	23

The air therefore contained 23 parts of oxygen by weight and 77 parts of nitrogen.

The results obtained led to the conclusion that air, purified from water, carbonic acid, and ammonia, had the composition by weight of 77 parts of nitrogen and 23 parts of oxygen.

Analyses on three successive days, viz. :

Proportion of Oxygen in 100 parts of Air.

	By Weight		By Volume
1841	1st Experiment	2nd Experiment	
April 27	22.93	22.93	20.729
„ 28	23.03	23.09	20.828
„ 29	23.03	23.04	20.828

It will be noticed in these numbers of Dumas and Boussingault that there are slight differences on different days in the proportion of oxygen present in 100 parts of air. With such a cumbrous apparatus and so elaborate a method of experimenting, it was not possible to make any very great

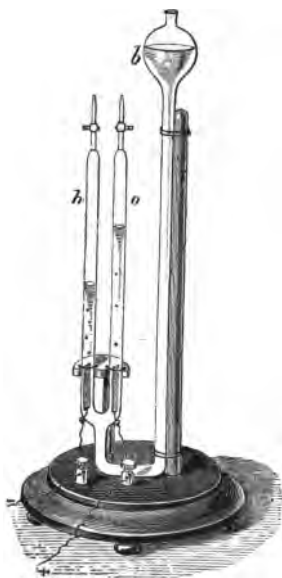


Fig. 9. Apparatus for the Decomposition of Water.

number of experiments, but by the later methods of gas analysis pursued by Bunsen, Regnault, and Lewy thousands of experiments have been made. This method depends on the fact that oxygen and hydrogen gases combine in the proportion of one volume of the former to two of the latter gas to form an inappreciably small volume of liquid water. By decomposing water with a galvanic battery, and collecting the two gases separately in a vessel made for the purpose

(fig. 9), it is easily seen that the hydrogen gas in the tube *h* is double the volume of the oxygen in *o*. That the gases when mixed and

then combined entirely disappear may be made obvious to the eye by exploding the mixture over mercury in the following manner:—The mixed gases are obtained from acidified water, which is decomposed in a voltameter (A, fig. 10) by means of a galvanic battery. The platinum plates B (the essential parts of the voltameter) are attached to the opposite poles of the battery by the wires C and D, and the gas is collected in the stout glass tube E, which has been previously filled with water.

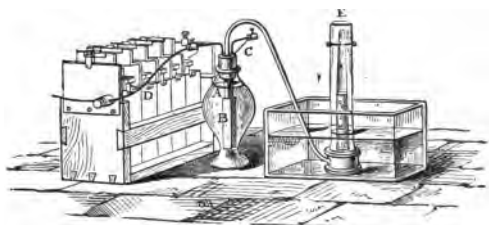


Fig. 10. Gases resulting from the Decomposition of Water being collected in a Glass Tube.

To conduct the electric spark to the interior of the tube two stout platinum wires are cemented into holes drilled through the glass near the closed end. When the tube is filled with the mixed gas it is closed with the thumb and removed to a thick porcelain mortar containing mercury, where it is firmly pressed down on a cushion of stout vulcanised india-rubber; the spark is now passed through the gas from one wire to the other by

means of an induction coil, an explosion takes place with a strong concussion but no noise, and since a vacuum is produced the cushion is found to be tightly pressed against the open end of the tube. On gently releasing this the mercury rushes up with violent rapidity, completely filling the whole space (see fig. 11); showing that neither oxygen nor hydrogen remains in excess. The two gases combine to form a trifling, in fact invisible, portion of water. Such is the main principle upon which the estimation of oxygen in atmospheric air de-



Fig. 11.
After the Explosion.

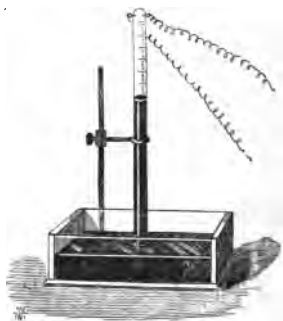


Fig. 12.

pends. In practice long tubes (fig. 12) are taken very accurately marked into an immense number (about 700) of small and equal divisions by an engraving instrument attached to a dividing engine. A specimen of air is introduced (2 or 3 cubic inches) after having attained the temperature of

the room ; the level of the mercury in the tube is carefully measured by means of a telescope fixed perfectly horizontal on a vertical stand ; at the same time the temperature and height of the barometer are noted. Hydrogen is then let up into the tube to mix with the air in the proportion of about half its volume ; the bulk of the mixture is again measured, the barometer and thermometer again observed. Next a spark is passed through the gas, whereby a certain amount of contraction takes place. Another set of readings as before are made, and the loss in volume of the original quantity of air is the volume of oxygen which has disappeared in company with twice as much hydrogen to form water. The barometric pressure and the temperature are necessary observations, because changes indicated by these instruments affect the volumes of gases to a great extent. The error made by an experimenter skilled in gas analysis never exceeds $\frac{1}{100000}$, and most frequently is not more than $\frac{1}{1000000}$.

Lewy analysed many hundreds of samples of air collected in Paris, at Havre, on the Atlantic, in New Granada, and on many mountain-tops. Regnault in 1852 published his 'Researches on the Composition of the Atmosphere' in the 'Annales de Chimie.' His plan was to send out instructions to his friends and acquaintances, and others

willing to help him in various parts of the world, to collect samples of air, to be sent to him for analysis. These were taken always at the same hour, towards noon, and on the first and fifteenth of each month during a whole year. At the same time the air of Paris was collected and submitted to analysis in the same way and with the same apparatus. The way in which the specimens of air were taken and preserved was the following:—Glass tubes were used, ending



Fig. 13. Tube for a Sample of Air.

in two open points, and these drawn-out points are very fragile. To avoid being broken in carriage the ends are covered by two little capsules, which are fixed on with wax or gum mastic, as in fig. 13. Each tube thus protected is placed in a cardboard case. To take a sample of air,



Fig. 14. Transference of Air to the Tube.

soften the wax or mastic, detach the capsules, and connect one of the points with a pair of bellows by means of a piece of india-rubber tube; then blow gently for two or three minutes (fig. 14).

The air of the tube is thus replaced by that of the bellows, which takes its supply from the surrounding atmosphere. It is necessary now to seal the tube hermetically. For that purpose withdraw the point from the india-rubber and heat it in the upper part of a spirit-lamp flame. When the

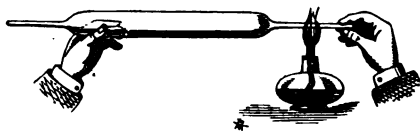


Fig. 15. Hermetically sealing a Tube.

glass is softened the point is drawn gently out so as to detach it from the tube, but without removal from the flame, and thus the tube is sealed on that side; the same operation is then performed at the other end. Specimens were taken in Berlin, Madrid, Geneva, Chamounix, on the Mediterranean, in the South Seas, and during the Arctic voyage of Captain James Ross¹ in search of Sir John Franklin.

The conclusions arrived at were :

1st. That the atmosphere shows sensible, although very small, variations in its composition, for the quantity of oxygen does not generally vary more than 20·9 to 21 volumes in 100 of air; but that in some cases, observed more frequently in

¹ Since Sir James Ross.

warm countries, the proportion of oxygen is as low as 20·3 volumes per 100 of air.

2nd. That the mean quantity of oxygen in 100 volumes of atmospheric air in Paris during the year 1848 was 20·96.

Taking into consideration the observations of Bunsen, Lewy, and Regnault together, we may consider the following fact established:—*Air from whatever altitude, and from whatever spot on the earth's surface, is very nearly of uniform composition.* This the following table from analyses of a later observer, Dr. Angus Smith, makes evident:

		Volumes of Oxygen in 100 of Air
Sea shore and open heath.	Scotland	20·999
Tops of hills.	Scotland	20·980
In a suburb of Manchester.	Wet weather.	20·980
The same	20·960
St. John's, Antigua	20·950

We see then the constancy of composition of the atmosphere notwithstanding that its constituent gases are not of the same weight, oxygen being a little heavier than nitrogen in the proportion of 16 to 14. This difference in weight may be well shown by the following arrangement: A jar of oxygen O, closed by a glass plate, is placed upon the table; a jar of nitrogen N, open top and bottom, both openings being closed by glass plates, is placed over it (see fig. 16), so that the two gases

may come together when the glass plates in contact are removed. The nitrogen will float for a few seconds above the oxygen, and then if a lighted taper be quickly introduced through the neck of the upper jar it is extinguished in passing through the nitrogen, rekindling brilliantly on reaching the oxygen in the lower jar, so that there can be no doubt as to the difference in weight of the two gases. It is possible to account for the constancy of composition by assum-



Fig. 16.

ing, as Döbereiner and Prout did, that the oxygen and nitrogen in the air are chemically combined, but then we must, before making such an assumption, assign good reasons for it. For the purpose of considering this matter let us examine the difference between substances which are chemically combined and mechanically mixed.

To this end take the red gas nitric peroxide, and see how different it is in all its properties to a mixture of its constituents. A large bottleful of this gas is seen to be of a reddish-brown colour, but in small quantities its tint may be described as

orange-red. It is heavier than air; so much so, that it can be poured out of the vessel containing it into another. When a lighted wax taper is plunged into it, it burns with a dull red flame, while a wooden spill is extinguished altogether; yet there is in a bottleful of this gas a much larger proportion of oxygen and a smaller proportion of nitrogen than exists in an equal measure of atmospheric air; for while this red gas contains 2 volumes of nitrogen and 4 volumes of oxygen, air, as already explained, contains 4 volumes of nitrogen, and only 1 volume of oxygen, and yet in atmospheric air the presence of oxygen is more evident from the behaviour of the substance towards combustible matter than it is here.

Furthermore, weight for weight the oxygen present in the brown gas is much greater, for whereas, 100 parts by weight of air contain 23 parts of oxygen, 100 parts of nitric peroxide contain more than $69\frac{1}{2}$ parts of the same gas.

Take as another comparison a mixture of 4 volumes of oxygen with 2 volumes of nitrogen; it is colourless, and on introducing a lighted taper the combustion is not feeble, but most brilliant, and so likewise with a half-extinguished wooden spill. In fact, the properties of this kind of air are such as we should assign to a gas containing so large a proportion of oxygen, supposing that we

had never seen it; things burn in it rather more than three times as rapidly and brilliantly, because it contains rather more than three times as much oxygen as air. But the reason the same conditions are absent when we take the red nitric peroxide gas is, that *when two substances chemically combine, the distinctive properties of each completely disappear, and new properties are developed.* This is why chemical research constantly discloses new substances with new properties, and why in nature we have an almost infinite variety of different kinds of matter, that is to say, substances with different properties. Another illustration of the difference between combined and mixed materials is thus again afforded: two bottles standing side by side, each containing a colourless gaseous substance, and each containing equal proportions of oxygen and nitrogen. The mixture behaves in no way different to one's expectations when the stopper is removed, and on plunging in a lighted taper it burns with brilliancy and indicates the presence of oxygen in a large proportion. Now another gas, nitric oxide, behaves in a very remarkable manner towards atmospheric air, so that when the stopper is removed from a bottleful it actually becomes red-brown in colour—in fact, is converted into nitric peroxide—and on plunging in a lighted taper it is extinguished. Yet another illustration:

one bottle containing a greenish-grey powder, and another a brilliant colourless, mobile, beautiful liquid, are each filled with the same materials. In each bottle there are 64 grammes of sulphur and 12 grammes of carbon; in the one only intimately mixed, in the other they are chemically combined to form the liquid carbon disulphide.

It is now evident that when two substances are mixed the mixture partakes of the properties of each of the constituents, and the properties of the mixture vary as the proportions of the constituents vary. But when two substances are combined, the properties peculiar to each constituent disappear, and a set of entirely new properties are developed. Again, substances combine only in fixed and definite proportions, so that given any chemical compound, no matter how it be made, the proportions by weight of its constituents are invariable and its properties are invariable in consequence. In this lies the distinction between a compound and a mere mixture, in which the constituents may be present in any proportion whatever. It has already been demonstrated how oxygen and hydrogen gases contract in volume on combination to form water: such contraction is a frequent occurrence during chemical combination. In fact, chemical combination is made evident to us by the following phenomena:

1st. By disengagement of heat, as in the burning of phosphorus.

2nd. By the contraction of the constituents to a smaller volume.

3rd. By the resulting compound possessing properties which differ strikingly from those of its constituents.

4th. By the constituents of a compound occurring in fixed and invariable proportions. These are identical with, or bear a simple relation to, certain proportional weights which are found to regulate the combination of simple substances.

That air has none of the properties of a compound may be easily shown. Four pint bottles of nitrogen, and one pint of oxygen are simply mixed in a tall jar standing over water. No contraction, in volume or evolution of heat takes place, and the product acts in every way as air. In short, the following are the arguments that air is not a compound but only a mixture:

1st. By the physical and chemical properties being the exact mean of those of its constituents; so that in the easiest manner possible, knowing as we do the properties of the substances of which it is composed, we can by calculation, readily foretell the properties of air.

2nd. A much feebler chemical action is suffi-

cient to separate its constituents than is effective with most chemically combined substances.

3rd. Solution in water, which is not a chemical process, is capable of separating to a large extent the oxygen of the air from the nitrogen, so that air which has been dissolved in water is found to be richer in oxygen than undissolved air. Humboldt and Gay-Lussac found in the air

		In 100 parts of Air	
		parts of oxygen.	
Of distilled water . . .	32		
Of river water (Seine) . .	31.9	„	„
Of rain water . . .	31.0	„	„

But the experiments of Boussingault are very decisive and of great interest. It was observed by Bischoff that the air entangled in snow contained 10 to 11 per cent. only of oxygen. Objecting to these numbers, on account of the way in which the analysis had been conducted, Boussingault repeated them, and found that when snow was melted the air remaining undissolved by the snow-water had this composition, but the air contained in solution in the water had 32 per cent. of oxygen. But the neatness of the work was completed when he showed *that the composition of the air dissolved in the water and that of the air undissolved gave figures which represented the composition of the atmospheric air as determined on the same day.*

4th. By a process of filtration through thin india-rubber, it is possible to partially separate the oxygen from the nitrogen in atmospheric air.¹ (See fig. 17.)

5th. Air may be easily made by mixing oxygen and nitrogen in the suitable proportions, when it is found that no evidence of chemical or physical change takes place, such as accompanies the formation of a compound substance.

A problem well worth solving is the separation of oxygen from atmospheric air, so as to make it available for many purposes, such as, for instance, the production of very high temperatures. This has at last been successfully accomplished by M. Tessié du Motay. The



Fig. 17.

¹ The process may be explained by the aid of the diagram fig. 17. B is a flat air-tight bag of waterproof silk, containing a piece of carpet to prevent the two sides coming in close contact. It is attached to the Sprengel pump A A A, which, by reason of a fall of mercury from the funnel down the vertical tube, drives the air out and delivers it at c. Each drop of mercury acts as a piston. After a good vacuum is attained, the gas delivered at c has more

process may be demonstrated in the following manner (see fig. 18): a copper tube *t* is filled with a mixture of manganate of soda and oxide of copper, and placed in a furnace consisting of a row of gas burners *b*. One end of the tube is connected with a T-shaped tube, so that by one limb *a*, a current of air, or by the other a jet of

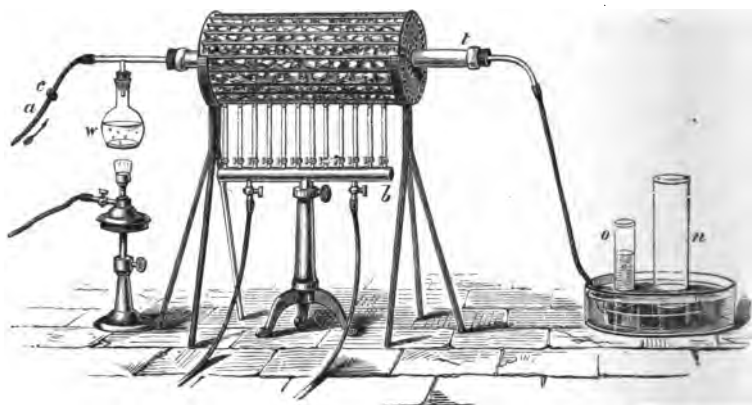


Fig. 18. Apparatus for Extracting Oxygen from Air.

steam, from the flask *w*, may be passed into the copper tube.

On heating the manganate of soda to dull redness, and passing a current of steam through the tube, oxygen is liberated and collected at *o*, while

the properties of oxygen than atmospheric air; in fact, the percentage of this gas rises from 23 to 47. This being the earliest and simplest form of Sprengel's pump, shows its mode of action best.

within the apparatus caustic soda and oxide of manganese are produced. When the current of steam is shut off and air is turned on, the oxygen of the air is abstracted by the contents of the tube, while the nitrogen passes out at the extremity, and is collected in the cylinder *n*.

The formation and the instability of manganate of soda may be readily shown. In a silver dish supported over a gas burner a little caustic soda is fused. Add to it a little black oxide of manganese, and the formation of a beautiful rich green colour gives evidence of the production of manganate of soda. This green substance now dissolved in water, if added to a jar of alkaline liquid, remains green, while one or two drops in a gallon of dilute acid give rise to the splendid red colour characteristic of Condyl's disinfectant, the permanganate of potash. Why the air is of such constant composition although its constituents are not of the same specific gravity, and are not combined, will come to light in a future chapter.

CHAPTER II.

Black's Experiments on the Carbonating of Lime in 1754—The Properties of Carbonic Anhydride, more commonly called Carbonic Acid—Its Production by Combustion, Respiration, Fermentation, and Decay—Of Aqueous Vapour—Its Condensation—Its Production, partly by Combustion and Respiration—Deliquescence illustrated by Iodide of Cobalt—Ammonia—Its Proportion in the Air—Some of its Properties—Ozone, its Preparation, etc.—The Experiments of Schoenbein and of Andrews establishing the Presence of Ozone in the Air—Houzeau's Researches—Effect of an Electric Discharge on Air—Dr. Angus Smith's Experiments on the Carbonic Acid in the Air—Methods practised by Pettenkofer and Angus Smith for determining the Carbonic Acid—The Simplest, or 'Household Method,' a Test of Ventilation—The Sources of the Minute Constituents—The Reciprocal Action of Plants and Animals.

AT the University of Edinburgh it is usual for graduates in medicine to contribute a dissertation on some scientific subject containing original observations and research. In 1754 the inaugural thesis of Dr. Black contained matter of such importance, treated in a manner so original, that it has raised the inexact observations of the alchemists and older chemists to the dignity of a true science. He showed that quicklime when exposed to the air becomes mild or carbonated, and at the

same time increases in weight ; that this increase of weight depends on the absorption from the atmosphere of a kind of air or gas (carbonic acid), which gas could be again expelled or recovered by heating the lime to redness ; and that the increase of weight of the lime was identical with the weight of the gas absorbed. This is the first instance of the employment of the balance in the investigation of chemical changes, the first beginning of quantitative research, and as chemistry is a science founded entirely on observations of the weights and volumes of matter participating in change of properties, it may emphatically be said that Joseph Black was the Father of Chemical Science, and that Edinburgh University was its birthplace. The classic research of Lavoisier, already mentioned, by which he proved the composition of air, is but an application of Black's mode of thought and experimenting to other substances. There can be little doubt now since the correspondence between these two illustrious men¹—dating 1789–90—has come to light, that Lavoisier was accustomed to look up to Black as his master, and regard his discoveries as leading to great reforms in chemical knowledge.

Black further observed that the gas absorbed by lime could be expelled by the addition of acids.

¹ Mentioned by Prof. Andrews at the Edinburgh meeting of the British Association. See Report for 1871.

He compared this air with that evolved during fermentation, and gave it the name of fixed air. After explaining the difference between caustic and mild alkalies as being due to the presence or absence of this fixed air, he next showed that it differed from atmospheric air in its chemical character and its relations to respiration and combustion; furthermore, that it had the power of combining with alkalies, and to a feeble extent neutralising them after the manner of acids, and finally that it was, in fact, the product of respiration and combustion.

When lime is burnt in a kiln, or more correctly speaking, very strongly heated, the fixed air (or as it is now called, carbonic anhydride, carbon dioxide, or carbonic acid) is expelled, and the lime becomes caustic.

There is this great difference between mild or carbonated lime and caustic or quicklime: the latter is soluble to some extent in water, and possesses the property of turning red litmus blue and yellow turmeric paper brown, while the former, better known as chalk, is insoluble and shows no such effects. This difference affords a ready test for carbonic acid. The production of carbonic acid (as I shall always call it) by fermentation is easily exhibited in the apparatus shown in fig. 19, where the gas is evolved and collected over water during the

process of fermenting sugar with a little brewers' yeast. That it is produced during combustion is



Fig. 19. Collecting the Gas evolved during Fermentation.

easily demonstrated by taking some lime-water, that is to say, a solution of caustic lime, and shaking it in a bottle in which a taper has been burnt: the clear liquid becomes turbid from the formation of chalk or carbonated lime. That it exists in the air is easily shown by placing a dish of lime-water in a room and it speedily becomes coated with an insoluble layer of carbonate of lime in the same way. That it is produced during respiration is evident by the arrangement of bottles shown in fig. 20: they both contain lime-water, and the tubes connecting them are so arranged that by placing the flexible one A in the mouth, and taking in

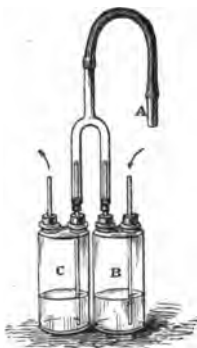


Fig. 20. Apparatus for showing Carbonic Acid in the Breath.

breath, air is drawn through the lime-water in the one vessel B, before entering the lungs, which causes little or no formation of chalk or carbonate of lime. By expelling the air from the lungs through the other tube C a very considerable quantity of chalk is formed, as may be seen by the increased milkiness of the liquid.

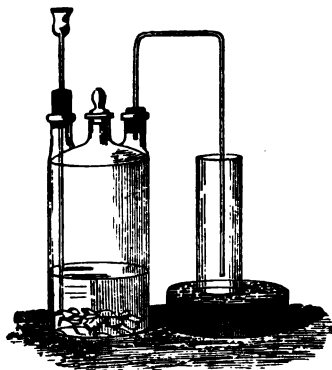


Fig. 21. The Preparation of Carbonic Acid.

Having thus considered the various sources of this gas, let us examine its usual method of preparation and chief properties. It is usually made from marble by the action of hydrochloric acid, in a two-necked bottle, such as that in fig. 21. The gas is invisible, being colourless; it is nearly odourless and tasteless. It is, however, more than half as heavy again as atmospheric air, and this may be easily made apparent by allowing it to flow

into a large jar: it will accumulate in the lower part, and can then be cautiously ladled out, and by taking care, poured into a glass beaker, which is carefully balanced; the balance then descends on that side into which the gas is poured, showing it to be heavier than the air which previously

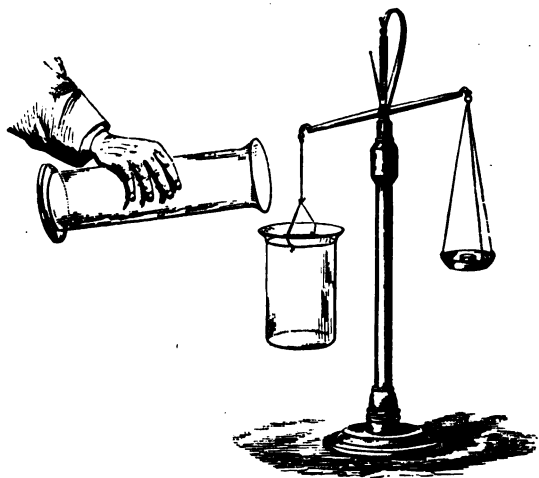


Fig. 22. The Weight of Carbonic Acid shown.

filled the beaker (fig. 22). This may further be demonstrated by allowing a light india-rubber balloon full of air, or a soap-bubble, to fall on to a vessel filled with the gas (fig. 23); the soap bubble will be seen to rebound from an invisible surface of air, and then remain tranquilly floating. This gas more resembles nitrogen than oxygen in

its non-support of combustion and power of extinguishing a flame. Thus by ladling out some of

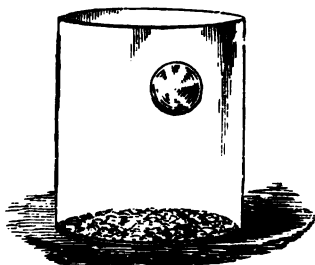


Fig. 23. Soap Bubble floating on Carbonic Acid.

the gas we may, by taking advantage of its weight, pour it some distance through the air on to a candle and extinguish it. The weight, too, and non-support of combustion may be shown by making a tree of lighted candles stand in a jar, to the bottom of which a supply of carbonic acid is allowed to flow by means of a tube; as the gas rises the candles are put out one by one. Then again a torch of blazing tow is easily quenched by immersion in the gas (fig. 24). The apparatus called 'l'extincteur' is an ingenious arrangement, by which water is forced out of a reservoir by means of the pressure exerted by carbonic acid as it is generated; the chemical action being in this instance the readiest means of obtaining mechanical power. On account of the extinguishing power of

carbonic acid a candle cannot continue to burn in a portion of air until it has exhausted the oxygen, but only until the combustion has produced sufficient carbonic acid to extinguish the flame. This may be shown by enclosing a candle in a bell-jar standing on a plate containing water so as to ex-



Fig. 24. Extinguishing Flame with Carbonic Acid.

clude air. On the plate and within the jar is placed a piece of phosphorus, and in the mouth of the vessel is a piece of hot bent wire. Now the candle burns till the product of its combustion causes it to flicker and go out. On turning round the jar so as to bring the hot wire and phosphorus in contact,

the latter will take fire and burn till it has consumed the remaining oxygen.

The other minute constituents of the atmosphere, besides carbonic acid, are water, ammonia, and ozone. That water is present in the air may be very easily visible by the effect of an ice-cold vessel. A deposit of dew takes place, and if the action is long continued a considerable quantity of water may be collected. If the temperature is lower than the freezing-point the deposit occurs in the solid form as hoar frost, as is easily seen when a mixture of ice and salt, having a temperature of -16° C. or 0° F., is put in a blackened metallic vessel. Now the source of this moisture in the air is evaporation from the ocean, and the extent to which it varies is due to temperature. Water, like all other liquids, has the property of giving off a portion of vapour at temperatures below the boiling point, and as the temperature rises the increase of vapour generated gradually gets greater until the whole of the liquid is converted into vapour. When water exists in the air as true vapour it is, of course, invisible; when, however, it is chilled we get a portion deposited as fog or snow. The extent to which water generally exists in the atmosphere is 1.5 volumes of aqueous vapour in 100 volumes of air.

This makes it a matter of great moment to

chemists, in the operations of analysis and other investigations, either to know precisely how much water vapour the air at the time of experimenting contains, or to have the means of getting air perfectly free from moisture. This latter end is arrived at by keeping oil of vitriol constantly exposed to the air. A piece of apparatus in constant use in chemical laboratories is a desiccator, that is, a bell-jar or other chamber, containing a vessel of oil of vitriol, over which is placed a floor of perforated zinc for substances which have to be kept dry. That atmospheric moisture is not entirely due to the existing oceans, lakes, and rivers, and the evaporation from their surface, is shown by the fact that a bell-glass held over a burning candle or gas jet is soon covered with condensed moisture, showing that it also is a product of combustion.

The presence of water vapour in the air is shown constantly by its effect on certain solid substances which have the property of combining with water, and in so doing becoming liquid. Chloride of calcium is one of these ; on account of this property it is said to be deliquescent. A compound of cobalt and iodine called iodide of cobalt—when perfectly deprived of water at a temperature of 140° C., or 284° F., takes the appearance of masses of plumbago or black lead, being black and lustrous.

Exposure to the air on a plate soon resolves small fragments of it into a rich green liquid, and this green liquid after a time increases in volume and becomes red, so that we have very pretty and curious phenomena accompanying deliquescence in this particular instance. A variation of the experiment may be made. A small quantity of the red solution is smeared over the bottom of two white porcelain dishes, in very small quantity; so little,

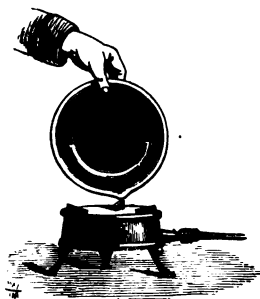


Fig. 25. The Dehydration of Iodide of Cobalt.

indeed, is necessary that the pink colour is scarcely visible even on close inspection. Heat the one dish over a gas burner, and at first a yellow tinge appears, but soon the whole of the bottom of the dish is covered with a rich, moss-like green tint; and now in the centre of this a black spot appears, which spreads gradually over the whole of the basin till it has the appearance of being black-leaded (fig. 25).

By dipping the outside of the basin into water it is cooled down, and while it is cooling the other dish may be heated, and thus the reverse operations go on side by side. By heating, the colours appear ; by cooling, they disappear ; and on close inspection it will be seen that their disappearance has resulted in the thin film of black shining substance being transformed into a pink liquid. The large, cool surface of the dish readily affords a means of condensing atmospheric moisture, and the peculiar properties of the cobalt iodide makes this condensation visible.

This is the most remarkable deliquescent substance at present known. When anhydrous, that is, entirely deprived of water, it is black. It combines with a certain definite proportion of water to form a green compound, and with a still larger quantity to produce a red substance ; hence the curious change of colours when hydrated, or combined with water to a greater or less extent. The little instrument commonly sold in the opticians' shops now, and called a chameleon barometer, simply indicates the amount of moisture in the air, and in its action depends on the fact that paper steeped in chloride of cobalt and chloride of calcium solutions has the property of becoming blue in a dry atmosphere, and red when moisture is present.

Without dwelling longer on the subject of atmospheric moisture, let us pass to the third minute constituent of the air, ammonia. This is a very light, colourless gas of a very pungent odour. Its specific gravity or relative weight, when compared with hydrogen, the lightest of all gases, is 8.5, or a little more than half that of atmospheric air. Its solubility in water is so great that at ordinary temperatures 1 volume of water absorbs 700 times its bulk of ammonia.

Notwithstanding the pungency of this gas and its extreme solubility, the fact of its being present in the air without our perceiving it is accounted for by its extremely minute proportion. This amounts to no more than $3\frac{1}{2}$ volumes in 10 million volumes of air. In other words, if we take 10 million gallons of air, the ammonia gas present would not be more than $3\frac{1}{2}$ gallons. To detect and accurately estimate the ammonia when in so small a proportion requires the use of very delicate chemical tests, very great care, and great patience in the performance of the experiments. From the numbers obtained by Fresenius, who is perhaps the highest authority we have on analytical matters, we learn that the ammonia present in the air varies by night and day, the amount being larger at night by about one-third. Mr. Horace Brown in 1870 published the latest experiments on the determina-

tion of the amount of ammonia in the air, and by making use of the valuable process known as Nessler's test, he obtained numbers which there is every reason to believe are accurate—firstly, because they are not widely different from those of Fresenius, and, in fact, are closer to them than those of any other experimenter; and secondly, the numbers are just a little higher—a variation one would be led to look for, as the method pursued by Fresenius would be slightly in error, probably from loss of ammonia. Here, in the form of a table, are the numbers given:—

Table showing the Ammonia in 10,000,000 parts of Air.

HORACE BROWN, 1870.

1869				By weight	By volume
Dec.	6	.	.	2·084 parts	3·542 parts
„	8	.	.	2·15	3·65
„	9	.	.	1·805	3·068
„	11	.	.	1·812	3·18
„	12	.	.	2·19	3·72

FRESENIUS.

By day	.	.	.	·98	1·666
By night	.	.	.	1·69	2·87
Mean	.	.	.	1·33	2·26

One remarkable property of ammonia is, its power of combining with acids to form solid compounds, called salts. This is shown by the following illustration. Two large glass globes contain

each about a gallon of a gaseous substance of very opposite properties; in the one is ammonia, in the other hydrochloric acid. They are at first closed by glass plates, but by bringing these in contact and gently removing them the two gases rush together (fig. 26) with great violence, combining instantly under great evolution of heat to



Fig. 26. The Combination of Ammonia with Hydrochloric Acid.

produce a white, solid, snow-like substance known as sal-ammoniac, ammonium chloride, or ammonia hydrochlorate, a substance in many respects resembling common salt.

By passing air through long tubes containing broken glass moistened with weak liquid acid, the ammonia may be abstracted, and so condensed in

the form of a solution. A very striking experiment arranged to show the exceeding solubility of ammonia gas in water is the following. The large glass flask (fig. 27), which has its neck inverted below water, coloured red with litmus, is full of ammonia, but the mouth of the flask is closed from contact with the water by means of quicksilver. By raising the

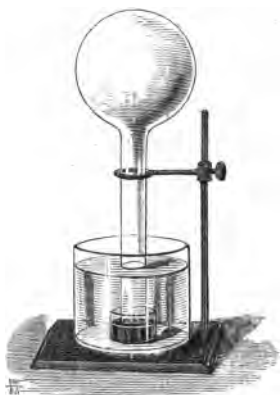


Fig. 27. Experiment to show the Absorption of Ammonia by Water.

neck of the flask out of the quicksilver into the water the liquid dissolves the gas so rapidly, that it rushes upwards with an alarming rapidity, the violence being almost sufficient to break the vessel ; at the same time it will be seen that the red litmus solution becomes blue by the action of the ammonia.

It cannot be doubted that the ammonia present in air is consumed by living plants ; indeed, it is this consumption that causes the difference in the amounts found by day and night. This difference was found by Fresenius to be in the ratio of 1 for the daytime and 1·7 for the night. The smaller quantity present during the day may be due to two causes : 1st. During the daytime more ammonia is absorbed in the nutrition of plants ; 2nd. The ammonia accumulating during the day and night together is dissolved and precipitated by the dew at sunrise.

The next and last gaseous substance to be considered as a constituent of the air is ozone. It is present in only very minute quantities, although in greater proportion than ammonia. Its nature and properties are most remarkable. The researches of many chemists have placed the supposition beyond doubt that it is a condensed form of oxygen with greatly exalted chemical activity. The energetic nature or oxidising power of ozone surpasses that of oxygen as much as the activity of oxygen excels that of air. The preparation of ozone on the lecture table in order that some of its properties may readily be demonstrated is thus effected (see fig. 28) :—The oxygen used is delivered from an iron cylinder of the compressed gas, but this is merely a matter of convenience and not of necessity.

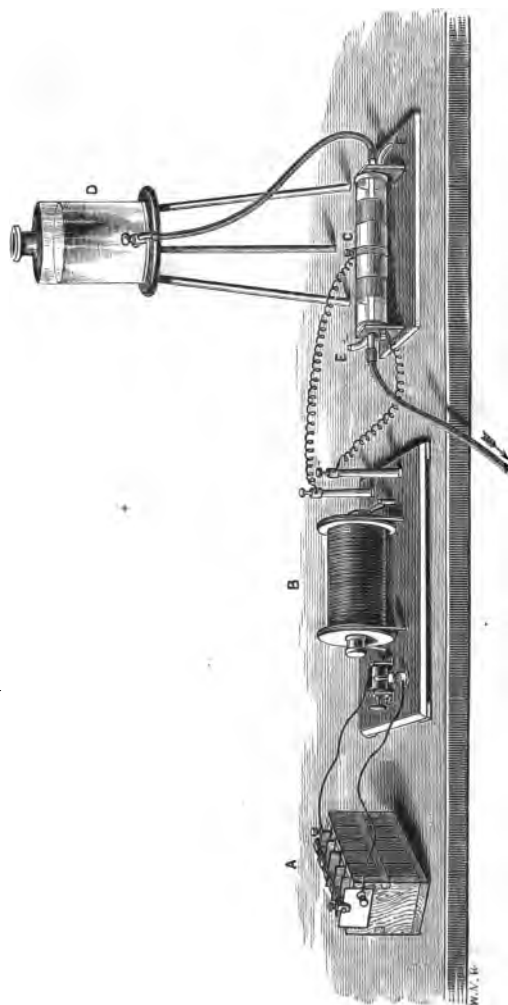


Fig. 28. The Preparation of Ozone.

A Galvanic battery to excite the coil. B Induction coil to produce an electric discharge. C Ozonising tube. D Reservoir of ice-cold water to cool the ozone tube. E and F Tubes for the passage of air or oxygen to be ozonised.

This gas when allowed to flow out is cooled by ice, afterwards dried with oil of vitriol, and passed into an apparatus known as an ozone generator.

The particular form of apparatus, made by Messrs. Tisley and Spiller (fig. 29) is the one by which the best effects are produced. It consists of an iron tube turned very truly on the outside ; within it passes a current of water by means of the tubes C C. Outside this metal cylinder is one of glass, which is only a very slight degree larger, and by an arrangement of tubes D D, the cooled air from the iron bottle can be passed through the annular space between the glass and metal cylinders. Part of the outer surface of the glass cylinder is covered with tin foil G. Now the outer tin-foil coating and the inner metal cylinder are connected with the two fine wires, from a Ruhmkorff induction coil, at the points E and F. The coil furnishes a constant electrical discharge ; this discharge passes through the annular space before mentioned, and the electrification of the current of oxygen or air endows it with new properties, which are very curious. It now possesses a very peculiar smell, a smell such as accompanies the sparks from an electrical machine, and which is often recognised when two stones are broken together with the emission of sparks. This odour is no doubt what is perceptible near a place which has been struck by lightning, and which is

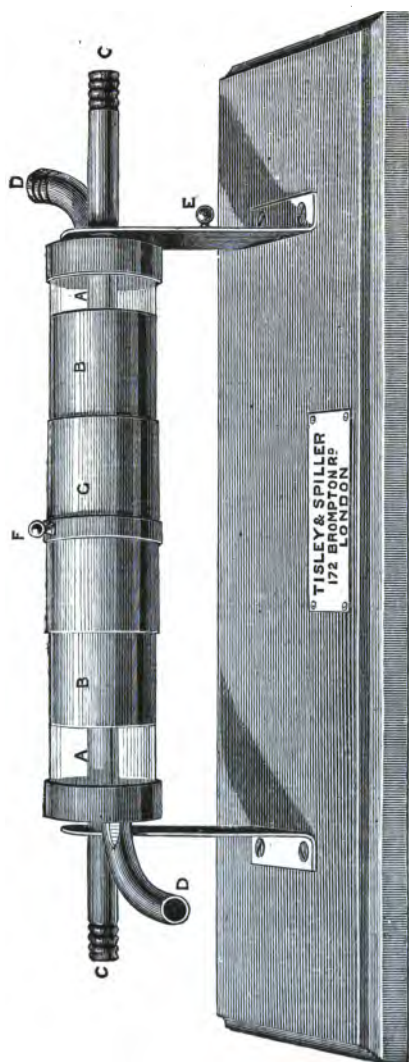


Fig. 29. New Ozone Generator.

Woodcut from 'Nature,' by permission of Messrs. Macmillan and Co.

ignorantly described as a smell of sulphur. Besides this peculiarity, the gas has the property of attacking the mucous membrane of the nose and throat, giving rise to a painful soreness. It eats through all india-rubber tube with which it comes in contact in a few minutes, it has the power of bleaching indigo, it oxidises silver and mercury, and of all these properties none are those possessed by common oxygen. It exercises its powers in such an energetic manner that the smallest possible quantity can readily be detected. There is one particular action which is peculiarly adapted for its ready detection: it sets iodine free from iodide of potassium. Now the merest trace of iodine in the free state has a remarkable power of striking an intensely blue tint when in contact with boiled starch; so that by placing starch paste and iodide of potassium together on paper, we have a mixture which if exposed to ozone acquires a deep blue colour. The depth of the blue depends on the amount of iodine liberated, and the amount of iodine liberated corresponds to the quantity of ozone present, so that by exposing iodised starch paper to a current of air in a dark place, if equal amounts of air are taken we can by the varying depths of tint compare the amount of ozone in one place with that in another, or the

quantity present on one day with that on another occasion.

For some time it was by no means certain, after all said and done, that ozone was really present in the air. Nitric peroxide and nitric acid, also chlorine, could act on iodised starch-paper and cause the same blue colour to appear.

The matter was finally set at rest, however, both by Professor Andrews of Belfast and Professor Schoenbein of Basle in the same year, 1867, and by different methods. Andrews had previously shown that ozone, whatever be the method of its production, is quickly destroyed at the temperature of 237°C. , or 458°F.

An apparatus was therefore fitted up by means of which a stream of atmospheric air could be heated to 260°C. , or 500°F. , in a large globular glass vessel. On leaving this vessel, the air was passed through a U-tube, the sides of which were moistened internally with water, while the tube itself was cooled by being immersed in a vessel of cold water. On passing air through this apparatus at the rate of three litres per minute, the test-paper of iodised starch was distinctly tinged in two or three minutes, provided no heat was applied to the glass globe. But when the temperature of the air, as it passed through the globe, was maintained at 260°C. , not the slightest action occurred upon the test-paper,

however long the current continued to pass. On the other hand, when small quantities of chlorine or nitric acid vapour, largely diluted with air, were drawn through the same apparatus, the test-paper was equally affected, whether the glass globe was heated or not. Hence the body in the atmosphere which decomposes iodide of potassium is identical with ozone.

The proof employed by Schoenbein was of a different nature. When white protoxide of thallium is exposed to the action of ozonised air, it is converted into brown peroxide. If we take a solution of sulphate of thallium to which a considerable amount of potash has been added so as to throw down the protoxide of thallium, papers dipped in this and exposed to the ozone instantly become brown; but it might be justly stated that if such papers be exposed to the air, a small quantity of sulphuretted hydrogen, a gas which escapes from decomposing organic matter, would also cause a similar effect. The difference between the browning in the two cases is this, that when due to peroxide of thallium, a paper steeped in tincture of guaiacum if pressed against it becomes blue where brought into contact with the brown thallium stain. The blue is the result of oxidation, caused by the peroxide of thallium; and it is needless to say that any other kind of brown stain would fail to show

such a reaction with guaiacum solution. This thallium test is not so sensitive as the starch and iodide of potassium. Schoenbein notices a curious fact, namely, that the air is very rich in ozone after heavy falls of snow; this he remarked especially on January 6, 1867, at Basle, when he detected ozone by means of the thallium test.

A further most valuable contribution to our knowledge of atmospheric ozone is afforded by the researches of M. Houzeau, who makes use of a new test for this substance. This consists of red litmus papers, half of which are dipped in a solution containing one per cent. of neutral iodide of potassium. Should this paper come in contact with ammonia gas it will entirely assume a blue tint; with chlorine or acids it will be unacted on, while with ozone only that part of the paper will turn blue which is impregnated with the iodide of potassium. This is occasioned by the oxidation of the potassium to potash, the alkali thus formed changing the colour of the litmus. Houzeau's most important conclusions are the following:—

1st. That country air contains an odorous oxidising substance, with the power of bleaching blue litmus without previously reddening it, of destroying bad smells, and of blueing iodised red litmus.

2nd. That this substance is ozone.

3rd. That the amount of ozone in the air at

different times and places is variable, but this is at most $\frac{1}{700000}$ of its volume, or 1 volume of ozone in 700,000 of air.

4th. That ozone is found much more frequently in the country than in towns.

5th. That ozone is in greatest quantity in spring, less in summer, diminishes in autumn, and is least in winter.

6th. It is most frequently detected on rainy days and during great atmospheric disturbances.

7th. That atmospheric electricity is apparently the great generator of ozone.

It has already been stated that ozone is a condensed form of oxygen. This was originally suggested by Odling and confirmed since by the experiments of Tait and Andrews, of Soret and Sir Benjamin Brodie. It has, indeed, been well ascertained that 3 volumes of oxygen become condensed by electrification to 2 volumes of ozone, and by heat we can expand it again into 3 volumes of common oxygen. Just as we have carbon exhibiting different properties in the three forms of charcoal or lampblack, plumbago or black lead, and diamond, so we have oxygen in the two forms of common oxygen and ozone. Such modifications of a simple substance are called allotropic conditions.

It is most probable that during thunderstorms

the violent or disruptive discharge of electricity from the clouds causes, besides ozone, a production of nitric acid, or at any rate of nitric peroxide, which, in contact with water, becomes nitrous and nitric acids. The evidence we have is, that rain-water contains minute traces of these substances, and that by making a miniature thunderstorm—as may be done by passing an electric discharge through a glass bulb (fig. 30)—the red fumes of nitric peroxide are rendered visible. Mr. J. M. Thomson, of King's College, has been engaged on experiments in this direction, and he finds that if a certain bulk of air be enclosed in such an apparatus as this, by passing electric sparks for a sufficient time all the oxygen of the air combines with

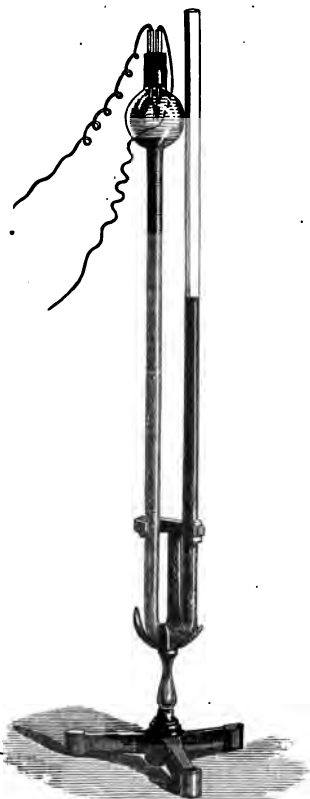


Fig. 30. Method of showing the Effect of Electric Sparks on Air.

enough of the nitrogen to produce the red nitric peroxide gas, the absorption of which by the water in the tubes causes a considerable contraction in volume of the original quantity of the air. When the discharge has gone on for twenty minutes within the glass bulb, the contraction amounts to that shown by the lower level of the liquid in the right-hand tube, fig. 30. The blue litmus, too, with which the tube is charged, though remaining unaltered in the right, becomes reddened in the left limb from the production of acid in the bulb.

So far the gaseous constituents of air may be divided into—

Oxygen and nitrogen	.	Chief constituents.
Water	} Constituents present in minute quantities.
Ammonia	
Carbonic acids	
Ozone	
Nitrous and nitric acids	

Besides these there are other gases present, occasionally, in more or less quantity, which may be called adventitious constituents, as hydrochloric acid from volcanic eruptions and chemical works ; sulphurous acid from the combustion of coal containing sulphur ; sulphuretted hydrogen from decaying organic matter ; and marsh-gas from the same source.

Returning to the consideration of carbonic acid and the amount present in the air, it has been observed that even in open places, many causes, amongst which is the state of the weather, give rise to very considerable variations in quantity, which are confined within very small limits.

To illustrate this variation are here tabulated some of the results obtained by Dr. Angus Smith, for a complete account of which reference must be made to his work on 'Air and Rain.'

	Volumes of Carbonic Acid in 10,000 Volumes of Air
On the mountains and moors of Scotland—mean of 57 analyses	3'36
In the streets of Glasgow—mean of 42 analyses	5'02
London, N., N.E., and N.W. postal districts—mean of 12 analyses	4'445
London, S. and S.W. districts—mean of 30 analyses	4'394
London, E. and E.C.—mean of 12 analyses	4'745
London, W. and W.C.	4'115
Manchester streets, ordinary weather	4'03
During fogs in Manchester	6'79

Here 3'36 volumes of carbonic acid in 10,000 parts of air may be taken as the normal amount. In towns the amount rises, and is greatest in those streets or parts which are the most populous. The state of the weather, again, influences the amount of carbonic acid; during fogs it may be increased to 8 volumes per 10,000.

It was stated by Dalton that in his day the air

of Manchester could not be distinguished by chemical analysis from that of Helvellyn, but Angus Smith has given us analyses which not only show the differences between town and country air—between the air of one town and that of another—but between that of various London streets, and even the variation in the air between the back and front doors of the house where his laboratory is situated in Manchester.

We have thus made evident the very great variability within very small limits, and the extraordinary constancy in composition of the air. All those who are susceptible to atmospheric influences prefer mountain air to that of the town, and in London prefer the air of the West End and the suburbs to that of the City.

Notwithstanding the evidence of one's sensations, it is impossible to say that the physiological effect which causes life to be more enjoyable in one place than another is due to the air we breathe, but it is now placed beyond conjecture by these very exact chemical analyses that such is actually the case. A vitiated atmosphere pervading the streets and hanging over cities is undoubtedly the chief cause of that longing for the country experienced by all dwellers in towns.

In the finest summer weather or in the depth of winter no one who takes the trouble to drive

from Bayswater to Whitechapel could, I think, fail to notice the deterioration in the air as the more thickly populated parts were reached, and chemical analysis alone has shown us the cause of this. In foggy weather, when the air of a town is not carried away by a wind, but hangs over the house-tops like a cloak, the carbonic acid is not disposed of by diffusion and gentle currents of heated air with the same rapidity with which it is produced, it therefore accumulates even to a very unwholesome amount. At the same time the condensing aqueous vapour of which the fog consists carries down sooty particles and tarry matter from partially burnt coal, which greatly aggravates the other evil.

There can have been but few people in London last November (1873) who did not during the heavy fogs experience some difficulty in breathing, in getting a sufficient supply of oxygen for the lungs to carry on the vital processes in the usual manner. On one of the last days of its duration many people one met in the streets seemed to find breathing painful ; they were apparently now and then gasping for breath, and the thought occurred that it was by no means impossible by a continuance of such weather for a whole cityful of people to be suffocated. The mortality at that time amongst the sick was very great, and the state of the air caused the death of many prize-fed

beasts sent to the Cattle Show at the Agricultural Hall at Islington.

It must not be inferred from the determinations of carbonic acid here made so much of, that there is something especially poisonous or dangerous in minute quantities of this gas; this is not so much the case, as that it is a measure of other accompanying impurities in the air; 'not that we know certainly of any positive evil, which it can do of itself in these small quantities, but because it almost always comes in bad company.' It should also be borne in mind that for every increase in carbonic acid there is a decrease in oxygen, so that we have a double effect, a subtraction of the life-giving principle of the air, and an addition of a somewhat noxious substitute. It will be seen further on, that under the worst conditions out of doors, fog alone excepted, people constantly within doors submit themselves to a worse atmosphere; it is on this account that the estimation of carbonic acid is of great importance.

The quantity of carbonic acid is estimated by placing the air in a graduated tube over mercury (see fig. 12). A ball of potash at the end of wire absorbs the gas, forming solid carbonate of potash, and causing a diminution in the volume of the air. Accurate measurement before and after the experiment, and corrections for alterations in temperature

and barometric pressure enable the operator to arrive at the true amount of the gas removed by the potash. All Dr. Angus Smith's numbers were obtained in this way. The method of Pettenkofer which I have practised is somewhat different.

Two graduated tubes, each of which is divided into 600 equal parts, the division being engraved on the glass, are two of the most essential pieces of apparatus. These instruments are called burettes, and are used for what is called the volumetric method of analysis. Chemists generally find it more convenient to use the French weights and measures, and in accordance with that system each 10 divisions on the glass would hold 1 cubic centimètre of water, which has the weight of 1 gramme; therefore the whole buretteful contains 160 cubic centimètres or 60 grammes. The instruments are filled with two different solutions, one oxalic acid and the other baryta water. The strength of the liquids is so adjusted that a measure of the acid will exactly neutralise an equal measure of the alkali. This neutralisation is thus determined. Taking, for example, 30 cubic centimètres of this baryta water, one drop removed from the flask with a glass rod leaves a brown stain upon a disc of yellow turmeric paper; if, now, 29 cubic centimètres of oxalic acid be run into the baryta water and well mixed up, the brown coloration is still pro-

duced, but it is fainter; but on adding oxalic acid drop by drop the brown stain finally ceases to be perceptible.

The end of the process is now come, the causticity of the baryta water has been entirely destroyed by the oxalic acid, or in other words it has been neutralised. Now the oxalic acid is made of such a strength that each cubic centimètre will produce exactly the same effect towards neutralising the baryta as $\frac{1}{1000}$ th of a gramme weight, or 1 milligramme of carbonic acid.

The next part of the process is to take an accurately measured flask, pump air into this from the desired locality with a pair of bellows, and then introduce from a burette 60 cubic centimètres of baryta water. An air-tight cap is then fitted on, the baryta water is swilled round the sides of the flask, and it is allowed to stand for at least two and not longer than twenty-four hours. The baryta water is then found to be quite milky, all the carbonic acid having gone to form carbonate of baryta. Pour out the turbid liquid, and allow it to settle, measure off half of the clear solution, and carefully proceed, as in the first operation, to add oxalic acid, and test with turmeric paper. If now 20 cubic centimètres of oxalic acid neutralise the liquid and only half the baryta water was experimented on, 40 cubic centimètres would have been

the whole quantity required ; this would indicate that 20 cubic centimètres had been neutralised by the carbonic acid. Such is an outline of the process. Its drawback is the bulkiness of the air-flasks, otherwise it is very convenient and very accurate.¹

It has already been shown, that the outside air contains an amount of carbonic acid varying between 3 and 6 parts in 10,000 volumes, but it will be seen later on in close places, such as crowded buildings, this rises to the extent of even 25 volumes in 10,000 of air. This, if the pollution of the air consisted of merely the addition of carbonic acid, would be a sufficiently alarming amount, but it is not ; there is likewise the abstraction of oxygen to be taken into account, and the various exhalations from the lungs which ought to be regarded as poison.

It has been experimentally proved that *when the heat is not excessive the organic matter charging the air of crowded places rises in amount as the carbonic acid increases*, so that in the estimation of carbonic acid we have a measure of the foulness of the air or, as it may be termed, want of ventilation. Coming from the outside into a room in the condition we call 'close' or 'stuffy,' we enter an atmosphere which does not contain less than 6 volumes of carbonic acid in 10,000 of air. But the 'closeness' is detected generally by the nose, and is the

¹ Fuller details are given further on.

effect of organic exhalations, rather than of carbonic acid; nevertheless the two are hand-in-glove, so the carbonic acid, which can be measured with greater certainty and ease than the other pollution, tells the story for both.

A very easy method of finding the extent to which the air of dwelling-houses is contaminated has been devised by Dr. Angus Smith. This 'household' method of testing the air depends upon the fact that a certain definite quantity of carbonic acid is necessary to cause a visible precipitate in half an ounce of lime water. A practical illustration will be the best explanation possible. Taking five wide-mouthed bottles measuring respectively 8·05 ounces, 9·13, 10·57, 12·58, 15·6, or in cubic centimètres 228, 259, 299, 356, 443, if into the first or smallest there be put half an ounce of perfectly clear lime water, the bottle being previously wiped quite clean and dry, and if after shaking so as to mix the air and lime water, on examining it carefully, there be 'no precipitate' pass on to the next bottle a size larger, if with the third it is somewhat doubtful, but with the fourth bottle the precipitate is well seen, let us take the exact point to be between the third and fourth bottles. Now on referring to a table previously constructed we find that this indicates between 5 and 6 volumes of carbonic acid in 10,000 of air, say 5·5, *i.e.* $5\frac{1}{2}$ —an

amount very likely to be present, and which indicates a fairly ventilated place.

'Let us keep our rooms so that the air gives no precipitate when a 10½-ounce bottleful is shaken with half an ounce of clear lime water.' We shall thus secure good ventilation, the carbonic acid will be less than 6 per 10,000. For sleeping apartments such a rule is most valuable, for it is there we spend about one-third of our lives, and during sleep we have great need of all the oxygen we can get; for, according to Pettenköfer and Voit, the body is then engaged in storing it up for future use.

The minimetric method of analysis of Dr. Angus Smith resembles the already mentioned 'household method' in its extreme simplicity. The principles upon which it depends are, that a measurable quantity of carbonic acid is required to produce an appreciable turbidity in half an ounce of baryta water, and that the *exact appearance* of a certain precipitate of carbonate of baryta *can be retained in the memory with marvellous exactitude*. If we place half an ounce of baryta water in a bottle, having space for two ounces of air, and attach by means of an india-rubber cork and tubes a ball-syringe fitted with valves, and holding two ounces (see fig. 31), we have a means of pumping air in measures of two ounces at a time through the baryta until the particular precipitate most easily retained by

the memory is produced. Such a precipitate is that yielded by shaking a 23-ounce bottle of air with half an ounce of clear baryta water, when the air contains 0.04 volume per cent., or 4 volumes of carbonic acid in 10,000 of air. The actual amount of carbonate of baryta formed is 0.00224 grammes. I have adopted the plan of making a solution which will always yield this precipitate. Dissolve 7.4536 grains or 0.4821 gramme of pure dry carbonate of soda—heated to redness in a crucible—in 5 pints of water; $\frac{1}{4}$ ounce carefully measured with a pipette will with $\frac{1}{4}$ ounce of baryta water solution produce the standard precipitate.¹

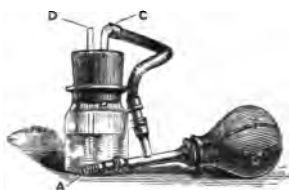


Fig. 31.

When the ball-syringe is squeezed the contained air is discharged by the tube with the free end A; on relaxing the pressure air enters from the bottle by

the india-rubber junction tube C; at the same time the external air, passing down the straight glass tube D bubbles through the baryta water.

In making an analysis first attach the syringe and cork to the empty bottle, give two or three strokes to get the air changed for that of the place

¹ Or the same quantity may be dissolved in 800 cubic centimètres of water, and only two cubic centimètres taken instead of $\frac{1}{4}$ ounce.

to be tested. Put in the baryta water and shake it well; this counts for one ballful. Close the entrance-tube with the forefinger, squeeze the ball, remove the finger, and relax the pressure. When the ball is fully expanded shake up again. This is the second ballful. Repeat the operation as often as is necessary until the liquid is of such a turbidity as that produced by the standard precipitate. Count the number of ballfuls of air, and refer to one of the tables here given.

Number of ballfuls of air	Volumes of Carbonic Acid in 10,000 of Air			
	With a 2-oz. ball	1-oz. ball	$\frac{1}{2}$ -oz. ball	$\frac{1}{4}$ -oz. ball ¹
1	—	—	—	—
2	—	—	—	—
3	—	—	—	—
4	—	—	—	—
5	8.8	—	—	—
6	7.4	—	—	—
7	6.3	12.6	—	—
8	5.5	11.0	—	—
9	4.9	9.8	—	—
10	4.4	8.8	17.6	—
11	4.0	8.0	16.0	—
12	3.7	7.4	14.8	—
13	3.4	6.8	13.6	—
14	3.2	6.4	12.8	—
15	—	5.8	11.6	23.2
16	—	—	11.0	22.0
17	—	—	10.4	20.8
18	—	—	9.8	19.6
19	—	—	9.3	18.6
20	—	—	8.8	17.6
21	—	—	8.4	16.8
22	—	—	8.0	16.0
23	—	—	7.7	15.5
24	—	—	7.4	14.8

¹ The utility of a ball of this size is doubtful.

The accuracy of the test diminishes rapidly when a 2-ounce ball is used with air containing more than 7 volumes of carbonic acid per 10,000. A 1-ounce ball may be used up to 12.6 volumes per 10,000. It is an immense convenience to have the half-ounce measure for the baryta water marked on the glass with a diamond, and to engrave the suitable reference table upon the bottle itself. It is also worthy of remark that a 1-ounce ball may be used with a 2-ounce bottle, but not *vice versa*; and under such conditions when the baryta water is first shaken up instead of 1 this must count as 2 ballfuls. Again, it is a great advantage to have a label written in lead pencil of such a shade that the standard precipitate just prevents its being seen through the bottle, gummed on to one side. If the words 'Not Done' or 'Unfinished' be taken, as long as they are legible the operation is uncompleted. Without the standard solution of carbonate of soda there is some inconvenience attending the *exact* determination of carbonic acid, though there is none in comparing the air of a room with the air outside the house, and saying that it contains twice, thrice, or four times as much. Again, we may make an experiment on the external air, and assume that it contains the smallest probable quantity—for instance, of a London street, say 4 per 10,000 volumes—test the

air of a room, and then by comparing it with that of the street it is easy to fix the minimum of carbonic acid present. For medical officers of health and others engaged in sanitary inspections who cannot be expected to know anything of analytical operations, especially those of such a delicate nature as gas analysis, an apparatus of this kind is invaluable in enabling them to express an opinion as to the wholesomeness or otherwise of an apartment, schoolroom, or public building.

So great is its simplicity and accuracy that a man who positively is ignorant of the operation he is performing may yet estimate the carbonic acid in the air of a room to the accuracy of 1 volume in 10,000 if the quantity be large, and the error is only half as great with small quantities. The error, too, is on the safe side, the estimated amount being rather under than over the actual quantity with a cautious operator. When the air of a place feels 'close' on first entering, a 1-ounce ball and bottle is most convenient; if very 'close,' such as the air of a crowded public building, the better size is $\frac{1}{2}$ an ounce, the same quantity of baryta water being taken in every case. Considerable experience has shown that the silk valves used in the T-tubes often cease to act properly, and the following contrivance is found to give what is for a time a ser-

viceable and more easily constructed apparatus. The ivory nozzle of a ball-syringe has about a foot of black india-rubber fitted to it; the best size is under $\frac{1}{4}$ of an inch in diameter. This has a longitudinal slit made in it with a sharp penknife of the length of an inch, and is attached to the bent tube C (fig. 31) of the bottle, and on to the tube D is fitted about $\frac{3}{4}$ of an inch of the same india-rubber. By bending this piece of india-rubber with the finger all exit from the bottle is closed, but when the ball is squeezed air escapes by the slit in the connecting tube, but cannot re-enter until the finger is removed from D. This simpler form of apparatus has been found to be most efficient.

Considering the carbonic acid exhaled by the animal world, poured into the air by furnace fires, and slowly evolved during the decay of organic matter; remembering also that this process goes on without ceasing, is it possible in course of time that the oxygen now at our disposal will be so used up and reduced in quantity as to be no longer capable of supporting life? Dumas and Boussingault have made a calculation to answer this question, and they consider that under the worst possible conditions scarcely less than 800,000 years would have to elapse before the mass of the atmosphere would be deprived of its oxygen by animals

living on the earth's surface, so vast is its extent. But the immense number of experiments of Angus Smith show numbers which exactly tally with those of Regnault, made twenty years previously: hence there is not the slightest diminution in oxygen in so long a time, though the average error in Regnault's and Angus Smith's work is not greater than $\frac{1}{100000}$, so that in eight years we should be able to detect some alteration in the numbers. It is evident, then, that there is not merely an absorption of carbonic acid by rocks or the waters of the ocean or rivers; there must actually be a restitution of the oxygen consumed in the production of carbonic acid. As Priestley originally discovered and made evident, plant life on the earth's surface is the cause of this. Just as a certain quantity of oxygen is necessary for the maintenance of the animal kingdom, so also an amount of carbonic acid is required as food for plants. The respiration of plants is effected through the leaves, which are filled with stomata, apertures which may be likened to small mouths; by these they inhale carbonic acid and exhale oxygen. The green colouring matter of foliage or chlorophyll performs a similar function, though the chemical process is different, to that of the red blood corpuscles, for while the red blood corpuscles are carriers of oxygen from the air to the internal organisation of animals, the

chlorophyll acts as a supplier of carbon to the tissues of plants ; the oxygen, not being wanted, is returned to the air. The solar rays play an important part in this chemical action of plants, for it is by the influence of sunlight that the nutritive process goes on. Regarding the changes which the oxygen undergoes, we have first its carbonisation and hydrogenation, by means of animal life, and return to the air ; then its decarbonisation by means of vegetable life, and restoration in its original state. Or if we consider the carbon in the same way, we have the consumption of vegetable products as food, the oxidation of animal tissues, and the oxidised carbon committed to the air, then again a reabsorption of the carbonic acid and its deoxidation in the form of vegetable products. So it follows that the air is a reservoir of oxygen for the use of animals, and a reservoir of carbonic acid for the use of plants.

In the round of operation by which the carbon of plants is derived from carbonic acid, water also plays as prominent a part. In carbonic acid and water there is contained far too much oxygen to admit of the building up directly of such substances as cellulose or woody fibre, sugar, and starch ; by a process of condensation or coalition the carbon and hydrogen combine, while some of this oxygen is got rid of. The reverse operation takes place

when we burn a piece of wood ; the oxygen entering into its composition is insufficient by any rearrangement with the carbon and hydrogen atoms to form again the substances from which it took its origin, but as the wood acquires oxygen from the air, its carbon and hydrogen atoms are separated, and it is again resolved into carbonic acid and water, and at the same time there is evolved an amount of heat corresponding to the chemical energy exerted by the sunlight when it promoted the growth of the trees. Thus can we store and concentrate the sun's rays so as to bring their action to bear on any one particular place at any given time. Such is the case where woods are planted and grown for twenty years to be cut down and thrown into a furnace. When primæval vegetation in the form of coal gives the warm glow of the fireside in winter it is the sunbeams of a bygone age whose influence we feel.

APPENDIX TO CHAPTER II.

For those who are interested in the method of Pettenkofer for determining the amount of carbonic acid in the air, here are given full directions, with some novel practical details which simplify this process.

Pettenkofer's Method.—The flasks used for containing

the air are the ordinary long-necked bolt-heads ; they should be able to hold 5 to 10 litres (1 to $1\frac{1}{2}$ gallon). Their capacity is determined by measuring in water at a temperature of 16° C. (or 60° F.) out of a litre-flask. When the water has risen up to the neck to a convenient height, an additional 60 cubic centimètres are added. A mark is then made on the glass at the level of the liquid with a sharp file or cutting diamond, and the application of a piece of hot glass or iron cuts off the superfluous portion of the neck. The complete capacity of the measured flask, supposing 7 litre-flasks full of water have been emptied, is 7,060 cubic centimètres ; its capacity for air, however, during an experiment is 7,000 cubic centimètres, because 60 cubic centimètres of baryta water are introduced.

Regarding the strength of the baryta water, this may be made by dissolving 7 grammes in 1 litre of water, if the baryta be the *crystalline hydrated* compound.¹

The oxalic acid is made by dissolving 2.8636 grammes of pure crystallised oxalic acid in 1 litre of water. As 126 parts by weight of oxalic acid ($C_2H_2O_4 \cdot 2 Aq.$) neutralise the same quantity of baryta or lime-water as 44 parts by weight of gaseous carbonic acid (carbonic anhydride), $\frac{126}{44} = 2.8636$, the whole of this solution is of the same strength as 1 gramme of carbonic acid ; consequently, 1 cubic centimètre = 1 milligramme.

Take 60 cubic centimètres of the baryta water, measured with a pipette, and compare it in the following manner with the oxalic acid. Run into it 30 cubic centimètres of the acid solution, and stir with a glass rod ; if no permanent precipitate or turbidity is caused, the

¹ That the baryta should contain no trace of potash or soda is of the greatest importance.

baryta should be diluted with an equal bulk of water. Supposing, however, that the liquid is no longer clear, take a drop out with a glass rod and let it touch a filter paper coloured with tincture of turmeric. As these test papers are bleached by exposure to light, it is best to store them away in an earthen jar in a dry place. Great care should be taken during the experiment to keep the air of the laboratory free from ammonia, even the quantity present in tobacco smoke is sufficient to vitiate the result. If a red or brown spot be seen, the baryta water is not yet neutralised ; more oxalic acid is cautiously added, and as in time the brown spots get fainter, increased caution is necessary in the addition of the acid. Finally, when the liquid shows no action on the turmeric paper, the baryta has been completely neutralised or saturated. If 60 cubic centimètres of the acid have been used, then 1 cubic centimètre of acid = 1 cubic centimètre of baryta. It is rarely that this occurs, and it is not worth while diluting or strengthening the solution so as to obtain very exactly this relation of the two liquids. If 90 cubic centimètres of acid were required for 60 cubic centimètres of baryta, then each cubic centimètre of the alkali would be capable of neutralising $1\frac{1}{2}$ milligramme of carbonic acid. There is, therefore, as much alkali in 60 cubic centimètres as there should be in 90 cubic centimètres, and the liquid should be made up to this bulk.

To make the analysis, take the flask in a clean and dry state, let it remain in the air for a quarter of an hour to gain its proper temperature, then with a pair of bellows blow air into it ; note the temperature, measure in 60 cubic centimètres of baryta solution, cover quickly with a capsule of india-rubber, or tin plate and india-rubber, such as are made for pickle and preserve jars, swill the

liquid round the sides of the glass, and let it remain from two to four hours. Pour out the liquid into a little beaker glass or flask, cover with a watch glass, allow the carbonate of baryta to subside, take off with a pipette 30 cubic centimètres of the clear liquid, and cautiously add oxalic acid from the burette until the alkalinity is destroyed.

Suppose it had been found in the first experiment that 63 cubic centimètres of acid were equal to 60 cubic centimètres of baryta, and that in the second case 26.5 cubic centimètres had been used ; then as only half the 60 cubic centimètres of baryta solution had been experimented on, multiply by 2 :

$$26.5 \text{ C.C.} \times 2 = 53 \text{ C.C. of oxalic acid.}$$

But 63 cubic centimètres of oxalic acid is the total quantity 60 cubic centimètres of baryta can neutralise, therefore $63 \text{ C.C.} - 53 \text{ C.C.} = 10 \text{ C.C. of oxalic acid, or 10 milligrammes of carbonic acid.}$

Now refer to the following table, which gives the volumes of 1 milligramme of carbonic acid at different temperatures.

If the temperature had been 17° C. , then—

$$.5403 \times 10 = 5.403 \text{ C.C.}$$

The capacity of the flask is 7,000 cubic centimètres, so the following proportion

$$\begin{array}{r} 7 : 10 :: 5.403 : x \\ \hline 7)54.03 \\ 771 \end{array}$$

gives the volumes of carbonic acid present in 10,000 of air. When several flasks are in use it is better not to mark their actual capacity on the glass, but the divisor and multiplier only, as in the above case, $\frac{1}{7}$.

Volume of 1 milligramme of Carbonic Acid at various temperatures.

° C	Cubic centimètres.	° C.	Cubic centimètres.
At 0 . .	=0.50863	At 19 . .	=0.54402
1 . .	=0.51049	20 . .	=0.54589
2 . .	=0.51235	21 . .	=0.54775
3 . .	=0.51451	22 . .	=0.54961
4 . .	=0.51608	23 . .	=0.55177
5 . .	=0.51794	24 . .	=0.55334
6 . .	=0.51980	25 . .	=0.55520
7 . .	=0.52167	26 . .	=0.55706
8 . .	=0.52353	27 . .	=0.55893
9 . .	=0.52539	28 . .	=0.56079
10 . .	=0.52726	29 . .	=0.56265
11 . .	=0.52912	30 . .	=0.56451
12 . .	=0.53098	31 . .	=0.566
13 . .	=0.53314	32 . .	=0.568
14 . .	=0.53471	33 . .	=0.570
15 . .	=0.53657	34 . .	=0.572
16 . .	=0.53843	35 . .	=0.574
17 . .	=0.54030	36 . .	=0.576
18 . .	=0.54216	37 . .	=0.578

CHAPTER III.

The Means whereby a Constancy in Composition is maintained in the Air—Currents produced by Unequal Temperatures—The Expansion of Gases by Heat—The Law of Charles—Dalton's Observations on the Mixing of Gases notwithstanding their Different Specific Gravities—Graham's Law : the Rate of Diffusion of Gases is Inversely Proportional to the Square Root of their Specific Gravities—Explanation of the Reason of this by the Kinetic Theory of Gases—Effects of Differences of Temperature on Diffusion—Of Ventilation, its Principles and Practice—Its Neglect in most Public Buildings—Evil Effects of Foul Air—The Air Permeating the Soil, or 'Ground-Air'—Passage of the Ground-Air into Dwelling-houses—Danger from Leaky Drainage or Gas-Pipes.

CONSIDERING the number of different gases possessing different properties, and having different relative weights or specific gravities, which being simply mixed together form the atmosphere, and considering that some of these gases, such as hydrochloric acid, carbonic acid, and sulphurous acid, are evolved in excessive quantities in certain districts, such as manufacturing towns, how is it that in a place like Glasgow for instance, the composition of the air is not greatly altered? Why does it not contain an excessively large quantity of all the

noxious gases named above, more especially as these substances are all much heavier than pure air? Why does not the heavy carbonic acid form the lowest stratum on the earth's surface, the oxygen next, and the nitrogen lie over these, while above all rests the ammonia? How is it that Babinet's calculation and Dalton's theory concerning the decrease of oxygen in the air at great elevations have no foundation in fact?

By studying the effect of rise of temperature on gases, one half or more of the problem will be solved. Take a bell-jar, and hold beneath it a roll of brown paper which is smouldering and giving off abundance of smoke; this smoke may be observed to rise and fill the bell-jar, that is to say, the air being heated rises and carries the smoke along with it. Having filled the bell-jar, remove the stopper, and there is immediately an upward rush outwards of heated air, which, by the disappearance of the smoke, indicates that the contents of the glass have been dispersed in the surrounding atmosphere. Again, by means of a charcoal fire one may easily obtain a supply of heated air, which consists for the most part of carbonic acid, a gas at ordinary temperatures heavier than air. If this be allowed to rise into an air balloon, it will so inflate it and render it buoyant that it will be carried to a considerable height before the gas cools. The

effect of rise of temperature is to cause gases to expand and become lighter.

The proportion in which they acquire this property is curiously related to the temperature to which they are heated. Thus on the Centigrade thermometer between the temperature of melting ice and that of boiling water there are 100 equal divisions or degrees. At 273 degrees below the melting ice temperature, we have a point representing a degree of cold which has never been arrived at, but which there is reason for supposing would cease to affect a gas in the same way that it does under all conditions at present known. This point is called the absolute zero, and the following statement is known as Charles' law :

The volume of a gas is directly proportional to its absolute temperature or temperature reckoned from the absolute zero. Thus if a gas measures 1,000 cubic centimètres at 16° C., its volume at 25° C. would be found in the following manner: first calculate the absolute temperature corresponding to 16°; this is $273^{\circ} + 16$ or 289° ; then that at 25° or $273^{\circ} + 25^{\circ} = 298^{\circ}$, after which the following proportion will establish the change in volume which would take place by this rise of 9° C. :

$$289^{\circ} : 298^{\circ} :: 1000 \text{ C.C.} : 1031 \text{ C.C.}$$

In other words, the effect of heat on gases may be

said to cause an expansion of $\frac{1}{273}$ of their bulk for each degree on the Centigrade scale. The effect of the sun on the earth's surface by heating the lower stratum of air causes it to be expanded so that it rises ; at the same time cooler air moves onward to supply its place. When the heating action is exerted on the ocean, evaporation from its surface takes place ; the vapour rises producing a similar effect. The heated air rising, the cooler would rush in a lateral direction, and if with considerable force this would constitute a wind. Winds are nothing more or less than such phenomena, but their direction is determined in certain cases by other causes. If, for example, the air between the tropics, where the sun's rays have greatest power, rushes upwards with great velocity, it is only natural that under so great a disturbance the air at the Poles, which is the coldest, should push forward to take its place. In doing this, however, it changes its entire position on the earth's surface.

Now, as already mentioned, the atmosphere rotates with the earth upon its axis, and the velocity of this rotation is of course smallest at the Poles ; in fact, we might fix on a particular point where it amounts to nothing ; but at the Equator the *surface of the earth* moves with a velocity of 1,000 miles per hour from west to east. If the

wind in travelling northward from the antarctic regions moves at the rate of 12 miles an hour, it would gradually acquire a south-westerly direction, that is to say, moving northward and eastward, as it changed its position on the earth's surface, to points where the rotation from W. to E. was more rapid. Beginning at the South Pole by a motion S. to N. = 12 miles an hour, W. to E. = 0, we should end at the Equator with S. to N. = 12 miles, W. to E. = 1,000 miles an hour, roughly speaking ; a combination of these two directions would give us a wind with more W. than E. in it. As a matter of fact, the trade winds owe their direction to such circumstances. As a wind travelling northward gets nearer to the tropics the more it diverges from its original direction, the increased velocity of rotation of the *earth's surface* being communicated to it. Such are the means whereby the air is thoroughly mixed, though it is not entirely by such violent means that a circulation is effected.¹ It is seldom we experience a day without some wind, and even when the air is apparently quite calm, and not the slightest ripple on the water or the movement of a leaf betrays the motion of the

¹ It is evident that the atmospheric tides must to some extent vary the direction and force of the wind, and if there be any truth in the popular notion that a new moon brings a change of weather, it is evident how the change may possibly be effected.

air, still there are currents ever ascending. Around our bodies, for instance, by means of delicate instruments, it has been proved that the animal heat causes a constant upward streaming of the air. If the smoke of a cigarette be carefully expelled from the mouth, it may be deposited in the centre of a room in a very small space measuring less than a cubic foot. On leaving it to itself it will speedily be dispersed, and if the door be opened it may even be smelt half over a moderately sized house after an interval of a minute. This is in reality due to the movements of the air, which carries the smoke along with it, and not to the motion of the smoke by itself. By bearing this in mind, the extraordinary way in which the so-called still air circulates is readily comprehended.

It has already been pointed out how the reciprocal action of plants and animals preserves a constant balance between the carbonic acid and oxygen of the air. It is necessary now to notice the last cause operating, so as to maintain a constancy in its composition. Dalton, who, as before mentioned, supposed the air of different altitudes to consist of different definite mixtures, completed in 1803 an investigation on the tendency of elastic fluids to diffuse through each other, and his results established this remarkable fact, that a lighter elastic fluid cannot rest upon a heavier, as is the case with

liquids ; but they are constantly active in diffusing themselves through each other until an equal equilibrium is effected ; and that without any regard to their specific gravity, except so far as it accelerates or retards the effect according to circumstances. An effective illustration of this consists in taking two cylinders, one of hydrogen, the other of oxygen. The specific gravities of hydrogen and oxygen are 1 and 16 respectively, that is to say, oxygen weighs 16 times heavier than an equal bulk of hydrogen. Place the heavier oxygen down, and on the top of this the lighter hydrogen, and now remove the glass plates dividing them (fig. 32).



Fig. 32. Interdiffusion of Hydrogen and Oxygen.

In about half a minute one is able to show that these two gases are no longer separated, for by applying a light each cylinder is found to contain an explosive mixture ; if they remained unmixed the hydrogen would burn quietly and the oxygen would increase the combustion of the light applied to it.

Yet another example. In a glass jar filled with air there is placed a piece of red litmus paper. Con-

tact with ammonia would render this blue. As the specific gravity of ammonia gas is little more than half that of air: if a porous vessel containing a supply of ammonia gas be raised upon a stand within the jar, the gas will have to make its way downwards to the paper, by virtue of this law of diffusion or intermiscibility as it might be termed. After a few minutes the red litmus is turned blue by the downward diffusion of the light ammonia



Fig. 33. Downward Diffusion of Ammonia.

gas in the smaller jar (fig. 33). Thus is shown the truth of Dalton's observations. Graham, who submitted these phenomena to a long and close investigation, discovered that the rate of diffusion of two gases is in the inverse proportion of the square roots of their relative weights or specific gravities; as, for example, in the case of oxygen and hydrogen, with specific gravities 16 and 1 respectively, as the rate of diffusion is in *inverse ratio* to the specific gravities, the gas with the higher number will have the lower rate of diffusion, but the proportion is as the

square root; now the square root of 16 is 4, so that we arrive at the conclusion that hydrogen will diffuse four times quicker than oxygen. Or if the time of diffusion be equal, four times as much hydrogen will diffuse out of a vessel as oxygen will diffuse in.¹ To establish this fact by experiment, a tube is filled with hydrogen. In order to prevent the hydrogen escaping merely by its lightness, the upper end of the tube is closed with a plug



Fig. 34. Upward Diffusion of Hydrogen.

of plaster of Paris. Now immerse the lower end in water, and leave the plaster plug exposed to the air: the hydrogen will presently diffuse outwards more rapidly than air can enter, consequently the liquid is drawn up into the tube (fig. 34). In order to be satisfied that the lightness of the hydrogen is not the cause of its

escape, the same experiment is arranged in a different way (fig. 35). A bottle with a plaster bottom is filled with the gas, and a curved tube proceeding from the neck dips downwards into water. By taking

¹ Rate of diffusion of hydrogen } : { Rate of diffusion of oxygen } :: $\sqrt{16} : \sqrt{1}$.

off the air-tight india-rubber cap which covers the

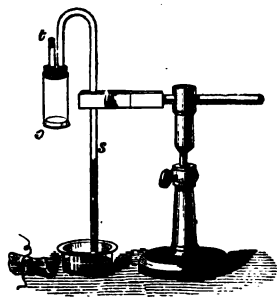


Fig. 35. Downward Diffusion of Hydrogen.

plaster bottom you have evidence that the hydrogen is escaping just as rapidly as before, by reason of the water ascending the tube. Here lightness of the gas cannot have influence, as the escape is downwards instead of upwards. Yet another example of the same phenomenon may be arranged (fig. 36). On a tube 5 feet long is

fastened, by means of an india-rubber ring, a porous

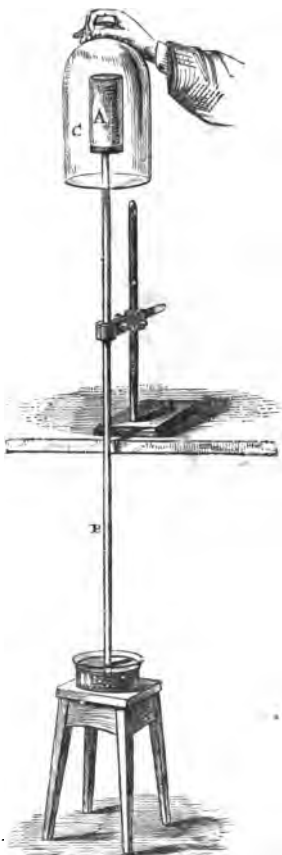


Fig. 36. Inward and Outward Diffusion of Hydrogen.

cylinder : the lower end of the tube dips under water. By bringing a jar full of hydrogen just over the cylinder, it will be observed that air is driven out of the apparatus, and through the water ; by removing the bell-jar the water rises in the tube, showing that air escapes by the porous pot. In the first case the

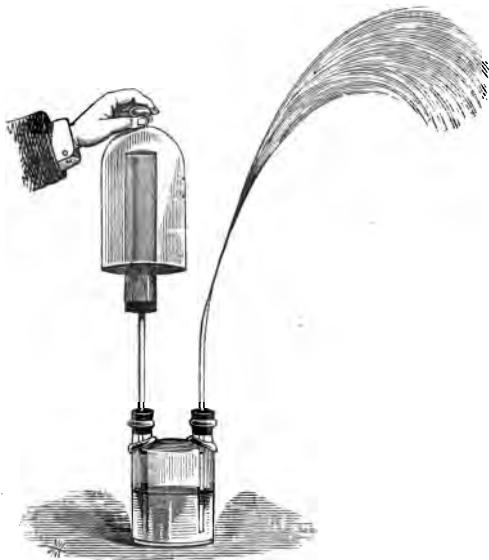


Fig. 37. A Diffusion Fountain.

hydrogen diffused into the porous vessel so much more rapidly than the air diffused out, that a pressure was created within which expelled some of the gas through the water ; but on removing the atmosphere of hydrogen, this gas diffuses so much

more quickly from the cylinder into the surrounding air than the air can take its place, that the water is drawn up the tube. That the pressure exerted is pretty considerable can be shown, but by another form of apparatus in which the long tube passes into a two-necked bottle (fig. 37), and in the other neck of the bottle is fitted a glass jet dipping below some water contained within, on holding the hydrogen jar over the cylinder a jet of water two feet in height is projected upwards, and then on letting air have access to the cylinder the hydrogen diffuses outwards, the fountain ceases to play, and air rushes violently through the liquid in the bottle to take its place in the porous cell.

The cause of these extraordinary phenomena is generally explained now by reference to molecular motion, or a constant movement or commotion amongst the particles of the gas. The kinetic or motion theory has been elaborated by Professors Clausius of Bonn and Clerk Maxwell of Cambridge. We have the best reasons for believing that equal volumes of gases at the same temperature and pressure contain the same number of molecules or particles, which molecules or particles are all of the same size. They exert their pressure in all directions, upwards as well as downwards, and thus it is that the pressure of the atmosphere is not felt by us under ordinary circumstances. If, then, a gas be

under a pressure of 15 lbs. on the square inch this pressure is exerted by the gas in every direction, perpendicularly and laterally.

Two pint bottles, filled the one with hydrogen the other with oxygen, are standing inverted, with their mouths immersed in water; the oxygen weighs 16 times as much as the hydrogen, yet the pressure in both is the same. If the particles of the gases be in a constant commotion and a hydrogen molecule strikes the vessel with only $\frac{1}{16}$ the force of the oxygen, the pressure must be equalised by many more blows being given in the same space of time, for the force with which a projectile strikes an obstacle is ascertained by multiplying its weight by the velocity with which it travels. Thus a weight of 16 lbs. travelling 1 foot per second would deliver a blow with a force of 16 lbs.; and a weight of 1 lb. at a speed of 16 feet per second would equalise this. But it must be remembered that if each hydrogen molecule hit the sides of the vessel with 16 times the velocity of the oxygen, or what is the same thing, if each hydrogen particle takes only $\frac{1}{16}$ the time of the oxygen to deliver its blow, it follows that 16 particles of hydrogen would strike in the time required by one particle of oxygen.

The pressures under such circumstances might erroneously be expressed as follows:—

	Weight		Velocity		Strokes in same time		Pressure
Oxygen .	16	x	1	x	1	=	16
Hydrogen .	1	x	16	x	16	=	16 x 16

or the mechanical power expended by the hydrogen under these circumstances would be 16 times greater than that of the oxygen. If, however, we reduce the velocity of the hydrogen molecules to 4 times that of the oxygen, then we find—

	Weight		Velocity		Strokes in same time		Pressure
Oxygen .	16	x	1	x	1	=	16
Hydrogen .	1	x	4	x	4	=	16

This is a rational explanation of Graham's law, 'The rates of diffusion of gases are inversely proportional to the square roots of their relative weights.'

Feddersen of Leipzig has made some interesting experiments on diffusion when influenced by temperature. His method of experimenting was as follows:—A porous stopper was inserted in a glass tube. The tube was placed horizontally, and the two ends projecting beyond the stopper were connected air-tight, by means of caoutchouc, with other glass tubes, also placed horizontally, in each of which a drop of liquid formed a moveable stopper. Thus every propulsive movement of the column of air in the middle glass tube would cause a movement of the drops in the same direction. One end of the stopper was subjected to constant source of heat, while the other remained cold or

was artificially cooled : whereupon a slow passage of the column of air through the stopper was uniformly observed having the direction from the cold to the warm end of the stopper. From the results obtained it appears to be a common property of porous substances, when acting as a diaphragm in gas, to occasion a passage of gas from the cold to the warm side. This phenomenon of diffusion, unlike the ordinary diffusion, takes place when on both sides the porous partition there is the same gas with the same pressure ; it is proposed to give it the name of *thermo-diffusion*.

Having thus given an account of the manner in which the atmosphere is variously affected, it will now be better to proceed to the consideration of the air of close places, the composition of this under various conditions, and the possibility of keeping it similar to that of fresh air. It has already been shown that the atmosphere contains even in the open country 3 volumes of carbonic acid in 10,000 of air, that this amounts to as much as 4 or even 5 volumes under ordinary circumstances in towns. It is not advisable to breathe air containing more than 6 volumes of carbonic acid per 10,000, and, indeed, such air is pronounced by those who enter it from the outside as close, that is to say, it is disagreeable—its impurity is detected by the nose—our senses warn us

not to breathe it. We cannot, therefore, conveniently increase the amount of carbonic acid in buildings by more than 2 volumes per 10,000 over that of the external air. If the amount be increased to 250 volumes the air would cease to support combustion, and 300 volumes would speedily extinguish human life, notwithstanding that 17 to 18½ per cent. of oxygen would be present, that is, 1,700 or 1,800 volumes of oxygen in 10,000. The



Fig. 38. An Experiment with Respired Air.

amount of carbonic acid in the breath is about 5 per cent., so that air once breathed is useless and poisonous so far as mankind is concerned. This may be easily shown. A bell-jar (A, fig. 38), sunk an inch or two in water, but filled with air, has a cork and flexible tube B, so arranged that one may

replenish the lungs from this supply after they have been purposely emptied. After again expelling the air from the lungs into the apparatus, if a lighted taper be introduced it is instantly extinguished, and on again taking this air into the lungs a suffocating sensation will be felt. A little lighted phosphorus, however, continues to burn in the confined space, showing that all the oxygen is not used up in the act of respiration. The amount of carbonic acid given off by a man in an hour is six to seven-tenths of a cubic foot, or 14 to 16 cubic feet in the twenty-four hours. Two sperm or paraffin candles have the same effect as one man, that is to say, a good candle yields three-tenths of a cubic foot of carbonic acid in an hour. One cubic foot of gas produces two cubic feet per hour of carbonic acid, and a good gas-burner will consume three cubic feet of gas in that time. A good oil lamp gives a little more than half a cubic foot of carbonic gas per hour. Dr. Odling has ascertained that for equal illuminating power candles yield a larger amount of impurity to the air than gas. From these data it is readily calculated¹ that a man seated in a

¹ The carbonic acid from the lamp and the man together amounts to $1\frac{1}{2}$ cubic foot, the excess of this gas over that in atmospheric air which is allowable is 1 in 5,000 cubic feet; if therefore we multiply $1\frac{1}{2}$ by 5,000 we arrive at the amount of air required to dilute the products of respiration and combustion.

room with two lighted candles or a good oil lamp would require the use of 6,000 cubic feet of air in an hour in order to keep the atmosphere from becoming uncomfortably vitiated, that is, prevent the carbonic acid from rising above 6 volumes in 10,000.

Generally speaking, in dwelling-houses this is accomplished by having lofty and spacious rooms, or, in other words, having a large amount of pure air to commence with.

Another cause leading to the same end is the discomfort produced by heat, necessitating an adjustment of doors and windows; this applies more particularly to apartments of smaller dimensions. According to Angus Smith, from experiments made in a leaden chamber, the rise in temperature of 170 cubic feet of air in one hour—

By the bodily heat of one man = $5^{\circ}6$ F.

By the combustion of a candle = $3^{\circ}8$ F.

Accordingly, in a room 8 feet high, 6 feet long, and 4 feet broad, a man burning two candles would in an hour raise the temperature from 60° F. to 73° F., which alone would be sufficient to produce uneasiness, which would be aggravated by the physiological action of the carbonic acid present. He would in consequence of the uncomfortable feeling open a door or window to get rid of the heat and thus supply himself with fresh air.

But the heating effect of the body and of a candle is much greater than is here indicated. Such an experiment is imperfect because it must necessarily last for a considerable time, and in the interval the metallic walls are parting with the heat they absorb from the contained air.

The amount of heat given off by the combustion of a candle has been accurately determined by other processes, and it has been found that a good candle, producing 0.31 cubic foot of carbonic acid per hour, will raise 150 cubic feet of air $118^{\circ}.4$ F. The bodily heat of a man, calculated from the amount of carbon consumed internally, which is indicated by the carbonic acid in the breath, will effect a rise of $139^{\circ}.6$ F. in the temperature of 150 cubic feet of air. This latter number is no doubt too small, because there is omitted the amount of heat resulting from the oxidation of hydrogen to water in the body, a process which must take place to some extent in the destruction of the tissues.

In default of more accurate data concerning this matter the following calculation is of interest: If the heating effect of a candle per hour on 170 cubic feet of air as found by experiment in the leaden chamber is $3^{\circ}.8$ F., and that of a man's body $5^{\circ}.6$ F., in a small quantity of air these numbers would be greater, and calculation shows that for

150 cubic feet they would be $4^{\circ}3$ F. and $6^{\circ}3$ F. respectively.

But the actual heating effect of a candle amounts to $118^{\circ}4$ F.—hence $\frac{118^{\circ}4}{4^{\circ}3} = 27.5$, or the quantity of heat observed in the leaden chamber must be multiplied by 27.5 to obtain the actual quantity. As the heating effect of a man's body is not widely different from this we may apply the same calculation, so that $6^{\circ}3 \times 27.5 = 173^{\circ}2$. A number somewhat higher, and probably more correct than that deduced from the carbonic acid in the breath, namely, $139^{\circ}6$ F. These figures are instructive, showing as they do the great heating effect of a mass of human beings in a close apartment where there is an insufficient circulation of air around them to carry off the heat with sufficient rapidity.

Under ordinary circumstances the rise of temperature alone is a self-acting medium of ventilation, as already shown. It causes, firstly, a diffusion of cold air into a warmer space, and secondly, an ascension of a heavy gas like carbonic acid to the upper part of a building. A very good illustration of this latter fact is afforded by the following experiment:—Two jars, holding each about half a gallon of carbonic acid, are tested with a lighted taper, which each is found

to extinguish. Into each jar is lowered a little flask, one containing hot water, the other cold (fig. 39); after a few minutes the jar with the flask of cold water will be found to contain sufficient carbonic acid to extinguish a taper, while

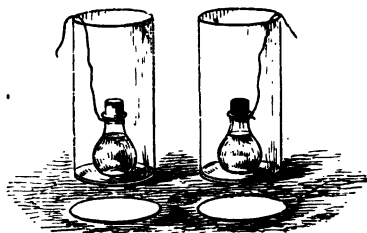


Fig. 39. The Ascension of Warm Carbonic Acid.

the air of the other jar will sustain combustion brilliantly.

Furthermore, to perform the experiment on a larger scale, take a glass vessel holding ten gallons of the gas; its presence within is made evident by the floating of soap-bubbles on the surface. Now introduce a quart flask containing warm water, and after a short interval it will be shown that the carbonic acid has escaped by the fact of the soap-bubbles no longer floating, but sinking to the bottom of the vessel.

All the foulest air of a room is near the ceiling; in fact, it is so bad there that unless an easy outlet be provided it becomes perfectly poisonous. A

few simple experiments will illustrate the principles of ventilation in buildings and mines in a more convincing way than mere description.

In a tall bell-glass (fig. 40) is arranged a tree of lighted candles ; the carbonic from all of them rises, the consequence being that the top one goes out first, the second next, and so on in succession, as the gas accumulates and fills up the space from the top downwards ; but if before the last and

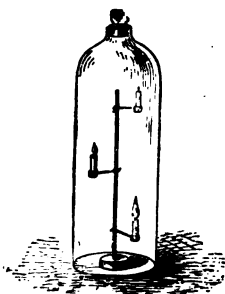


Fig. 40.

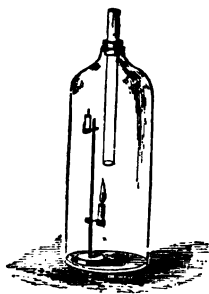


Fig. 41.

lowest is extinguished the stopper be withdrawn and the jar raised half an inch, the heated vitiated air escapes, fresh air enters, perfect ventilation results, and the light continues to burn brightly.

A similar jar has a glass chimney fitted into the neck (fig. 41) ; one taper is raised on the stand to the upper part of the jar, and one is placed below the chimney, and the lower rim of the jar

is raised a little from the taper. The upper candle will go out, while the lower one gets a fresh supply of air from below, and discharging its products of combustion by the chimney continues to burn in the usual manner. In mines, where fresh air has to be carried downwards, the burning of fires is resorted to as a means of introducing fresh air. The principle can be well shown in the following manner:—A small box has a glass chimney

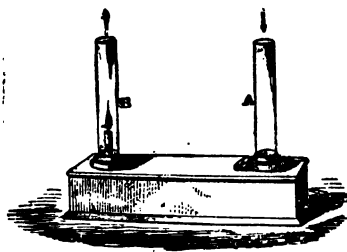


Fig. 42.

fitted into each end (fig. 42). The one distinguished by the letter A has a taper suspended within it; this continues to burn brightly, the entrance of fresh air being downwards by the shaft or chimney B. A piece of smoking brown paper shows this: the smoke pours down the one shaft and passing the candle rises out of the other. Now by closing B with a small glass plate the entrance of fresh air downwards is prevented, and the candle is extinguished. On repeating the experiment so far, and

just before the extinction of the taper, if the glass plate be removed the taper is restored to life again. But in mines it is not often convenient to have two shafts, and ventilation is accomplished by dividing one by a partition. Thus in a bell-jar closed below by water, if there be fitted a glass chimney a candle will burn until the air is so vitiated that it is just about to die out; now on inserting a piece of metal so as to divide the chimney into two halves the light is revived (fig. 43), and the smoke from brown paper will give evidence of there being two draughts, one in an opposite direction to the other. The smoke descends one side of the partition, curls round in the jar, and makes its exit by the other side. The causes determining which shall be the upcast and which the downcast shaft in this case are so small, the candle being as nearly as possible between the two, that we cannot detect them. In rooms the chimney is always a most important ventilating shaft, while accidental openings in doors and windows, mere chinks or cracks as they may be, yield a large supply of fresh air. There is, however, a source of ventilation quite as efficient as the opening of

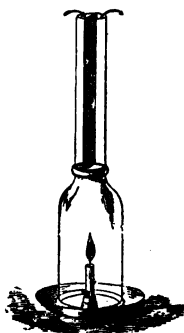


Fig. 43.

doors and windows, or the accidental openings that occur in buildings; it is a self-regulating supply of fresh air provided by diffusion through the porous walls. We are unconscious of its operation, and though the amount of air supplied by passage through bricks or stone, even when plastered and papered, is very considerable, yet one cannot detect it by any sensation of draught.

Here is the value of this means of ventilation. In winter, when it would be uncomfortable to open windows, but when, having doors and windows closed, and lamp or gas-light in our rooms, there is really more urgent need of fresh air, then this unfelt and unsuspected supply is greatest. An account of an experiment of Pettenkofer's will give the best evidence of the value of penetration through the walls, and the influence of temperature on this diffusion. An ordinary room in his house with a German stove instead of an open English fire-place, the walls being of brick, and the capacity being 2,650 cubic feet, had its entire contents changed once in an hour, when the difference between the inside and outside temperature was 34° F. (66° inside, 32° outside), the doors and windows being shut. With the same difference of temperature, but with the addition of a good fire in the stove, the change of air rose to 3,320 cubic feet per hour. Now here is the most important fact, when

all openings and crevices in doors and windows were pasted up and rendered thoroughly air-tight, there was still a change amounting to 1,060 cubic feet per hour, which was owing to diffusion through the walls. On lessening the difference between the internal and external temperature to only 7° F. (71° inside, 64° out), the change of air was reduced to only 780 cubic feet per hour, but on opening a window 8 feet square the change rose to 1,060 cubic feet per hour. These numbers are most interesting, showing as they do that houses may be better ventilated in winter with doors and windows shut than in summer with windows wide open, because it is here proved that a difference of temperature of 34° with all openings carefully closed is of greater influence, effecting a more complete change of air than large and uninterrupted communications with the outer air at a small difference of temperature.¹ Hence a crowded ball-room in summer with windows wide open has a more unwholesome atmosphere than one would at first be led to suppose. With a roaring fire the room was ventilated by the draught of the stove to the extent of only 700 cubic feet. The free wall of a room on being examined for its ventilating power gave the fol-

¹ It is the custom in some parts of Germany for the plastered walls to be coated thinly with paint instead of paper. It is doubtful which material was used in the case of this room.

lowing numbers :—The room contained 2,650 cubic feet, and with a difference between external and internal temperatures of 4° F., the spontaneous ventilation or diffusion through each square yard amounted to no less than 7 cubic feet or 43 gallons per hour. It has been found by Märker and Schultze that the spontaneous passage of air through 1 square yard of wall, the difference of temperature being 4° F., varies with different materials according to the following table :—

	Per hour.
Sandstone	4.7 cubic feet
Quarried limestone	6.5 „
Brick	7.9 „
Fufaceous limestone	10.1 „
Mud	15.4 „

The extent of free wall surface is therefore an important matter as regards the ventilation of a house ; the larger the proportion of external wall surface to the number of inhabitants the better will be the ventilation. The inhabitants of the lofty houses in the splendidly proportioned streets of Edinburgh and Paris, where families live on floors or flats, cannot have the advantage of so large a supply of fresh air through the walls in proportion to their numbers as those who dwell in the small suburban villas around London. No doubt the much lighter style of building has much to do with the matter, but setting this aside, the circulation of the wind

round a house has a great effect on its ventilation, no doubt, in causing a transpiration through the walls, whether through the porous materials or through chinks and crevices, and such circulation is reduced to a minimum if the houses are many-storied and constitute fine regular streets.

If dwelling houses are not over-crowded, there is no occasion to resort to artificial ventilation. The natural ventilation caused by difference of temperature, facilitating the passage of air through the dry and porous walls, together with the assistance of windows and doors, should be sufficient, provided the greatest cleanliness throughout, and avoidance of pollution of the air is maintained. It is a curious fact, as Pettenkofer has pointed out, that the more impervious a building stone is to air the larger is the quantity of mortar generally used with it. Mortar which is made of sand and lime is, when dry, exceedingly porous, we therefore get a compensation in the larger quantity used for the lesser porosity of the stone used with it. Thus sandstones are generally so porous that water and air readily permeate them. Solid or quarried limestones are most impervious to air; but as they are very irregular in shape, they require more mortar, and such is the reason why a wall built of such material does not differ much from a wall made of regular bricks and thin layers of mortar.

Observations have shown that a wall of quarried limestone would consist of one-third mortar, tufaceous limestone one-fourth, bricks one-fifth to one-sixth, sandstone one-sixth to one-eighth.

There is a style of building now coming into vogue which consists of putting up an iron framework, which is filled with concrete; when this material has hardened, the iron supports are taken away, and a very durable wall remains. Practical experience has yet to prove whether this material is sufficiently porous to constitute the walls of healthy houses. Pettenkofer mentions the fact that a workman's house was built near some iron-works, from large regularly shaped pieces of slag, which may best be described as an opaque glassy substance, not in the slightest degree porous. When completed the house dried rapidly, but when the family had lived in it a short time the damp became excessive and remained, thus making this the worst house on the establishment. Moisture condensed on the walls just as it does on the glass of a greenhouse. To keep a house dry the walls should be thick and porous, and the rooms well warmed. When damp shows itself on the walls of a house after a long spell of wet weather, it is no indication of wet having come in from outside, but it is due to a condition of things similar to that causing the house built of slag to become damp

after it had been inhabited. The walls, when thoroughly wetted, are no longer porous ; they are also better conductors of heat, and are consequently constantly parting with the warmth derived from the room to the outside air, which process of course keeps them cold. The aqueous vapour with which the warm air of the house is saturated is thereby condensed on the walls, and makes itself evident to the eye.

A remarkable case in a London house has come to my knowledge, which gives a distinct proof of the much greater passage of gas through the walls in winter than summer. A small room occasionally used was noticed sometimes to have an unbearably bad smell ; this was never noticed in summer nor in winter unless a fire was lighted in the room ; the drainage was suspected and examined, but was found perfect, yet here was this extraordinarily foul air making its way into the room whenever the interior was warm and the exterior cold. The cause was a dust-bin built against one of the walls, and the filtration of the air through this and the house wall into the room.

Having thus mentioned the causes constantly in operation which effect a change of air within our dwellings, it is necessary to lay emphasis on the fact that the reckless opening of windows does not constitute good ventilation, though it may result in

the supply of fresh air. Those motions of the air which are so gentle as to be unperceived by us, causing it to stream upwards from our bodies and through the walls of houses, must be imitated to secure good ventilation; that is, the supply of fresh air *must be free from draught*. If the rate of an almost imperceptible breeze be measured, it is sure to be found travelling at a higher speed than 19 inches per second. If the hand be passed through the air so that a space of 18 inches may be traversed in a second, about equal to 1 mile an hour, no draught or current of air will be felt. The effect is the same if the hand be still and the air moving at this rate. Everyone knows what a draught is from having experienced the discomfort, if nothing worse, of the chilling sensation caused by its continued action; good ventilation does not admit of draughts, generally speaking. Ventilation in temperate climates is the art of supplying fresh air at a rate not greater than 18 or 19 inches per second, and in sufficient quantity to reduce the proportion of carbonic acid present to 6 volumes in 10,000 of air. In India, where the external temperature is so high that punkahs to create a draught are a necessity, an essential point in ventilation is the supply of air at a much greater speed than in this country would be comfortable.

A person does not inhale and exhale more than

18 cubic feet of air per hour, but the supply of fresh air necessary, according to General Morin's experiments made in Paris, is not less than 2,120 cubic feet per hour. I have already made a calculation which will answer either for two men or a man and two candles, showing that in a small apartment of 170 cubic feet capacity 6,000 cubic feet of fresh air would be required, or 3,000 cubic feet per head. This is the amount fixed on by Drs. Parkes and De Chaumont. In addition to this every ordinary gas-burner consuming 3 feet of gas per hour necessitates a supply of 5,400 cubic feet of fresh air. Generally speaking, in this country, the change of air in a room cannot be effected oftener than 3 or 4 times in the course of an hour without a sensation of draught; hence, as 3,000 feet of air are required for each individual, a space of 750 to 1,000 cubic feet per head must be provided in an apartment. The following is a practical example of the kind of calculation necessary to form an opinion as to the ventilation of an apartment:—A bed-room $12\frac{1}{2}$ feet square and 10 feet high had an air capacity, after deducting for the space occupied by furniture, of 1,500 cubic feet; while one wall communicated with a dressing-room of 560 cubic feet capacity, two of the other walls were against the open air; there was free communication between the room and the chimney, but both

windows and door were shut. The room had one occupant for 9 hours, and during that time the carbonic acid had increased only 1 volume in 10,000 of air. Taking the carbonic acid expired during the night as 5·4 cubic feet, the air which had passed through the room was 54,000 cubic feet ; that is to say, the entire contents of the room had been changed four times in every hour without the help of open windows.

The ventilation of ordinary dwellings is a simple matter ; but in considering the case of public buildings and workshops, where the numbers of people crowded together is very great, and the amount of gas consumed extremely large, there is under such circumstances a state of things almost beyond the control of natural ventilation. Fresh air should be provided by well devised means, and the greatest possible care should be taken that natural ventilation should have full play ; but how seldom this is the case the following tables will show :

*Carbonic Acid in close places in London, according to
Dr. Angus Smith :—*

	1864	Parts in 10,000 of Air
Chancery Court, 7 ft. from ground, closed doors,		
March 3		19·3
Same, 3 ft. from ground		20·3
„ door wide open		5·
Strand Theatre, gallery, 10 P.M.		10·1
Surrey Theatre, boxes, „ „		11·1
„ „ 12 „		21·8

1864	Parts in 10,000 of Air
Olympic, 11.30 P.M.	8.17
„ 11.55 „	10.14
Victoria Theatre, boxes, 10 P.M.	12.6
Haymarket, dress circle, 11.30 P.M.	7.57
Victoria Theatre, boxes, April 4	7.6
City of London Theatre, pit, 11.15 P.M.	25.2
Standard Theatre, pit, 11 P.M.	32.0
St. Thomas's Hospital, Queen's ward, 3.25 P.M.	4.0
„ „ Edward's ward, 3.30 P.M.	5.2
Lambeth Workhouse wards	1
St. Luke's (Chelsea)	7.6
East London (Homerton)	7.6

In April 1874 I made experiments to ascertain the amount of carbonic acid in the air in Covent Garden and Drury Lane theatres during the performances of Italian opera, and also in the Money Order Office in Aldersgate Street. In the theatres it was considered that from the amount of gas burnt and the small difference between the outside and inside temperature, besides the fact that no extensive provision for ventilation being in use, that the air would be very bad, much worse than in winter, when the outside temperature is low. In the theatres the escape of burnt gas is not sufficiently provided for, and the products of combustion strike the roof and roll over into the galleries. This arises not so much from want of outlet as for lack of inlet of fresh air to displace that rendered hurtful by heat and impurities. Public buildings too often resemble the

bell-jar on page 111, fig. 43, in which, though a candle is burning beneath an ample opening, the vitiated air does not escape for want of sufficient inward draught. But as soon as the chimney is divided, upward and downward currents are established and ventilation ensues. If shafts were carried downwards from the roof to the basement of a theatre there can be no doubt similar effects would follow. The gallery and boxes, too, of a theatre may be likened to a number of small and over-crowded apartments, being as they are so shut in overhead and at the back. That this is the case anyone may prove for himself by passing the hand up from floor to ceiling when, between the acts, the door is open on a warm night. The cold air rushes in at the bottom, and at the top the hot air escapes; two currents are playing in opposite directions, one over the other.

Analyses made in April 1874.

	Volumes of Carbonic Acid in 10,000 of Air
Covent Garden Theatre, amphitheatre	22.0
„ „ near an open door, time 10.10	17.6
„ „ near a ventilator, time 10.20	14.8
Drury Lane—in the lower gallery	27.0
Money Order Office	17.4

Here it is easily seen that the air in the theatres is very bad; but after the doors had been open for a short time between the acts it rapidly improved;

indeed, in Covent Garden in the second case, near an open door, the people exclaimed, by force of contrast, how delightful the fresh air was ; nevertheless, this *fresh air* contained 14·8 parts of carbonic acid in 10,000, or from $2\frac{1}{2}$ to 3 times as much as it should have had. Drury Lane was the first place experimented on, and having entered at the commencement of the performance, the bad effect of the air as it became vitiated was only gradually experienced, but it produced a listlessness and headache. All the audience around were evidently affected in the same manner, and appeared to be constantly sighing and gaping, or in other words gasping for breath. In the upper boxes of the same house, a rather more confined space, it seemed as if the breath of those in front was passed on from mouth to mouth, the last recipients being those seated farthest behind. On entering Covent Garden from the street in the middle of the performance it was scarcely possible to remain in the building for more than twenty minutes, and the feeling of relief on coming out again was immense. The effect experienced by many persons after breathing such foul air for three or four hours results in a feverish state and headache, even after so long an interval as twelve or twenty-four hours. The immediate effect is to weaken and decrease the rate of the pulse, and at the same time quicken the

respiration. There can be no question that the effect of the performance and the enjoyment of the audience are very seriously damaged by the listlessness and dulness of comprehension and feeling induced—a condition which is only occasionally dispelled for a time by some increased exertion on the part of the performers.

The bad effect of carbonic acid in the air alone, without taking into account organic matter, has been shown by Dr. Angus Smith, who ascertained that 1 part of the gas in 1,000 of air produced in 15 minutes an increase in the number of respirations from 18 to 19 per minute, which increase remained the same up to 45 minutes; the pulse was lowered in 25 minutes from 73 to 71 beats, while at 45 minutes it was 72 per minute. With $2\frac{1}{2}$ volumes of carbonic acid in 1,000 of air the pulse, at first 70 beats per minute, increased to 73 at the end of 10 minutes, and at the end of half an hour was lowered to 69, while the inspiration increased from 17 to 21 per minute. With 5 volumes of carbonic acid per 1,000, the pulse at first 76, and the inspiration at 17, at the end of 40 minutes were represented by the numbers 71 and 24.

Passing over extreme cases of poisoning by foul air, such as the well-known one of the Black Hole of Calcutta, and of the prison in which 260 out of 300 Austrian prisoners died after the battle

of Austerlitz, it is necessary to lay great stress upon the now well-ascertained fact that air only moderately vitiated, if breathed for a long time day after day, produces most serious results. The effects are seen in pale faces, partial loss of appetite, a lowering of the spirits, and a decrease of muscular strength. Very frequently such affections of the respiratory organs as bronchitis and phthisis result. Dr. Parkes, in his work on 'Practical Hygiene,' mentions the fact that in the badly ventilated prison of Leopoldstadt in Vienna, in the years 1834 to 1847, the proportion of deaths was 86 per 1,000, out of which number 51·4 per 1,000 were due to phthisis, while in the well ventilated House of Correction of the same city the deaths were 14 per 1,000, of which 7·9 were occasioned by phthisis; 43·5 cases of phthisis may be therefore attributed to foul air. The origin of phthisis and other lung diseases may reasonably be considered due to the inhalation of organic particles thrown off from the lungs of unhealthy persons, and it is now a well ascertained fact that a bad atmosphere assists the rapid spread of several specific diseases, such as typhus, small-pox, and scarlet fever. From personal observations convincing evidence was obtained that such was the case with the cattle plague in 1866. In cattle sheds containing twenty or thirty cows—which the owners kept closed to such an

extent that all chinks in the doors and windows were stuffed with straw and matting, under an ignorant belief that thus the plague could be kept out—very frequently the entire stock died in two or three days after the first appearance of disease ; while in other cases, where the animals were housed in a well-cleaned and tidily-kept shed, with a plentiful supply of fresh air, not only did some of them escape the disease altogether, but the deaths were reduced to one-third of the number of beasts attacked.

The large supply of air necessary in hospitals for contagious diseases is fully recognised by medical men, and more especially so in America.

The next matter—of greater importance than it may at first sight appear—is the composition of the air permeating the soil. Boussingault and Lewy, in 1853, first made researches in this direction, and found that this air is very rich in carbonic acid, no matter whether the soil be sandy, loamy, or rich in organic matter. The amount of carbonic acid is much greater than that of the atmosphere, though of course earth, with a large proportion of decomposing organic matter in its composition, has a much greater proportion of this gas. The following table illustrates this :—

the carbonic acid in the air, may be compared with that in the second flask produced by the carbonic acid in the soil. In fact, after a very short time, it is made perfectly evident that earth dug out of a meadow renders to the air passing through it at least ten times as much carbonic acid as the air originally contained. Pettenkofer has thought it advisable to investigate the nature of the ground air twice a week all the year round in Munich, and he considers it a matter of much importance and worthy of great attention. His ascertained facts are surprising, and a much larger number of experiments is necessary before we can account for them. In sterile land, consisting of chalky rubble, he finds the quantity of carbonic acid is smaller at 5 feet than 13 feet throughout the year, excepting in the months of June and July, when an inverse proportion arises. In the lower stratum, however, a considerable increase soon begins again, which continues until the maximum is attained. The carbonic acid in the soil of Dresden is far greater in amount—nearly twice as much as in that of Munich. We cannot connect these facts with any other knowledge which would lead to their explanation, but they teach us an important lesson concerning the surroundings of our houses. The soil is very easily permeated by the outer atmosphere—more especially if some

pressure, such as that of the wind, is brought to bear upon it. That this is the case may be shown thus :—A long cylinder is filled with earth, to the bottom of which goes a tube. At the upper end the tube is bent like a U, and a little water is placed in the bend ; now, by blowing with the mouth through a tube on to the surface of the soil, the liquid is seen to be raised, which indicates that a measureable amount of pressure was exerted. It has been shown that the bulk of a gravelly soil consists of one-third air ; that is, the space between the stones and particles of sand which is filled with air amounts to one-half the space filled by the gravel. Let us now consider this soil to be the surroundings of a house, and imagine a drain leaking into such soil ; the ground air would be charged with the products of decomposing offensive matter, and those abominations which are to be so much dreaded as the cause of zymotic diseases. The house built upon this soil is full of warm air, and up the chimneys a draught rushes, causing, when the doors and windows are closed in the lowest rooms, a certain amount of the ground air to enter by way of the floor, carrying with it the germs of disease and perhaps death. Pettenkofer mentions the fact that persons were poisoned and killed by coal gas in a house where no gas was laid on. It had travelled twenty feet under the

street, and then through the foundations, cellars, and flooring of the lower rooms. A most serious explosion, which demolished at least one house, occurred in the latter part of the year 1874 at Southgate, near London. The houses had no gas laid on, but the main pipe which ran along the road was damaged, the coal gas permeated the soil, and escaped into the dwellings, where it was smelt for some days before the final disaster happened. These cases both occurred in winter, when the house being warmed acted as a chimney shaft. It shows that a badly made drain and ill-fitted gas-pipes may be the death not only of the inmates of the house to which they belong, but also of the neighbours who have no control over them. These remarks do not apply to buildings with properly ventilated cellarage, but they show the importance of placing ventilating gratings in the wall between the soil and the flooring of basement rooms. In the suburbs of London it is not at all uncommon for speculative builders to erect houses the drain-pipes of which have no proper outlet, so that until they become filled with sewage, and the surrounding soil saturated—which may take about twelve months—the unfortunate inhabitants are unaware of the fact. In the meantime the mischief so much to be dreaded may have occurred. Against such practices there is no protection at present, or

what is the same thing, there is nothing to prevent such reckless dishonesty. It is a question how far a clay soil is or is not impervious to sewer gases, and therefore to what extent the air of houses built on clay may be free from contamination with impure ground air; but enough has been said to show how dangerous is defective drainage, and how easy is the escape of sewer gases through the soil into dwelling-houses, the earth actually forming a reservoir for them. It was pointed out some years ago that evil effects arose from living in houses built on sites originally gravelly, but from which the gravel had been removed and replaced by the refuse from dust-bins. From complaints recently made this practice seems to be still carried out. It is unpleasant to think of living on a dust-heap, and the preceding remarks have made the danger of it obvious.

CHAPTER IV.

Erroneous Ideas concerning the Origin of Minute Organisms—**Redi's** Experiments on the so-called Spontaneous Generation in 1638—**Needham's** first Experiments in Closed Vessels in 1745—**Ap-pert's** Food-Preserving Process—The Researches of **Schwann**, of **Schultze**, and of **Schroeder** and **Dusch** on Atmospheric Germs—**Pouchet's** Theory of Heterogenesis—**Pasteur's** Researches on the Solid Particles Contained in the Air—Germination of the Dust Suspended in the Air—Experiments on the Air of Different Localities—Later Evidence Against the Evolution of Life from Lifeless Matter.

THE facts connected with putrefaction and decay have from superficial observation led to the erroneous belief which, taking its origin in the earliest times, has survived to some extent until the present day—that minute organisms, the lower order of living things, spring from dead matter; that worms, for example, are derived from putrefying flesh, that they are not reproduced from parent organisms, but are generated from decaying substances. A few quotations will illustrate this.

Lord Bacon, in the 'Novum Organum' (Book II., Aphorism 13), says: 'All putrefaction exhibits some slight degree of heat, though not enough to

be perceptible to the touch : for neither the substances which by putrefaction are converted into animalculæ, as flesh and cheese, nor rotten wood which shines in the dark, are warm to the touch.' The joke of Dr. Johnson on Tom Davies, a bankrupt bookseller who took to authorship, that he was 'an author generated by the corruption of a bookseller,' is evidently a hint as to his connection with 'Grub Street' through an allusion to the popular belief.

In Shakspeare's 'Antony and Cleopatra' (Act II. Scene VIII.), the mud of the Nile is said to give birth to serpents and crocodiles by the operation of the sun. In 1662 Van Helmont relates how a green leaf placed between two bricks was transformed into a scorpion, and states that living things—fish, newts, and frogs—in ponds that have been dried up, spring from the mud ; he further tells how to generate mice out of some sawdust and an old shirt ; and in 1870 Dr. Bastian (a Fellow of the Royal Society) relates how various microscopic fungi may be developed from solutions containing only definite chemical compounds. Van Helmont enters into an explanation of the phenomena he describes, and attributes the cause to the presence of an essence which impregnates the mass, fertilises it, and makes it bring forth life. This is really a very important theoretical assumption of

Van Helmont, which has much more apparent truth in it than the facts offered in its support would lead one to believe. Most careful experiments made within the last fifteen years did not lead to any more definite result than a confirmation to some extent of this theory.

Redi in 1638 made the first discovery which undermined the belief in spontaneous generation. He found that when meat putrefied, the maggots were not generated from the flesh, but were the larvæ developed from the eggs of the common blow-fly ; that when the meat was covered with gauze the maggots did not appear, but the flies settled on the gauze nearest to the meat and there deposited their eggs. The common wire-gauze dish-cover is the practical application of Redi's discovery to household purposes. At the beginning of the eighteenth century microscopical research revealed in infusions of vegetable and animal matter a great variety of minute living things, the existence of which not being satisfactorily accounted for led the discoverers to believe that, under certain favourable conditions, matter which had once lived preserved after death a sort of vitality, under the influence of which it underwent a re-arrangement and re-construction, producing a different class of organisms. There appeared in London, in 1745, a work by Needham supporting

the doctrine of spontaneous generation by arguments derived for the first time from direct experiments. Needham's views were adopted by Buffon, but they had both partisans and opponents. The latter advocated Bonnet's theory of the pre-existence of living germs in the atmosphere, and amongst them the most distinguished advocate was the Abbé Spallanzani, who, repeating Needham's experiments, obtained opposite results. It suffices to say regarding these experiments that vessels containing putrescible substances were hermetically sealed and heated to the temperature of boiling water. Needham heated his vessels for a space of five minutes only, and they contained organisms, while in Spallanzani's experiments they were heated not less than forty-five minutes, and were found barren. Needham then complained that Spallanzani 'had put to the torture' and 'not only much enfeebled, if not totally annihilated the vegetative force of the infused substances, but also that he has totally destroyed by the exhalations and heat of the fire the small portion of air remaining in the empty part of the flask.' It is to be observed that Needham has no evidence whatever to support such statements. Unfortunately, the discovery of the composition of air by Lavoisier had not yet been made, consequently eudiometry had not yet been invented. No information, there-

fore, concerning the validity of this objection was at this time obtainable.

The real value of any laboratory experiment is put to a very severe test when it is carried out on a large scale for commercial purposes. Appert applied to domestic purposes Spallanzani's results, carried out according to Needham's plan, and the process has now developed into the preservation and importation of immense quantities of Australian meat, and the peas and other vegetables procurable in winter in a Parisian restaurant. Gay-Lussac reported on Appert's process of preserving food, and having found the air contained within the vessel to be destitute of oxygen came to the conclusion that the absence of oxygen was a necessary condition of the preservation of animal and vegetable substances, a doctrine which has been very generally received, but which is in reality most erroneous. Schwann, in 1837, first showed this, and refuted the objection of Needham that Spallanzani had destroyed the air in his flasks by prolonged heating. His experiments show that fresh, cool air, if previously heated, has no effect on the juice of meat which has been boiled, and his conclusion is, that fermentation and putrefaction are not caused by oxygen, or by means of oxygen in the air, but by a principle included in ordinary air which heat can destroy.' Schultze modified

Schwann's experiments by purifying the air admitted to his infusions by passing it through energetic chemical reagents, such as oil of vitriol and caustic potash. Schröder and Dusch, guided by the experiments of Lœwel, which showed that ordinary air was not able to provoke the crystallisation of sulphate soda when filtered through cotton wool, applied the same principle to the preservation of infusions, and by means of flasks, fitted up in the same manner as those used in Lœwel's experiments, they operated on the following materials :—

1st. Meat with the addition of water.

2nd. Beer wort.

3rd. Milk.

4th. Meat without water.

In the first two cases the air filtered through cotton wool left the liquids unchanged even after many weeks ; but the milk soon curdled and the meat putrefied. Five years afterwards, in 1859, Schröder alone returns to this subject, and after recounting numerous and various very interesting experiments, concludes thus :—‘ It must be admitted that fresh air contains an active substance which provokes the phenomena of alcoholic fermentation and putrefaction, a substance which heat destroys and cotton wool arrests. Must this active substance be regarded as formed of organised microscopic germs disseminated in the air? Or

is it a chemical substance still unknown? I cannot tell.'

In 1859 the many difficulties besetting this subject were attacked by M. Pasteur in a most impartial spirit, free from preconceived ideas, and with the object in view of founding a well-grounded opinion upon this important question of heterogenesis. The conclusions of Schwann and Schröder could leave no doubt in the mind of the reader of the existence in the air of a principle which is the necessary condition of life appearing in infusions. Those who stated that this principle was nothing but germs had no further proofs in support of such an opinion than those who thought it might be a gas or a fluid, and who consequently inclined to the belief in heterogenesis. Schwann, Schröder, and Schultze failed when experimenting in certain liquids, and Pasteur failed constantly and with all liquids when, as will be presently noticed, his liquids came in contact with mercury, the cause of these failures being at the time quite unknown, no cause of error being even suspected.

Thus we see that in times as recent as 1859—the date of Schröder's memoir—we had really no more definite information beyond that given in Van Helmont's statement in 1662, *but then it had been established by careful and exact experiments.*

About this time M. Pouchet, a renowned natu-

ralist of Rouen, published a work advocating the theory of 'Heterogenesis.' He supposed that when water, a putrescible substance, and air are brought together, certain so-called plastic forces of matter beget germs, which germs undergo the same process of development as those resulting from the normal process. *Animalculæ*, he said, were formed from the débris of others which preceded them. He repeated Schwann and Schultze's experiments, and failed to get the same results. The liquid mostly used by Pouchet was an infusion of hay, and his arrangements for excluding germs from the air in his experiments were very insufficient. He objected that if the air were a reservoir for germs it would be so overloaded with organic matter as to be quite foggy, and that these germs would not escape microscopic observation. He examined dust which had settled in different places, and found it contained grains of starch, silica, soot, and spores of fungi, but the latter, he said, were very few in number.

In 1862, two hundred years after Van Helmont, Pasteur's investigation of all the principal foregoing facts was published. The ingenuity displayed and the remarkable chain of evidence afforded by these most carefully executed, most complete, and, therefore, most trustworthy experiments, gives the whole paper, when read in detail, more the air of a romance

than the record of scientific laboratory work, at the same time that it completely sets at rest the doubts concerning the existence of countless numbers of living organisms in the air, and furnishes conclusive evidence against the spontaneous generation of living things.

*Pasteur's Microscopic Examination of the Solid
Particles diffused in the Atmosphere.*

The question which Pasteur first set himself to answer was—Is it possible to gain an approximate idea of the relation a volume of ordinary air bears to the number of germs that this air may contain? That there are germs in the air is admitted by the staunchest partisan of the doctrine of spontaneous generation—M. Pouchet. He has recognised the spores of cryptogams, and particles of starch in dust, but he adds that their number is very small. It is, however, not the dust deposited on surfaces such as Pouchet took which should be examined, but that floating in the air, for, as Pasteur remarks, the organised particles may be considered the lightest and would be carried about by draughts whilst the heavier particles would settle. This fact has been overlooked by some who, in criticising the remark of Professor Tyndall, that floating dust consisted chiefly of organic matter, which conclusion

he came to from the fact that it was combustible, said that Dr. Percy had found the dust settled on the walls of the British Museum consisted of 50 per cent. of incombustible matter. Now it is perfectly evident that, as a whole, organic matter has a lower specific gravity than inorganic, and organised organic matter may be expected to be still lighter: on this account if 50 per cent. of the dust deposited on surfaces was organic, a still higher proportion of that floating would be of the same nature, while organisms would for the most part float in ordinary air which is subject to disturbance.

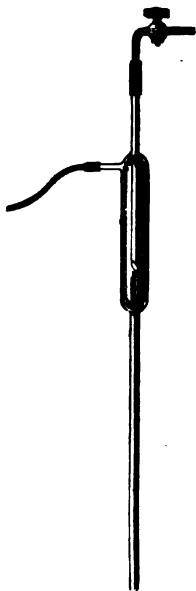
Water,¹ which is more than 800 times heavier than air, remains suspended in the form of very fine spherules, causing a mist or fog when near the earth's surface, and constituting the clouds which float high above. Solid particles of carbon are discharged into the air by chimneys in the form of smoke. Taking these facts into account, we cannot be surprised at finding the dust (which is generally speaking so fine as to be unseen) to consist of such very finely divided solid matter as is most abundant around us.

But let us see by what means Pasteur sought

¹ One cubic foot of water weighs 1,000 oz., 13 cubic feet of air weigh 16 oz. $\frac{16}{13} = 1.23$ and $\frac{1000}{1.23} = 813$, the proportion by which water is heavier than air.

to determine the nature of floating particles in the air, and what conclusions he arrived at.

By means of a water aspirator (see fig. 45), he drew air from a quiet street and from the garden of the École Normale in Paris at some distance from



The fall of water through the wide vertical tube carries bubbles of air before it, the air being drawn through the narrow horizontal tube attached to the wide one. At the point where the water is discharged the air also escapes : this can be collected in a flask full of water inverted in another vessel. It is easy thus to measure the quantity discharged in a minute.

Fig. 45. Water Aspirator.

the ground through a tube containing a plug not of cotton wool as in the experiments of Schröder, but of soluble pyroxyline, such as is used for making collodion. Pyroxyline is cotton wool which has been steeped in a mixture of strong nitric and

sulphuric acids, and afterwards carefully washed with water and dried. It is so altered in chemical properties as to be easily soluble in a mixture of strong spirits of wine and ether. It burns rapidly, and if prepared in a particular manner is converted into the highly explosive substance—gun-cotton.

By drawing air through a plug of this soluble cotton, inserted in a small glass tube, all the solid particles were intercepted or filtered off. The amount of air aspirated in a given time was accurately measured, and after a sufficient interval the soluble cotton plug was removed and treated with its solvents, alcohol and ether. After allowing the dust to subside in a tube, the collodion was syphoned off and more alcohol and ether was repeatedly added, to effect the perfect removal of the collodion. The completely washed dust was placed on a microscope slip, and examined in a drop of water. By ordinary methods the action of different reagents, such as iodine water, potash, sulphuric acid, and colouring matters, on the particles was tried.

The accompanying drawings represent some of the organised corpuscles collected during twenty-four hours on November 16 and 17, 1859. Immediately after washing the dust a small portion was placed on a slip of glass and tested with a drop of potash-solution containing 5 parts of the sub-

stance in 100 of water (see fig. 46). It was then covered with a microscopic thin glass cover and



Fig. 46.

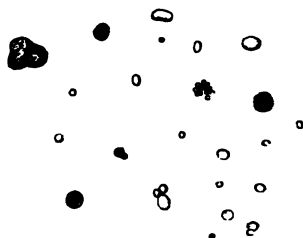


Fig. 47.

examined, and when an evidently organised particle was noticed this was drawn. The potash-solution was then withdrawn from under the glass by means of absorption with filter paper, and iodine water was substituted. Fig. 47 shows the action of iodine water and fig. 48 the subsequent

effect produced by ordinary concentrated sulphuric acid.

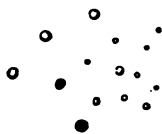


Fig. 48.

This process disclosed the fact that there is in ordinary air a variable number of corpuscles of such forms and structures as show them to be

organised: Their dimensions range from extreme smallness up to the diameters of 0.01 millimètre to 0.015 millimètre or more. Some are perfectly spherical, others ovoid. Many are translucent,

others opaque, with internal granulations. Those translucent particles of a regular shape so closely resembled the spores of the most common fungi that the most able microscopist could not see any difference. These corpuscles were evidently organised, resembling completely the germs of the lowest organisms, and so diverse in size and structure as to belong without doubt to very various species. The employment of iodine¹ shows that amongst those particles there are always some starch granules, but it is easy to separate all such granules by treatment with sulphuric acid, which dissolves them. Doubtless other particles also are dissolved, but still there remain a great number, which are often seen more clearly by reason that they are freed from the carbonate of lime and other particles of amorphous dust. But by direct experiment it was found that ordinary oil of vitriol did not dissolve the spores of common fungi, even after prolonged contact. It



Fig. 49. Starch coloured with Iodine.

¹ The iodine, as before remarked (see p. 60), strikes a deep blue, indeed inky black colour, with minute portions of starch, so that a test-tube full of starch solution in half a gallon of water with a little iodine, gives this intense colouration. See fig. 49.

is needless to say that the soluble cotton used was previously tested and found to contain no residue insoluble in alcohol and ether beyond a fibre or two.

By exposing for 24 hours a plug of pyroxyline to a current of air passing at the rate of a litre the minute after a succession of fine spring days, it was found that many myriads of organised corpuscles were collected. It may be easily understood that the time of the year, and whether before or after rain, and a thousand and one causes, may increase or lessen the number of solid particles which every one has seen dancing in a ray of sunlight entering a darkened room. Pasteur found that in the winter months, particularly once on the



Fig. 50. Dust collected on June 25-26, 1860.



Fig. 51. Dust collected during an intense Fog in February 1861.

occasion of so low a temperature as -9° C., to -14° C., a very small number of germs were collected.

The drawings, figs. 50 and 51, show both organised and amorphous particles, such as are seen in

the field of a microscope giving an enlargement of 350 diameters, the moistening liquid being ordinary sulphuric acid.

Pasteur remarks near the end of his second chapter, 'I think it would be of great interest to multiply researches on this subject and to compare the organised corpuscles disseminated in the air at one place during different seasons and at different places at the same time. It appears that our knowledge of contagious diseases, especially at periods when epidemics rage, would be increased by work carried out in this direction.' The influence of this suggestion may be traced in recent advances in surgical and medical science. By following up his own ideas Pasteur was enabled to prescribe a means of preventing the disease known as 'pébrine,' which made such havoc amongst the silkworms in France.

As mention will have to be made of those organisms which develop in decomposable organic liquids, it is necessary to give a slight description of them. The commonest and most easily observed are the minute fungi commonly known as mould or mildew. When a previously boiled putrescible liquid is left open to the air, exposed to a warmth of about 30° C. (95° F.), there generally appears a slight speck of down upon the surface: this increases to a tuft, which on examination with a

low microscopic power not unfrequently is resolved into an appearance like that seen in fig. 52. The growth consists really of two differently disposed



Fig. 52. A common form of Mould.

parts, the mycelium and the fructification. The former (fig. 53), which performs the function of

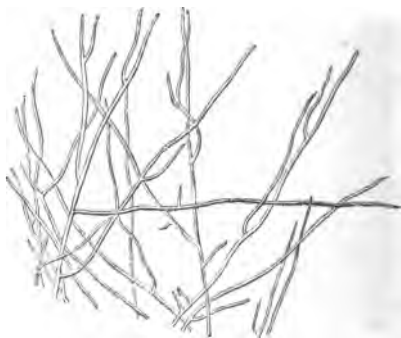


Fig. 53. Mycelium of Fungus. After Pasteur.

nutrition, is a network of tubes immersed in the liquid. The latter, raised on fine threads, consists of spores serving for the propagation of the species. In the *mucorini* the threads are terminated by sporangia (fig. 54) containing sporidia, which are set

free when mature by the bursting of the sporangium (see fig. 55).



Fig. 54. Sporangia of *Mucor*. After Pasteur.

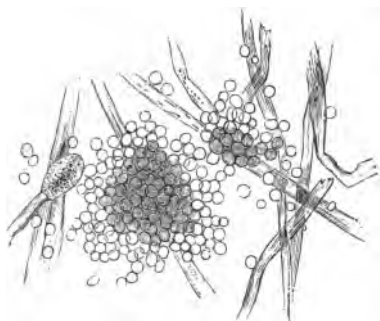


Fig. 55. Sporidia set free. After Pasteur.

Figs. 56 and 57 are forms of fructification very commonly met with. Under certain conditions fungi can be directly reproduced from the myce-



Fig. 56. Fructification of *Penicillium*. After Pasteur.

lium without spores. There are two great divisions of fungi, and the commonest forms of mould, referred to their family, order, and genus, are given here—

I. *Sporifera*, spores naked. Family, *Hyphomycetes*. Order, *Mucedines*. Genera, *Penicillium* (fig. 56) and *Aspergillus*.

II. *Sporidiifera*; Sporidia in vesicles. Family, *Physomycetes*. Order, *Mucorini*. Genus, *Mucor*.

A very frequently occurring growth, especially in saccharine liquids, is the torula¹ (fig. 57). This has



Fig. 57. Torula.

no mycelium, but consists simply of cellular organisms, which frequently elongate and branch: it is essentially the cause of alcoholic fermentation, the conversion of sugar into alcohol. Besides these there are certain organisms of much smaller dimensions seemingly allied to the Algæ, which are the constant attendants on putrefaction.

These are bacteria, and according to Cohn's

¹ According to the Rev. M. J. Berkeley, torula is but an altered form of penicillium, but De Bary does not confirm this.

researches and his definition they are 'cells free from chlorophyll of spherical, oblong, or cylindrical form, sometimes twisted or bent; they multiply exclusively by transverse division and occur either isolated or in cell families.' Division only occurs longitudinally, and such a thing as branching is never seen. They frequently form a jelly-like mass, which is called the *Zooglæa* form. Attracted by oxygen, bacteria also collect like a layer of oil upon a liquid, or as a tough pellicle, a modified *Zooglæa*, the little rod-like organisms being packed on end in rows. After a certain time they are generally to be found at the bottom of a vessel as a pulverulent precipitate; this is when the liquid no longer yields them nutriment. In a perfectly clear liquid, such as a carefully prepared turnip infusion, they make their first appearance as a slight turbidity. This is really due to their incessant and rapid motion, for after they have subsided the liquid again becomes clear. Most bacteria present a motile and a motionless condition.

The following illustration gives an idea of the appearance of different organisms classed as bacteria by Cohn, though only A and D are what may be called bacteria proper. These have all been observed with one of Messrs. Beck's best one-eighth inch objectives and the second eyepiece of one

of their microscopes. Magnifying power 750 diameters.

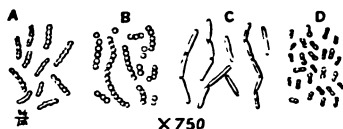


Fig. 58. Some of the most minute Organisms found in organic infusions, &c.

A. *Bacterium Lineola*—the *Vibrio lineola* of Ehrenberg—commonly found in stale milk, 'Pasteur's solution,' and other putrefying liquids. Motion very rapid.

B. *Micrococcus ureæ*. The necklace fungus supposed by Pasteur to convert urea into carbonate of ammonia.

C. *Bacillus ulna*. Seen abundantly in perfectly fresh hay infusion. Jointed rods. Motionless.

D. *Bacterium termo*. The motile form constantly seen in early stages of putrefaction.



Fig. 59. *Bacillus subtilis*.
Vibrio lineola. Pasteur.

The following (fig. 59), taken from Pasteur's drawing, is what he describes as *Vibrio lineola*; it is the ferment of sour milk, and the cause of that action which converts lactic into butyric acid. Cohn calls it *Bacillus subtilis*.

Pasteur's Experiments with Heated Air.

Although it appears there are in air organised corpuscles in great numbers which are indistinguishable from the germs of the lowest organisms, is it really a fact that amongst these there are particles capable of germination? This interesting question was answered in a conclusive manner. Firstly, the facts announced by Schwann were firmly established, although they had previously been attacked by Mantegazza, Joly and Musset, and Pouchet. The solution, sealed up in flasks, was one extremely liable to change; its composition was—

Water	100 parts
Sugar	10 „
Albuminoid and mineral matters from yeast	0·2 to 0·7 parts

Boiled for two or three minutes, and then placed in contact with air previously heated to redness, not a single doubtful result was obtained, although repeated at least fifty times; not a single trace of any organised production was seen even after eighteen months, keeping at a temperature of 77° to 86° F.; while, if the liquid be left to ordinary air for a day or two, it never fails to become filled with bacteria or vibriones, or covered with mould. The experiment of Schwann applied to this sugar solution is, therefore, of irreproachable exactitude.

Schwann, however, did not always succeed so well as he wished, and the experience of Mantegazza and Pouchet was at variance with his general conclusions ; even Pasteur himself in some experiments failed to preserve his liquids. These are the particular instances :—Five flasks of 250 cubic centimètres capacity, containing 80 cubic centimètres of the sugar solution, were boiled, and during ebullition sealed up. The points were broken under mercury, and pure gases in all cases but one let into the flasks. Organisms were found in every case after four days. In all these experiments, as in those likewise of Schwann, which were contrary to the result of his first experiment with extract of meat, it was the mercury that introduced the germs. In making such experiments with a mercury trough, preservation of the liquid will not always succeed, even if it succeeds sometimes. If the sugar solution be replaced by milk and treated by either of the methods above described, the milk putrefies. These results, so different and contradictory, find a natural explanation further on, but so far they are facts of a troublesome nature.

Germination of the Dust which exists suspended in the Air, in Liquids suitable to the Development of the Lowest Organisms.

The facts ascertained so far are :—

1. That there exist suspended in the air organised corpuscles exactly like the germs of the lowest organisms.

2. That sugar solutions with the liquor from beer yeast, a fluid extremely alterable in ordinary air, remains unchanged and limpid, without giving rise to bacteria or fungi, when left in contact with air previously heated.

The question now arises, how is it possible to sow an albuminous sugar solution with germs collected by means of pyroxyline in the manner already described ?

Taking a flask containing such a sugar solution kept at 77° to 86° F. for one or two months unchanged, in contact with previously heated air, the sealed-up end is connected by means of a caoutchouc tube with one part of a T-tube, while another is in connection with an air-pump, and a third with a platinum tube heated to redness. Between the T-tube, however, and the flask is a wide tube containing a very narrow one within it, holding a plug of gun-cotton, through which a large volume of air has been passed. The tap in connection with the

heated platinum tube was closed, and the one in connection with the air-pump opened ; after exhausting air was admitted through the red-hot platinum, the tap was closed, and the air again pumped out, fresh air being admitted through the heated tube ; this was repeated three or four times. The stop-cocks were then closed, and the sealed beak of the flask was broken within the india-rubber connection ; the plug of gun-cotton was shaken into the liquid, after which the flask was sealed up again. All experiments so performed resulted in the liquid, after three or four days' exposure to a temperature of 77° to 86° F., decomposing, and being found to contain bacteria, vibriones, and fungi, exactly like those in flasks exposed to ordinary air. There was no difference in the length of time requisite for the change, the forms of life occurring, or the nature of the change resulting in flasks so treated, and those with the same liquid exposed to common atmospheric air. These experiments can scarcely be surpassed for beauty in their arrangement, or for the importance and clearness of the evidence they afford. Yet thinking that it might be objected that the gun-cotton had given rise to the changes produced, Pasteur made use of plugs of asbestos, and found a like result ; but when the plugs of asbestos were heated red-hot previous to being put into the flasks

the liquids remained unchanged in every case, and so constantly and with such perfect exactitude after an immense number of trials did the results remain the same, that the experimenter himself was astonished.

Extension of previous results to other very alterable.

Liquids—Urine, Milk, and Albuminous Sugar Solution mixed with Carbonate of Lime.

The facility with which urine exposed to the air becomes altered, and the change which takes place is well known. It becomes turbid and alkaline, sometimes filled with bacteria, or covered with patches of *Penicillium glaucum*, *Aspergillus glaucus*, or *Mucor mucedo*. Often there is formed, when the temperature is not higher than 59° F., a pellicle consisting of a remarkable mould closely resembling *torula*, but which is believed by Pasteur to be a different species. It consists of transparent cells, often without a nucleus, and considerably smaller than the cells of beer-yeast. There is also present in urine, when alkaline from the carbonate of ammonia resulting from the changed urea, a peculiar fungus in necklace-like groups, and this organism Pasteur is fully persuaded is the cause of urea being converted into carbonate of ammonia. An interesting observation was made with regard

to the turbidity of liquids, which generally is the first sign of alteration; this is caused not merely by the presence of minute organisms, such as bacteria, but by their movements in the liquid; for when they are dead they settle to the bottom of the vessel, and the liquid becomes clear again. Many flasks of urine were treated in the manner already described—that is to say, they were boiled, and heated air was admitted to them. After preservation for months at 77° to 86° F. without change, plugs of asbestos through which air had been drawn were introduced; and then in cases where the liquid was alkaline, strings of this peculiar fungus were found invariably, and crystals of ammonio-magnesian phosphate were deposited. It was observed that *Bacteria* (see fig. 60) appear in a liquid before any other organism. These



Fig. 60. *Bacteria*.
As seen by Pasteur.

organisms are so small that it would be impossible to distinguish their germ; but even if the appearance of the germs were known, it would be still less possible to recognise them among the various particles of organised dust collected from suspension in the atmosphere.

In experimenting with milk boiled in flasks and exposed to heated air, it was found that generally in from eight to ten days, but in one case after so

long a time as a month, the milk was found to be curdled. Microscopic examination showed that the whey was filled with vibriones, often of the species *Vibrio lineola* (see fig. 59, p. 152) and bacteria. The air of the flasks showed that the oxygen was replaced by carbonic acid ; yet swarms of these vibriones were living in an atmosphere without oxygen. The most important observation which leads to an explanation of the extraordinary behaviour of milk in these experiments, is the fact that no mucor, torula, or penicillium—nothing but bacteria or vibriones—were found in the liquid. The obvious conclusion is, that these organisms or their germs are not destroyed by a temperature of 212° F. when the heated liquid which serves to develop them enjoys certain properties. To test this supposition, the milk was boiled under pressure so that the temperature was raised during ebullition to 230° F., and then heated air was admitted, of course at the usual atmospheric pressure ; flasks treated in this way were kept an indefinite period without the production of any life whatever. The milk preserved its flavour, its odour, and all its properties. Sometimes a slight oxidation of fatty matter took place, as could only be expected in such a considerable body of air ; this was proved by an analysis of the air. In such cases the milk had a slightly suety taste. But what condition prevents

the development of vibriones in sugar solutions and urine when heated to 212° F.? It is the fact that they contain a trace of acid. Milk is an alkaline liquid.¹ If a liquid of the following composition :—

Sugar	10 grms.
Yeast water	100 c. c.
Carbonate of lime	1 grm.

be boiled in flasks at 100° C., filled with heated air and sealed up and left to itself at 25° to 30° , in from two to four days it becomes turbid from vibriones, which have a very lively motion. It was found that a species of mucor after a time covered the surface of the liquid. It seems, therefore, that under these particular conditions the germs of this cryptogam had resisted the temperature of boiling water. An important confirmation of these experiments regarding the failure of a temperature of 212° F. to destroy certain germs here follows. Milk which had been preserved some months had a plug of asbestos presumably containing germs introduced into it by the method already described ; it was sealed up, and the flask was then plunged into boiling water ; in eight days bacteria and vibriones were found in swarms. It was further discovered that 226° F. was too low a temperature to effect the preservation of these liquids.

¹ More strictly speaking, milk is one of the few liquids, according to Prof. Lister, which show both an alkaline and acid reaction to litmus.

It cannot be too forcibly impressed on the reader by what means and with what success Pasteur demonstrated the fact of myriads of organisms occupying comparatively small volumes of air. This is a point to which his detractors have willingly made themselves blind ; they tell us the organisms are few in number without any experimental proof ; while, on the other hand, as will be presently seen, Dr. Angus Smith and Mr. Dancer estimated that there were $37\frac{1}{2}$ millions of organisms, many of which were recognisable, in 2,500 litres of Manchester air.¹

Another Method for showing that all the Organisms produced by previously heated Infusions have for their origin the particles which exist suspended in ordinary Atmospheric Air.

Says Pasteur, 'I believe it to be rigorously established in the preceding chapters that all the organised productions of infusions previously heated, have no other origin than the solid particles which are always carried in the air and left deposited constantly upon everything. Could there still remain the least doubt of this in the mind of the reader, it will be dissipated by the experiments I will now describe.'

¹ 'Air and Rain,' p. 305.

The experiments consisted in placing in glass flasks the following liquids, all of which are very changeable in contact with ordinary air, yeast liquor, sugar solution and yeast liquor, urine, beet-root juice and infusion of pears; the flasks were then drawn out so as to have a long neck with many bends in all directions. The liquid is boiled for some minutes, while the steam escapes plentifully from the open neck; the flasks are then left to themselves without being sealed, and, strange to say, though the air enters, the liquid may be preserved for an indefinite period—an interesting fact for those who are accustomed to make experiments of such a delicate nature as this subject requires. There is no fear of transporting these flasks from place to place, or submitting them to the varying temperature of the seasons; the liquids show not the slightest alteration in taste or smell; they are truly specimens of Appert's food-preserving process. In some cases there was a direct oxidation of the matter, a purely chemical process. But it has already been shown how this action of oxygen was *always limited when organised productions were developed in liquids*. The explanation of these new facts is, that the air on first entering comes in contact with water vapour at the temperature of 212° F., and is so rendered harmless; what follows enters but slowly, and leaves its germs or particles

of active matter in the moist curvatures of the tube-neck. After remaining many months in a warm place, the necks of the flasks are cracked off by a file mark without other disturbance, and in twenty-four to thirty-six or forty-eight hours, fungi and other organisms make their appearance in the usual manner.

The same experiments can be made with milk, but then the milk must be boiled under pressure ; milk has been kept for months in these open flasks without change at a temperature of 77° to 86° F. The production of organisms can always be started in these flasks by briskly shaking the liquid or by sealing during ebullition, and after cooling allowing the air to enter suddenly by breaking the point of the tube.

Many such flasks, exhibited at the Academy of



Fig. 61. Solutions in flasks similar to those of M. Pasteur.

Sciences, were preserved with their contents unchanged for eighteen months, although extremely

prone to decomposition. Here, in the annexed wood-cut (fig. 61), are represented two flasks similar to those of Pasteur, and two plugged with cotton wool after the plan of Schröder and Dusch ; in the centre is an open flask of liquid made at the same time and treated in exactly the same way as the others, but instead of being still bright and clear like them, it has now become a mass of brown filth. When the photograph for this illustration was taken, the flasks with their contents were three years old.

‘ The great interest of this method is, that it unquestionably proves that the origin of life in infusions which have been boiled is solely due to solid particles suspended in the air. Neither a gas, divers fluids, electricity, magnetism, ozone, things known or hidden causes, there is absolutely nothing in ordinary atmospheric air which, failing these solid particles, can be the cause of the putrefaction or fermentation of the liquids which we have studied.’ It has so far been definitely proved by Pasteur, and stated in the following manner :—

‘ 1st. That there are constantly, in ordinary air organised particles which cannot be distinguished from the true germs of the organisms found in infusions.

‘ 2nd. When these particles and the amorphous *débris* associated with them are sown in liquids, which have been previously boiled and which re-

maintained unchanged in air previously heated, there appear in these liquids exactly the same forms of life as arise in them when they are exposed to the open air.'

Such being the case, could a partisan of spontaneous generation wish to uphold his principles even in the face of this double proposition? He might, but then his argument would necessarily be of the following kind, of which I leave the reader to judge for himself. There are in the air, he might say, solid particles, such as carbonate of lime, silica, soot, fibres of linen, wool, and cotton, starch granules and besides these organised corpuscles having a perfect resemblance to the spores of the Mucedinesa or the germs of Infusoria. I prefer to attribute the origin of Mucedinesa and Infusoria to the first amorphous substances rather than to the second.'¹

This has actually been urged. Could there be more eccentric reasoning? Reasoning it is not. That question is beyond the pale of argument, to which common sense dictates the answer.

¹ In mention here made of *Infusoria* it must be recollected that Pouchet stoutly maintained that the true *Infusoria* were subject to heterogenic reproduction.

It is not exactly true that the smallest quantity of ordinary Air gives rise in an Infusion to the Organisms peculiar to this Infusion. Experiments on the Air of various Localities. Inconvenience of employing Mercury in Experiments relative to Spontaneous Generation.

If the smallest quantity of air in contact with an infusion gives rise to organisms, and these organisms are not of spontaneous origin, then it follows that in the minute portion of air there must exist a multitude of the germs of very different organisms ; in such numbers, too, that, as Pouchet says, the air would be so loaded with organic matter as to form a thick fog. Strong as this reasoning is, it would be still stronger if it were shown that different forms of life are derived from different germs : this may be so, but it has not been proved.

Experimental proof of this statement, the error in which lies in gross exaggeration, was made by sealing up during ebullition flasks of 250 cubic centimètres capacity containing about 80 cubic centimètres of various liquids. On breaking the points of these flasks in certain noted places, the air entered with a rush into the empty space, carrying the germs along with it ; after resealing, the flasks were placed in a warm situation and any

change noted. In some cases the decomposition followed, and the production of the usual forms of life; in other cases the flasks remained as if they had been filled with heated air, quite unchanged. In two experiments made in the open air after a slight shower in the month of June, both resulted in the production of organisms; in four others, after a heavy rain in the same place, two of the flasks had their contents remain unchanged for at least thirteen months afterwards. These experiments were made, it is easily seen, in an agitated air, but Pasteur carried his labours into the cellars of the Paris Observatory, where the air is quite still except when agitated by the movements of the experimenter, and in that region below the surface of the earth where the temperature is unaffected by the changes of the seasons. It is to be expected that air, in which there is so little to cause its disturbance, would have deposited on the ground the germs which at one time floated in it. A greater proportion of flasks, therefore, if opened and resealed in such an atmosphere, should have their contents preserved. Out of ten experiments made under such conditions with yeast water, in only one was any living thing found; while eleven experiments made in the court-yard of the Observatory at a distance of 50 centimètres from the ground, and at the same time, rendered in every case the

usual forms of life ; a modification of these trials was made by letting air into flasks of liquid at various mountain heights. Eighty-three flasks, prepared in the manner already mentioned, were experimented on : twenty of these were filled up with air at the foot of the heights which form the first plateau of the Jura ; twenty others on one of the peaks of the Jura, 850 mètres above the sea-level ; and the remaining twenty were carried to Montanvert, near the Mer de Glace, at an elevation of 2,000 mètres. The result was, that of the twenty opened on the lowest level, eight contained organisms ; of the twenty on the Jura, five only contained any ; and, lastly, of the twenty filled at Montanvert, while a strong wind blew from the deepest gorges of the Glacier de Bois, one only was altered. The method of opening the flasks was to hold them with outstretched arms and the point of glass turned from the wind, and by a pair of iron forceps, which had just been heated in a spirit-lamp flame, the point was broken. The drawn-out point had been previously scratched with a file and heated ; otherwise, particles of dust adhering to the glass would have been carried into the liquid by the in-rush of air.

A remarkable and interesting fact connected with these experiments was, that on one occasion Pasteur opened his flasks, and, on account of not

being able to see the flame of his lamp against the brilliancy of the snow, it was impossible to re-seal them; the flasks were necessarily carried back to the little inn at Montanvert to be closed up. Every one of these flasks contained organisms after keeping for a short time. On the Glacier, then, there are no germs in the atmosphere, but at the neighbouring inn the air warms with life, and life from all parts of the world, brought by the travellers. On opening the flasks they were held above the head, so as to prevent the possibility of germs attached to the person being deposited in them.

Explanation of the Cause of Failure of the Experiments in which Mercury is used.

Flasks containing liquids which had been kept for a great length of time were connected with an air-pump and a red-hot platinum tube: after repeated exhaustion and re-filling with heated air, the communication was made between the flask and the platinum tube, and a globule of mercury taken out of a mercury trough in a laboratory, which had previously been introduced into the connecting-tube of india-rubber, was made to roll into the flask; on re-sealing and keeping for a few days, fermentation ensued in every case, just as certainly as when the asbestos plugs and the adhering germs

were sown in similar liquids. This case leaves no doubt regarding the cause of failure of experiments in which the liquid comes in contact with mercury by the flasks being broken under the surface of the quicksilver.

There are other facts which Pasteur established, of great interest and importance in connection with the nutrition of minute fungi. Instead of experimenting on milk, urine, or solutions containing the liquor from yeast, he made use of such an infusion as the following; that is to say, a mixture of perfectly definite chemical substances:--

Pure water	100 parts
Sugar-candy	10 „
Tartrate of ammonia	0·2 to 0·5 part
Ashes of yeast	0·1 part

On impregnating such a liquid, when supplied with heated air, with germs collected from the atmosphere, bacteria, vibriones, and fungi, &c., were soon developed; the albuminoid and fatty matters, the essential oils, and pigments belonging to these organisms being derived from the elements of the ammonia salt, the phosphates, and the sugar. These complete organisms were built up out of the material afforded by such a mixture of simple substances, a fact which is quite contrary to Pouchet's declaration that ovules or germs were evolved from a sort of vitality remaining in lifeless, or, rather,

dead matter—that is to say, matter deprived of life.

A solution consisting of—

Pure water	100 parts
Sugar-candy	10 „
Tartrate of ammonia	0·2 to 0·5 part
Yeast ashes	0·1 part
Pure calcium carbonate	3 to 5 parts

showed much the same phenomenon, in fact, differing only by a more marked tendency towards the changes called lactic, viscous, and butyric fermentations; and all ferments, whether vegetable or animal, characteristic of these changes were produced, simultaneously or successively.

Professor Tyndall, in 1870, gave us a means of investigation, supplementary to the microscope, and of extreme delicacy. Aided by Professor Huxley, he proved that particles in a liquid, quite invisible under an object-glass readily showing bodies $\frac{1}{100000}$ of an inch in diameter, were revealed with the greatest ease by means of a beam of light. If the air were pure, a beam of sunlight traversing a darkened room would be invisible except where it struck upon the wall. It is the scattering of the light by the floating dust which makes the track luminous; the larger and more numerous the particles the greater the luminosity. Hydrogen, coal-gas, air passed through cotton-wool, and the air of

still places, were found to be free from floating matter. Having devoted much attention to this subject since 1865, I made use of this discovery to aid me in a very careful repetition of some experiments published by Dr. Bastian.

The following description of the method pursued in this research is taken from the 'Proceedings of the Royal Society for 1872 : '— .

Experiments concerning the Evolution of Life from Lifeless Matter.

On June 30, 1870, there appeared in 'Nature' a paper by Dr. Bastian, entitled 'Facts and Reasonings concerning the Heterogeneous Evolution of Living Things ;' the perusal of this, and its continuation, led to the belief that another interpretation might be put on the results obtained by Schwann, Pasteur, and others, not so much by virtue of the arguments made use of, as by accounts of experiments given in detail. The most remarkable case was that of Exp. 19, in which the author gave a drawing of a large organised mass obtained from a solution of sodium phosphate and ammonia tartrate, which had been exposed to a temperature varying between 294° F. and 307° F. for four hours. This organism was seen to grow within the flask till it attained a certain size, beyond which it did

not increase. Now a fact so distinctly stated as the production of an organism, and its development to a considerable size, from a liquid containing nothing further than phosphate of soda and tartrate of ammonia, in a flask from which the air had been most thoroughly withdrawn, and which, when containing the liquid and hermetically sealed, had been heated to so high a temperature, was (admitting the conditions and performance of the experiments to be faultless) an absolute proof of the evolution of living matter *de novo*.

While adhering as strictly as possible to the conditions of Dr. Bastian's experiments, it was necessary to devise some refinement on the mode of examining the liquids experimented on without exposure to atmospheric air, for even the presence of one single fungus spore might be regarded as evidence of this process of generation.

The most promising plan seemed to be to open the sealed vessels in an atmosphere artificially prepared so as to be free of living matter. Hydrogen being fourteen times lighter than common air, may remain in contact with it without risk of contamination by floating matter; indeed, Professor Tyndall's demonstration, by means of a powerful beam of light, that such an atmosphere is free from dust, was sufficient to warrant its use. The means whereby this fact was made of further practical value are the following :—

1st. *The experimental tubes* in which the infusions and solutions were heated were made of ordinary combustion tubing drawn out at the lower end, first to a finer tube $\frac{1}{4}$ the diameter of the original, and after a space of an inch or so to extreme smallness. The solution or infusion was then made in a flask with distilled water, drawn by a siphon from a carboy after standing at rest for many hours, the siphon dipping into the middle of the liquid. The flask was, after the usually careful cleansing that chemical vessels require, rinsed out with a solution of potassium dichromate, mixed with strong sulphuric acid, then washed out with distilled water. A further quantity of distilled water collected in this vessel was used for the solution. The experimental tube, with the lower end drawn out but open, was cleansed with acid dichromate, and afterwards with hot distilled water. The fine point was open, in order to let the liquids run through, otherwise a drop might have collected in its capillary portion, and, not being dislodged, would have interfered with the experiment. This tube, then, best described by fig. 1, p. 175, had the fine end sealed up, and the liquid to be experimented on poured in till it about half filled the tube. The upper part was then drawn out, so as to serve a purpose yet to be named, and also in such a manner that it could be adapted to the Sprengel exhaustor ;

after this operation it had the form of fig. 2. It was then fitted to the Sprengel by means of a tube shown in fig. 3, which admitted of connection by means of two Sprengel joints, the one over the experimental tube being made air-tight with water. The reason for water being used was this: had there been a leak, water would have entered the

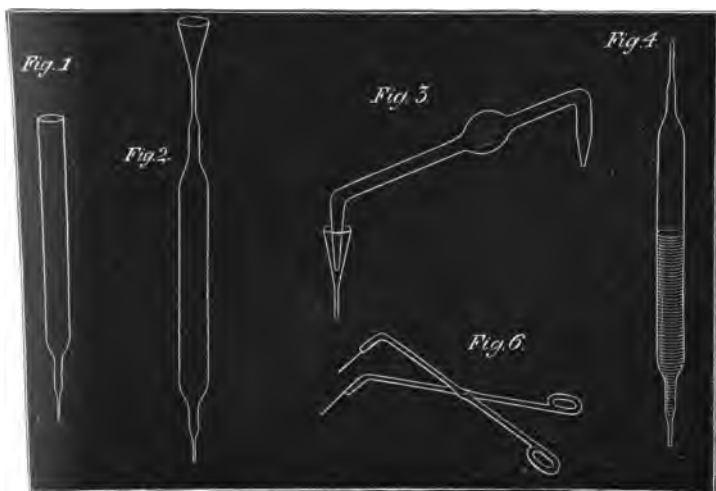


Fig. 62.

tube, and so no damage would have resulted; but had it been mercury or glycerine, the tube would necessarily have been rejected. The use of the bulb on the connecting-piece was to catch the water which boiled or distilled over during the exhausting

process. After complete exhaustion, recognised by the clicking of the falling mercury, a blowpipe-flame was cautiously applied to the fine tube till it fused thoroughly, and it was then drawn away from the other portion. What was so far accomplished was the sealing of a solution or infusion in a vacuous tube of the annexed form (fig. 4). The tubes thus prepared were heated in an air-bath, consisting of horizontal iron pipes surrounded by two iron jackets. The tubes lying horizontally were not in contact with the liquid on the upper part ; so, after being heated in the one direction for a period, at Professor Tyndall's suggestion they were turned over and heated anew, so as to bring every part of the tube into contact with the heated liquid. These tubes had generally a bead of glass fused to one side, so that if the tube were heated with the bead uppermost, it had to be turned over and heated again, the bead being downwards. Generally four tubes were heated at a time ; and one of these was soon after cooling opened to allow access of air, in order to observe whether any change occurred differing from any that might take place in the sealed tubes.

2nd. *Apparatus for Examining the Contents Microscopically out of Contact with Air.*—This consisted of a bell-jar closed at the top with a bung, and supported on a tripod ; this bell jar was kept

filled with hydrogen by means of a self-regulating apparatus continually passing a gentle stream of gas into the upper part of the jar by means of a glass tube. The bung was bored with an eccentric hole; through this passed the tube containing the liquid for examination; the end of the tube passed very little below half-way down the bell-jar.

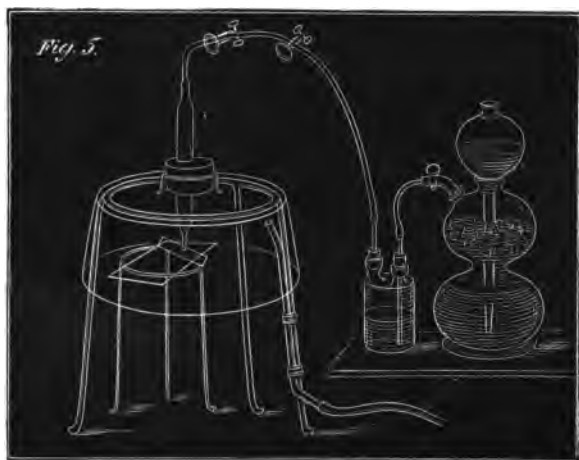


Fig. 63.

Under this was a small tripod, on which rested a glass plate to be used as a stage for the microscope slips. The tube then being *in situ*, over the upper point was tightly slipped a piece of non-vulcanised india-rubber tube, connected with a constant hydrogen generator. It was not deemed advisable

to make use of coal-gas, because, had any lifeless organism been found in a tube, it might have been objected that a trace of benzole or naphthalene vapour or other impurity had been fatal to the experiment. On the india-rubber tube were two burette clips; now, by breaking the fine point within the india-rubber (a scratch with a file being previously made upon it), hydrogen flows into the vacuous tube. Both clips are now closed, and by means of forceps, the hand being beneath the bell-jar, the lower end of the tube is broken off. No liquid, however, escapes, because the internal pressure is not much in excess of that of the atmosphere. The condition of things now can be explained only by the aid of a drawing (fig. 5—in fig. 62—p. 175); the whole arrangement consists of a pipette containing the experimental liquid, above and below which is an atmosphere of hydrogen; each drop of liquid expelled is received on a slip of glass in such an atmosphere. A drop of liquid is deposited by squeezing the space of india-rubber between the two clips, that clip nearest the bell-jar being open; before removing the pressure, it is again closed, and the partial vacuum made by compressing the india-rubber is filled up by allowing gas to flow in from the apparatus. This precaution prevents the possibility of the atmosphere of the bell-jar bubbling up into the tube after a little fluid has been discharged.

3rd. *Treatment of the Glass Slips.*—These are heated in an air-bath to about 390° F., taken out while hot with forceps, and placed on the stage in the hydrogen to cool, and kept there till wanted.

4th. *The Glass Covers.*—These are washed in the acid-dichromate solution, then in distilled water, and finally in alcohol, and, being picked up by a pair of peculiar forceps, are heated over an argand burner till very hot; they are then held in the glass vessel full of gas till cool enough to use. The forceps I devised (fig. 6—in fig. 62—p. 175) have points of watch-spring steel, so that a thin glass cover may be firmly gripped without breakage. They are made by cutting a small pair of crucible tongs an inch or so before the part where they bend at right angles; they have then two pieces of thin brass rivetted on, which are bent at right angles an inch or so from their ends; the points are made by rivetting on pieces of watch-spring steel a little more than an inch in length. When these tongs are held in the position of scissors, that is to say, with the thumb above the forefinger, the ends point downwards. To take up a glass cover, the forceps being in the position mentioned, the wrist is turned over from right to left while the elbow is raised, the glass resting on the lower point while the upper is closed down on it, may be safely held

and conveyed to where it is wanted. After a little practice these pincers are easily used.

The advantage of the bung closing the neck of the bell-jar being bored eccentrically is, that by simply turning the bell-jar horizontally the pipette point delivering the liquid may be shifted from a glass slip on which a drop of liquid has been deposited to another clean one, or be made to turn through such an angle as to be out of the way of the glass stage, in order to transfer the solution to a flask for further experiment.

Tartrate of ammonia was prepared by neutralising a tartaric-acid solution with aqueous ammonia; this was mixed with phosphate of soda solution, made by dissolving carefully washed crystals of the salt in hot distilled water. The mixture containing about 5 per cent.¹ of the two salts was slightly acidified with tartaric acid. It was found that in no case should filtering be resorted to if possible, as the finest Swedish paper transfers myriads of its fibres to the liquid. With proper care, however, filtration is unnecessary when dealing with solutions of definite salts.

¹ No mention was made by Dr. Bastian of the strength of his solutions, so that this, it is natural to suppose, he considered of no importance. In order to guard against error, these liquids were always exposed to the air to ascertain the fact that fungi would be developed in them, and some form of life was found to make its appearance in every case.

Modifications of Experiments.

In the renewed examination of liquids kept some time in sealed tubes, commenced in July 1871, a slight modification in the original method of proceeding was used. A bell-jar was chosen, the upper mouth of which was ground perfectly flat at the edge. Instead of inserting a bung with a hole in it to receive the tubes, a metal disk, with a wide metal tube placed eccentrically and projecting half an inch, was luted on to the mouth of the jar by means of grease, or, better still, what is known in pharmacy as *resinæ ceratum*. The glass sealed tube was then slipped into an india-rubber conical stopper, or rather ring;¹ for the thickness of it was so slight that a tube of any size could be made to fit it, either by the india-rubber stretching when the tube happened to be large, or by binding with a piece of copper wire when it fitted loosely. The pipette was scratched with a file at each end, and over the upper one was slipped a piece of india-rubber tube, attached to a tube of glass about $\frac{1}{4}$ inch in bore and 4 inches long, tightly packed with cotton-wool. The caoutchouc tube was pinched by a burette clip, and the extremity of the tube enclosed by the caoutchouc was broken

¹ These things are made and sold for the purpose of fixing the taps into beer-barrels.

at the file-mark. The vacuum was considered good if the india-rubber tube collapsed completely; the burette-clip was opened, and filtered air thus admitted into the vacuous space. In order to render anything that might be attached to the interior of the india-rubber tube harmless to the experiment, it was dipped in glycerine and the glycerine squeezed out of it, or treated in the same way with melted bees'-wax or paraffin. The pipette fits into its place in the disk by means of the flexible stopper. By closing the burette-clip, the tube can be broken at the lower point without more than a drop or two of the liquid escaping. After about one-third of the liquid had been examined, one-half of the remainder was allowed to run into a flask which had been previously heated to between 390° F. and 570° F.¹ The tube was then removed, and the fine capillary point, when possible, sealed at a gas-flame. The finer the point the more easily is this accomplished. A portion of liquid remains in the tube. On heating rather strongly a little of this is driven out, and then no air can pass to the remaining liquid without passing over red-hot glass, which readily melts together. The tube and flask were then placed side by side in a warm place to undergo further

¹ That is to say, baked in an oven the bottom of which was red-hot.

observation. If the tube, or class of tubes, were called A, after opening it was labelled A', and the liquid out of it exposed to the unfiltered air, A''. The tubes and flasks labelled thus were kept in a cupboard, the bottom of which was the metal lid of a long water-bath. It was thought better not to place the flasks or tubes in water, because the aqueous vapour which would thus surround the mouths of the flasks would create an abnormal atmosphere which might or might not affect the experiments; besides, such a plan is not so cleanly.

The objective made use of was obtained from Messrs. R. and J. Beck. It was a $\frac{1}{8}$ glass, without any immersion arrangement, and gave, with the second eyepiece of one of their microscopes, a magnifying power of 750 diameters. Occasionally, for convenience in drawing, a power of 420 diameters was employed.

Method of examining a Liquid which it was difficult to retain in the Pipette-tube.

When it happened that the finely drawn-out end of the pipette was too large to retain the liquid, it was allowed to run into a small glass vessel, really a beaker cut down so as to measure about $1\frac{1}{2}$ inch in diameter and 1 inch high. Drops of the solution were removed from this to the glass

slides while it stood on the glass stage in the jar of hydrogen, by means of a tube like a very long-legged siphon, at the lower end of which was a piece of caoutchouc tube, with one end stoppèd by a little piece of glass rod. This was nothing more than a bent pipette; by compressing the india-rubber when the pointed tip of the shorter limb was dipped into the beaker-glass, and then releasing it, the liquid entered for the space of an inch or so, and could then be easily transferred to a glass slip. It was thought as well to blow hydrogen through the tube before use; and, of course, like all the other apparatus, it was carefully washed and heated.

Preparation of Solutions.

The water used was very pure distilled water taken from a carboy, the contents of which had been tested with a beam of light, and found to reflect chiefly the blue rays. A previous attempt to obtain pure water by distillation with sulphuric acid and potassium permanganate, in glass vessels and an atmosphere of hydrogen, did not yield better specimens. It is impossible to prepare solutions of salts which do not show abundance of floating matter to a ray of light, even when such pure water is made use of. Solutions filtered through the finest Swedish paper are crowded with

fibres, which may readily be seen by filling a globular flask with the solution, the eye and an argand burner being on the same horizontal line, and about a foot apart. The flask is interposed, and gradually lowered till the particles are seen brilliantly illuminated on a dark ground. The phosphate of soda used was recrystallised immediately before being dissolved, and the tartrate of ammonia was prepared from recrystallised tartaric acid and the strongest aqueous ammonia. When the solutions were mixed, the alkaline reaction was neutralised by tartaric acid, or rendered faintly acid. Of course it is of the first importance that the tubes, after being sealed, should be heated immediately to the temperature necessary to destroy life, and this was done in every case as soon as possible.

Without entering into details concerning each particular tube, it may be stated shortly that in the liquids when kept in vacuo for from six to twelve months, and afterwards exposed to filtered air for periods of from two months to a year, nothing at all was discovered, while those portions of the liquids let out into perfectly clean flasks, previously heated to the temperature of boiling mercury, became always filled with life.

The figs. 64, 65, and 66 were noticed in the contents of different tubes ; also fig. 58 B, p. 152.

It may doubtless appear to some that this particular mode of experimenting involves the



Fig. 64. Confervoid Growth.



Fig. 65. Torula (rough sketch).

introduction of needless complications, that others have obtained the same results without the use of such apparatus ; but the incorrectness of this will be admitted when it is pointed out that the necessity of being guarded, of being safe, indeed, from

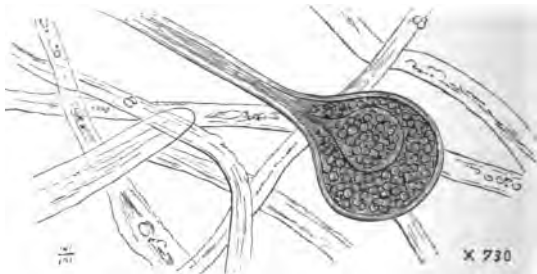


Fig. 66. A species of Mucor.

every possibility of error, is absolute. Had living things been discovered, that the experimental

method and apparatus was a matter of the first importance would have been evident, inasmuch as it excluded all possibility of atmospheric contamination of the experimental liquids during examination. In a word, it was necessary to be prepared for any result, and be guarded on every side.

The appearances described in Exp. 20,¹ by Dr. Bastian, are, with the exception of the fungus-spores and bicellular bodies, exactly what one sees in silica. Having ascertained the fact that phosphate of soda, and especially when not neutralised with acid, attacks glass tubes at a temperature of 302° F., and as in this case sodic phosphate and ammoniac carbonate² were heated for four hours at 294° F. to 307° F., some of the silica deposited from these experimental tubes was examined, and gelatinous matter, resembling that met with in solutions containing organisms, was noticed: there were two or three transparent spherules also, most probably drops of water enclosed in silica; they are often seen in pectised silica. As for the matter becoming stained by magenta, that is no evidence of any organic nature, this property being shared by silica.³ It may further be remarked that magenta would be precipitated on addition to such

¹ 'Nature,' vol. ii. p. 200.

² The reaction of this liquid is described as being neutral, which is evidently a mistake, because both these salts are very alkaline.

³ 'Journal of the Chem. Soc.' vol. ix. p. 452.

a solution by the alkaline phosphate and carbonate ; those parts more deeply stained than the others would be those where rosaniline was precipitated ; it would be impossible to use a salt of rosaniline for the purpose of detecting albuminous matter.

Dr. Bastian considers he has established by experiment the theory that living organisms, amongst which are vibriones, and fungi of the genera *Mucor*, *Penicillium*, and *Torula*, and algæ, such as conferva, are evolved *de novo* from lifeless matter ; he brings together a number of reasons, of a more or less decided kind, to show not only why he thinks it should be an intelligible process, but also why others, and particularly M. Pasteur, have obtained results leading directly to an opposite conclusion. These arguments are not drawn from experimental evidence ; they do not therefore fall within the bounds of this discussion ; but it should be observed in one case not only do Dr. Bastian's own experiments deny the truth of a most important assumption of the evolutionists, but also at the same time these experiments prove the contrary. He says :¹—‘ The disruptive agency of heat is fairly enough supposed by the evolutionists to destroy some of the more mobile combinations in each solution—to break up more or less completely, in fact, those very complex organic pro-

¹ ‘ Nature,’ vol. i. p. 176.

ducts whose molecular instability is looked upon as one of the conditions essential to the evolutionary changes which are supposed to take place.' Before granting such a supposition, it would be necessary to know, first, what are 'the very complex organic products' of such peculiar 'molecular instability' existing in a solution of tartrate of ammonia, sodic phosphate, acetate of ammonia, oxalate of ammonia, in a solution of sugar and calcined yeast, in turnip infusion, or any other putrescible liquid. These experiments show that there is no such disruptive agency in a high temperature, that it does not influence the 'more mobile combinations,' either in solutions of organic salts or vegetable infusions, for certain tubes in which the ordinary forms of mould developed were the same as the others of the series, contents identical, heated at the same time to the same temperature, in fact taken from among them indiscriminately, the only difference being that one was exposed to the air and the others were not. Besides, there is, in addition, the evidence afforded by the tubes being kept, firstly, in a vacuous condition, secondly, supplied with filtered air, and, finally, freely opened; yet we find the changes occurred in them as readily as in the unheated original solutions. *Dr. Bastian records¹ the development of organisms in a liquid*

¹ 'Nature,' vol. ii. p. 200.

heated as high as 307° F. ; yet the assumed 'disruptive agency of heat' is supposed to have influenced the results of Schwann and Pasteur at a temperature of 212° F. ! His experience is contradictory to his own theory, and at the same time to the experiments of others, to which his theory raises objection.

If it be asserted that dead nitrogenous organic particles in the air, and not germs, are the cause of organic changes, how are we to account for the action of various antiseptics? In the Report to the Cattle-Plague Commissioners 'On Disinfection and Disinfectants,' by Dr. Angus Smith, an account is given (p. 10) of the action of a number of essential oils, such as oil of bitter almonds, oil of mustard ; amylic alcohol, cresylic and carbolic acids, and ether, the vapour of which was diffused in air surrounding pieces of meat. Many substances had the property of preserving the meat for a great length of time, especially amylic alcohol and oil of bitter almonds ; chloroform and carbon tetrachloride also share this property. If these nitrogenous particles are not living things, how are we to account for the action of these substances? It cannot be a chemical action, because the substances are chemically inactive. In some cases the less active preservative agents having diffused away out of the bottle, mould formed on the meat, or putrefaction commenced ; and this always happened on that

part nearest the cork, showing that particles causing the change came in with air after the preservative agent had escaped. That these preservatives were fatal to the growth of mould, or the spread of putrefaction, was ascertained by placing mouldy paste under a bell-glass the atmosphere of which contained a small quantity of the vapour.

If those particles which are the origin of life are themselves lifeless, they could not influence the nature of the organisms developed; in that case different solutions would give rise to different organisms, and, conversely, the same solution would yield the same form of life. But we do not find this to be the case; for in the same infusion of turnip, there are found on exposure to the air, even at the same times, different forms of life. In the one case minute vibriones and masses of *torula*-cells occurred, whereas in the other confervoid growth and some minute motionless organisms which had not previously been observed were the most noticeable forms. With different portions of a solution of alkaline phosphates and tartrates, in one case we got a fungoid and in the other an algoid form. Were it not for the great preponderance of sound experimental evidence to the contrary, it would be easier to believe the theory that life was evolved *de novo*. It is the crude idea which a superficial observer of everyday phenomena would entertain;

the popular error that maggots are bred out of corruption illustrates this. The theory involves the discovery of a new property of matter, the property that certain compounds (undefined nitrogenous particles in the atmosphere) must have of decomposing molecules of other substances with which they are in contact, and building out of their constituent atoms, substances of a much more complicated nature, without the exertion of external forces; even beyond this, they must be capable of arranging those compounds into definite forms. On this account a very great deal of thoroughly sound experimental evidence is necessary to establish the doctrine of evolution of life *de novo*. But, so far as our present knowledge guides us, whether we term it spontaneous generation, abiogenesis, or archebiosis, the process by which living things spring from lifeless matter must be said to be only ideal.

This is so far a *résumé* of the discussion of this subject published in 1872. Since then Dr. Bastian has written two books, 'The Origin of the Lowest Organisms' and 'The Beginnings of Life,' the direct reasoning of which rests on his own experiments, which, placing the highest possible estimate on them, may be said to be insufficient. Perhaps the further discussion of one or two points may not be without interest.

If a little liquid from one of those tubes which were exhibited at the Royal Society when the foregoing paper was read is let out even now into a clean superheated flask, although at first perfectly clear, sweet, and fresh in flavour and odour, there soon appears a growth of mould upon the surface, and it acquires a very foul smell. Indeed, these experiments prove, 1stly. That Dr. Bastian probably mistook silica for organic matter in at least one case; 2ndly. That there is no so-called disruptive agency in a high temperature; 3rdly. That hay infusion contains at the outset so many organised corpuscles of various forms that no safe and conclusive experiments on the subject of evolution *de novo* can be made with it; 4thly. That fungi and bacteria are not generated in infusions previously deprived of all living organisms under any circumstances unless unpurified air gains access.

It is impossible to obtain infusions, and even solutions of crystalline salts free from organic impurities; still it is possible and indeed necessary to free them from solid particles of any considerable size, and it may here be remarked that Dr. Bastian's flasks being merely washed with hot water could not be clean. No experiments can be called trustworthy if they have been made with liquids containing solid lumps of dirt or organic matter, such as cheese or muscular fibre.

Cheese, from its greasy nature, would not easily come in contact with the surrounding fluid, and from its slow conductivity for heat, would take a long time before reaching the temperature of the surrounding fluid. The most minute particles might indeed be those least likely to get heated, by reason of air bubbles surrounding them, as is the case with *Lycopodium* spores and powdered arsenious acid when thrown into water. Yet Dr. Bastian, in his 'Beginnings of Life,' draws conclusions as to the evolution of living things generally from lifeless matter by his own experiments on turnip infusion, containing solid visible lumps of cheese, which mixture had been boiled merely at the ordinary temperature for so short a period as ten minutes.

The immense importance of Pasteur's work can only be appreciated when we consider that he unquestionably established the fact not only that there exists in the air a very large number of *living* germs of divers species—a fact in itself quite sufficient to make us extremely cautious in admitting the theory of heterogenesis—but furthermore that liquids protected in every way against contact with these germs could be preserved for any length of time.

The arguments lately raised that Pasteur used too low a microscopic power in his experiments

really do not hold good : instead of weakening they only make his case so much the stronger. The change which takes place in a liquid when it becomes charged with living organisms is perfectly apparent to the naked eye, as well as evident to the nose in most cases. If Pasteur detected so many organisms floating in the air, with a magnifying power of 350 diameters, what must have been the number brought under his notice had he used a glass magnifying 1,000 diameters? Dr. Bastian used an immersion glass magnifying 600 times (linear), and the particles of matter seen in his liquids were in many cases so trifling that they could very easily have come from the air. Very strong evidence that in some cases this was their origin, exists in the fact that his method of opening his flasks was identical with that employed by Pasteur for infecting his liquids with the germs of different localities, viz., to let air rush into the vacuum.

CHAPTER V.

Examination of Arguments raised in Opposition to Pasteur's Conclusions—On Saline Solutions and the Formation of Crystals—The Effect of Atmospheric Dust on Crystallisation—Crystals and Organisms—The Antiquated Doctrine of Catalytic Change—On Arctic Dust—Researches on Atmospheric Dust by Dr. Angus Smith and Dr. Cunningham—On the Cause of Hay-fever—On the Ageing of Roquefort Cheese—Nourishment of Plant and Animal Life—Relation of the Minute Forms of Life to Putrefaction and Decay—Conclusion.

ALTHOUGH it is obvious that the arguments deduced by Dr. Bastian from his investigations are not deserving of any full discussion, seeing how a more careful repetition of his experiments has failed to afford that corroboration which is absolutely necessary before reliance can be placed on them, in short, before they can be quoted *as evidence*, yet it may be worth while to examine some of his objections to Pasteur's conclusions and some of the arguments he considers favourable to or explanatory of the doctrine of evolution of living matter *de novo*.

There is one striking characteristic in his method of reasoning : his premisses are all *possibilities*, while

he exacts from others not merely *probabilities* but *absolutely proved facts*.

Thus on pp. 93-95, 'Origin of the Lowest Organisms,' he says, in speaking of the preservation of liquids by means of cotton wool, 'The plug of cotton wool or the narrow and bent tube may, it is true, protect the boiled fluid from subsequent contact with living germs; but that the liquids do not undergo change on account of such deprivation cannot be safely affirmed, when the same means would also filter off from the fluid some of the multitudinous particles of organic matter¹ (dead) which the air undoubtedly contains, and which may act as ferments.' Which is the more probable, that living or dead matter should act as a ferment? We know that living matter does, while we know of no dead matter in the air that can do so.

The paragraph above quoted takes it for granted that there are no living germs in the air, ignoring the fact altogether that not only Pasteur proved this, but every one who has investigated the subject, including M. Pouchet, has confirmed this observation. In short, taking the statements of Bastian and Pouchet together—that there are living germs in the air, and that fermentations take place by means of living organisms—the former wishes us to believe that it is not the living parti-

¹ 'Origin of the Lowest Organisms,' p. 21.

cles which give rise to fermentation, but the dead, the cotton fibres, the starch grains, bits of wool, hair, grit, and soot. This is contrary to common sense. If the dead particles, let us say, for the sake of argument, initiate a fermentation, they would do this after the living ones had been killed by boiling water; consequently the liquids in Schwann's apparatus would not be preserved. On p. 98 he seeks refuge from the action of heat in assuming merely that 'molecular mobility,' whatever that may be, of these 'dead nitrogenous particles is impaired by the agency of heat.' Which is, again, the most reasonable, that living or dead matter should thus be changed? There comes on p. 98 this statement: 'Portions of organic matter can always be demonstrated amongst such atmospheric dust; whilst living bacteria, or other organisms, such as are first produced by the supposed sowing of spores, either cannot be demonstrated, or would soon prove other evidence to be very sparingly distributed.' Compare this with Pasteur's statement concerning atmospheric dust, also that of Dr. Cunningham, Mr. Blackley, and Dr. Angus Smith, which has yet to be referred to, and then its fallacy is evident. The ambiguous expression 'organic matter' can only be translated as germs.

Pasteur says of the *bacterium termo*, 'This

organism is so small that its germ must be indistinguishable, and if its germ were known it would not be possible to recognise it amongst the very various and numerous particles of organic dust collected from the air.¹ The late researches of Sanderson and Ferrier prove that bacteria germs are invisible under the highest powers of the microscope.

But to return to the examination of Dr. Bastian's arguments; there occurs the announcement on p. 101, 'Origin of the Lowest Organisms,' 'Some of the very fluids which remain pure in the bent neck apparatus will become fetid if shut up in vacuo.' This is only Dr. Bastian's assertion deduced from his own experiments. My experiments completely disprove this, as I made use of the same liquids as he did. Again, if we grant it, how can the nitrogenous particles in the atmosphere affect a liquid in vacuo? If some liquids, and they include, according to the experiments already quoted, infusion of turnip, and ammonium tartrate with di-sodic phosphate solution, are capable of yielding life without the action of dead nitrogenous organic particles, it is evident that when they are kept at first in vacuo, and subject to only filtered air, no objection can be raised against them whatever, yet they never showed signs of living things until exposed to the open air.

¹ 'Ann. Chim. Phys.' lxiv. p. 57.

With regard to any supposed analogy between the origin of organisms and the origin of crystals, a moment's consideration is sufficient to show how entirely it is without foundation (p. 108), 'living matter develops into organisms of different shapes, whilst crystalline matter grows into crystals of diverse shapes.' This sentence cannot be said to throw any light on the origin of life, for except in the fact that a difference of form results in both cases, there is not the slightest analogy between the cases. When an organism is produced in a liquid there is first of all a new kind of matter or several new kinds of matter formed by the splitting asunder of the compounds already existing, and a re-combination in a different way and in a more complete manner. These newly formed compounds are then moulded and distributed in a certain form, and this form has the power of performing definite functions; it can decompose ready formed compounds, and re-combine their elements in a different way so as to add to its own substance by the processes of growth and nutrition. The development of an organism in a liquid involves very complicated analyses and syntheses. In the crystallisation of a solution nothing similar occurs. We have a solution of a definite substance; this solution is in unstable equilibrium between the solid and the liquid states of aggregation; by a slight disturbing cause of a

physical nature, very often such as change of temperature or vibration, the balance is upset and the substance becomes a solid. We know perfectly well that sulphate of copper will not crystallise with 5 molecules of water until its solution has been evaporated down so that it practically has the composition of the crystals *plus* a solution of the crystals in water. So again with nitrate of potassium. When the crystals are formed it is because the whole liquid has the composition of a solution of nitre *plus* nitre, which separates out; or with alum a solution on the point of crystallising consists of a solution of the salt *plus* the actual substance in a liquid condition. When the temperature is lowered the substance no longer retains its liquid state, and in assuming the solid it takes a perfectly definite form. There are no analytical and synthetical processes at work here; the only change is from the one state to the other, from the liquid to the solid. Something analogous occurs when a gas is dissolved in water: it is in a state of unstable liquefaction, and a slight disturbing cause makes it assume its gaseous condition.

The effect of solid particles on the crystallisation of solutions is strikingly shown by the following experiments derived from the results of Lœwel and Schröder and Dusch. A large flask containing a gallon of alum solution is uncovered, and

observation will show that the first trace of dust which touches the surface causes a quantity of octahedral alum crystals to shoot throughout the liquid till it becomes a solid mass. It usually happens that a trace of the salt either off the stopper or the neck of the flask causes this crystallisation. With another solution, which is sulphate of soda, the flask is tightly corked. On removing the cork the air rushes in, the surface of the liquid is slightly disturbed by the fall of a speck of dust, and from this one single point crystals spread throughout the whole mass.

That it is not the entrance of air, but the entrance of solid matter with the air which causes crystallisation, can easily be seen with another flask, which is simply plugged with cotton-wool, so that air obtains free access. On removing the wool from the neck of the flask crystallisation commences at some point on the surface of the liquid. That the solid matter is dust floating in the air may thus be shown: a flask of sulphate of soda is provided with two tubes, one descending to the bottom of the liquid, the other just passing through the cork; these tubes are stopped with cotton wool. Now on removing one of the plugs and drawing air filtered by the wool through the liquid no crystallisation takes place. If, again, the air be taken from the lowest depths of

the lungs and blown through the liquid without the intervention of the cotton wool there is still no change ; but the air from a pair of bellows, ordinary dusty air, at once causes the solidification to occur. That the dust causing crystallisation is destructible by heat is demonstrated in the following manner. A cylinder of a supersaturated solution is carefully uncovered and the end of a glass rod which has been heated, and afterwards cooled, is thrust in. No change takes place, but the rod is now turned round and the other end inserted, when the liquid crystallises at once. What is the nature of this dust in the air destructible by heat which induces crystallisation in this manner ?

From the experiments of Mr. Liversidge,¹ it seems more than probable, for it is the view favoured by the majority of investigators of this subject, that they consist of sulphate of soda. It is a salt which we know to be in the air, and solutions of sulphate of soda, according to all experience, are those which crystallise most readily on exposure. Moreover, if precautions be taken to remove the sulphate of soda from the air by means of baryta solutions, the remaining dust is powerless to cause crystallisation. Again, solutions of this salt, when uncovered in a very quiet room, the dust of which has settled for the most part, do not crystallise on

¹ 'Proc. Roy. Soc.' vol. xx. p. 497.

shaking, but immediately on bringing a little powdered sulphate of soda into the room crystallisation takes place.

We therefore see that the sentence (p. 108), 'It would appear that specks of living matter may be born in suitable fluids, just as specks of crystalline matter may arise in other fluids. Both processes are really alike inexplicable,' is incorrect; crystallisation is as fully an intelligible process as any physical phenomenon ever investigated. The reason why crystals are formed in some places and under certain conditions is a matter undecided, though why they are formed at all is as evident as why water boils, or steam becomes water.

In order to believe that the cases are parallel, the imagination must be stretched to admit the idea that the tissues of the fungi said to have been evolved *de novo*, from disodic phosphate, and ammonium tartrate solution, consisted entirely of all or any of the four substances existing in such a liquid. 'Both crystals and organisms in such cases under suitable conditions' 'appear at first as minutest visible specks in solutions which were previously homogeneous.' 'In the one case we have to do with crystallisable matter in solution, and in the other with those big-atomed, unstable compounds which constitute the so-called *colloidal*'¹

¹ With regard to the terms colloid and crystalloid, they are

state of matter.' How are we to reconcile this statement with the evolution of life from some of the first order of crystalline substances, such as the solution of ammonium and sodium salts? Surely Dr. Bastian would not have us believe in spite of all human knowledge that these compounds are unstable and colloid substances? Yet this would seem to be the case, since he mentions the fact of cyanate of ammonia being converted by boiling into urea, and then states:¹ 'This would seem to show that the passage from the crystalloid to the colloid mode of molecular collocation is by no means a difficult one, that it may be brought about, in fact, by very slight determining causes. As urea is well known to be one of the best crystallising substances in existence, this sentence is otherwise incomprehensible.

applied to solid substances soluble in water, which may be represented by gelatine on the one hand and common salt on the other. Colloids are not resolved into any definite crystalline form when assuming the solid state, whereas all the particles of a crystalloid substance are built on its own peculiar geometrical model, that of common salt being the cube, of alum the octohedron. A colloid can generally be resolved into a condition half-way between the solid and liquid states of matter, and is then commonly known as a jelly, but a crystalloid passes from the liquid to the solid suddenly and without any intermediate condition. A solution of a colloid is incapable of diffusion through a membranous material like bladder, while the reverse is the case with crystalloids, so that in this difference of properties we have a means of separating crystalloid from colloid substances.

¹ 'Nature,' vol. ii. p. 95.

No one can say that there is any further resemblance between the growth of organisms and the formation of crystals, than that their origin is caused by similar bodies, and that under certain conditions they increase and multiply ; a shallow and useless comparison. A substance which is soluble in water is deposited when the water is removed by evaporation : were organisms developed from solutions like crystals they would also be deposited from suitable liquids by evaporation, which is absurd. We must therefore have some statement, not only less vague but more definite, in fact a *definition* of what takes place, before we can listen to any such theory. As it stands now, too much is left to a very unscientific use of the imagination.

The shiftiness of the arguments advanced by the defenders of heterogenesis is very striking. For instance, when liquids are found to be preserved through being protected by means of cotton wool, it is stated that the cotton may filter off, besides the living spores, the dead organic particles which it is inferred produce the change. When it is asserted that no change takes place in hermetically sealed tubes, then it is by reason of the pressure of the confined gases that fermentative and putrefactive change is prevented. Next, on its being shown that fermentation and putrefaction go on under great pressures, it is because some fermentations

are not prevented by pressure. When, again, the objection of pressure is removed from the experiments, the disruptive agency of heat is resorted to as an explanation. Having got rid of this difficulty, the next refuge is that sufficiently strong solutions were not employed. This argument failing, another is advanced, that the liquid was not an infusion of hay, and therefore not of the first order of fermentable matter (no experiments distinctly bearing on any one of all these statements being put in as evidence); and then as a last and least cogent reason, it is said that too low a power has been used in the microscopic examination of the liquids. The weakness of such vacillating reasoning is obvious.

I was really the first to test the true value of Dr. Bastian's experiments. Since the publication of my investigation, however, other workers have come into the field, whose evidence, more particularly from a physiologist's point of view, is of greater value than my own, such as that of the late Dr. Pöde, of Oxford, and Prof. Ray Lankester, the well-known microscopist. Their mode of operating was in every respect exactly the same as that of Dr. Bastian, but they employed much higher magnifying powers ($\frac{1}{33}$ and $\frac{1}{30}$ -inch object glasses) and proceeded more carefully; for instance, in examining the liquids *before as well as after*, they were

sealed in tubes. It is gratifying to find that their evidence is concordant with mine in every respect. From the results of 53 experiments made with turnip-infusion mixed with cheese; and hay-infusion, they produce ample evidence to prove the inaccuracy of Bastian's observations. He always supposes that having boiled his liquids for 10 minutes, he has really deprived them of living things; but he never makes any attempt to prove this, and thus he begs the whole question. It is necessary that he should give unquestionable proof of his liquids being free from life when sealed in flasks, and such treatment as he subjects them to is, as Pasteur proved, totally inadequate for the purpose.

From the earnestness with which Dr. Bastian puts forth his conclusions, his experiments have received more attention than they would have gained on their own merits, for there is certainly nothing new, either in his mode of experimenting or in his results. Relying implicitly on their accuracy, he has revived an old theory and given it a new name. 'In a¹ not-living organisable fluid we have good reason to suppose that a living unit may originate; and this being so we should have in such a case a veritable instance of the passage of the not-living into the living. Life would here begin *de novo* owing to the occurrence of certain

¹ 'Beginnings of Life,' vol. i. p. 232.

new molecular combinations. To this process we propose to apply the name of *Archebiosis*.' In his communications to 'Nature,' and two works 'The Origin of the Lowest Organisms,' and 'The Beginnings of Life,' he makes unwarrantable assumptions and suggestions rather than distinct statements; in fact, he argues from premises which he has not first established. The evidence too which he quotes is not of the latest date, neither is it the product of exhaustive research, such as that of Pasteur.

He suggests or indicates the following arguments: 1st. That germ life, if it exists in the air at all, 'is only very sparingly distributed.'¹ 2nd. That the particles in the air which produce fermentative changes are not germs. 3rd. That those particles are fragments of dead nitrogenous unstable organic matter. 4th. That all living germs can be destroyed in liquids by boiling at 212° for ten minutes. 5th. That dead nitrogenous unstable organic particles in contact with suitable organic solutions can exert such an amount of chemical action as to decompose the constituents of the solution, recombine them in a different way, and out of such newly formed material build living organisms. Not a single point has he proved. The

¹ Compare this with the account of Dr. Cunningham, Dr. Blackley, and Dr. Angus Smith's work.

burden of proof lies with him ; those small ovoid particles found floating by myriads in the air which bear so close a resemblance to the spores of the minute fungi that the most able microscopists can even assign names to them, must be shown by him to be lifeless. They have, on the contrary, been shown to be living, because when heated in water above 212° F. they ceased to have the power of reproduction in those solutions the fertility of which was not affected by this temperature. The very theory on which the supposed dead particles in an assumed unstable condition are dependent for their power of communicating change to other matters, is the useful but now long antiquated fiction of *catalysis*. This theory originated with Berzelius, and was adopted by Liebig to explain the processes of fermentation and decay ; and concerning it Graham wrote thirty-two years since.¹ 'It would be unphilosophical to rest satisfied by referring such phenomena to a force of the existence of which we have no evidence. The doctrine of catalysis must be viewed in the light of a convenient fiction, by which we are enabled to class together a number of decompositions not provided for in the theory of chemical affinity as at present understood, but which it is to be expected will receive their explanations from new investigations.'

¹ 'Elements of Chemistry,' p. 197.

If on the one hand we are asked to believe the arguments of Dr. Bastian, we cannot admit his experiments, and on the other if we are asked to believe his experiments, we cannot assent to his arguments; but as his reasoning and his experiments should both be in harmony to make out a case, we see that he has failed completely in doing this. Again, if we compare both his reasoning and his experiments with those of M. Pasteur, Prof. Lister, and Mr. Ray Lankester later, we can admit neither and so find that he has doubly failed.

Dr. Bastian apparently believes in a hypothetical substance producing hypothetical changes by means of a hypothetical force. His theory of archebiosis, resting as it does on the fictitious catalytic action of supposed 'dead nitrogenous organic particles,' on so-called 'big-atomed unstable molecules' passing from the crystalloid to the colloid mode of molecular collocation,' is suggestive of the house which a man built upon the sand.

Further Researches on Atmospheric Dust.

Some idea of the varied nature of the dust floating around in the air, and which we are frequently conscious of swallowing in unpleasant quantities, may be given by just referring to the microscopic observations of several independent workers who

have given attention to this matter. Dr. Angus Smith has shown that in this country, many miles from the sea, crystals of sea-salt are found in the rain ; they are, however, most abundant near the shore. Many years ago it was noticed that after a strong westerly gale the windows of Drayton Manor, near Tamworth, the residence of Sir Robert Peel, were covered with minute crystals of common salt, no doubt deposited by the evaporation of salt-water spray carried high into the air and across the country from the Welsh coast. The air of Manchester contains a considerable amount of sulphate of soda ; so, indeed, does that of all towns where much coal is burnt. It is found as a deposit, together with particles of iron rust, when rain water is evaporated.

Sulphate of soda, however, abounds most in the neighbourhood of chemical works or those towns which form the centres of chemical industry. When coal is burnt the sulphur is discharged into the air in a gaseous form, which ultimately becomes converted into sulphuric acid ; therefore the rain of all large towns is found to be acid, and it has a slow but distinctly destructive action upon such building stones as consist of magnesian limestone. Now if there be in the air traces of common salt, this in presence of the sulphuric acid would be converted into sulphate of soda ; hence the salt in

country air would be converted into the sulphate of soda when it arrived at a large town. The air of mines, which has a peculiarly irritating action on the respiratory organs, is found to hold in suspension minute crystals of saltpetre and also probably of sulphate of potash. These crystals are derived from the gunpowder used in blasting, and at a time when the 'shots' or charges are fired it is blown into the air with the smoke. Blasting with gun-cotton has been found advantageous in some cases on this account, but the use of nitro-glycerine, on the other hand, has been the cause of severe headache, from a small quantity of the vapour of the substance being driven into the air, the inhalation of which produces this effect on the workmen. Dr. Angus Smith has examined the glittering dust seen in the sunbeams during rapid railway travelling, and found it to consist of minute rolled plates of iron detached by friction from the rails and wheels. During a railway journey near Birmingham, Mr. Sidebotham remarked the gritty nature of the dust in the carriage, and securing a portion of it ascertained that 50 per cent. by weight was composed of minute fragments of iron, some particles being evidently the result of wear and tear on the rails, while others were undoubtedly either the produce of neighbouring works or from the furnace bars of the engine, as they were of the

nature of burnt iron and clinkers ; there were also many small angular particles like cast-iron, having a crystalline structure. Other portions of the dust consisted largely of cinder, bits of coals, fragments of yellow metal, opaque white and spherical bodies, and bits of glass. It would be difficult to imagine anything in the way of inorganic dust more irritating to the eyes, nose, and air passages. Mr. Charles Stodder, of Boston, U.S., examined the dust from a beam in the United States armoury at Springfield, and found it consisted largely of minute fragments of iron. To prevent such dust circulating in the air, he recommended the use of magnets fixed close to the grind-stones and polishing wheels in the workshops, a plan which was put in practice, and abandoned some years previously in this country.

An interesting communication was made by M. Nordenskjöld to the French Academy of Sciences in the summer of 1873. In December 1871 he observed that snow, collected at the end of a five or six days' continuous fall, was mingled with a large quantity of sooty-like dust consisting of an organic substance rich in carbon. It bore some resemblance to the meteoric dust which fell together with meteorites at Hessel near Upsal in the beginning of the year 1869. Suspecting the railways and houses of Stockholm of polluting the

snow, he got his brother, who was living in a desert district in Finland, to give his attention to the matter, and the result was that he collected a similar powder. The snow gathered from floating ice in the Arctic regions and on the glaciers of Greenland, leaves, on melting, a greyish residue, consisting partly of the minute organisms known as diatomaceæ, but there are also particles of a considerable size containing iron, cobalt, nickel, silicon, carbon, and phosphorus.

The source of this dust is doubtful ; two of its constituents, cobalt and nickel, are of rare occurrence in terrestrial matter, while on the other hand they are very commonly found in meteorites. It may be suggested that the dust may either be blown up from sandy plains situated at some distance from the place where it is deposited, or else discharged from the mouth of some far-off volcano.

Further Microscopic Examinations of Air disclosing the presence of Germs.

To examine ordinary air according to Dr. Angus Smith's plan, a wide-mouthed bottle, which may vary in size from four to sixty ounces, perfectly clean and dry, is filled with air of the desired place by means of the concertina bellows before mentioned, and into this is put about $\frac{1}{4}$ ounce of

the purest possible distilled water, a commodity difficult to obtain free from floating particles. Very pure block ice perhaps yields it most easily. Without following the somewhat elaborate but undoubtedly successful plan of Mr. Cottrell, devised at the Royal Institution, one may take some carefully washed lumps of ice in a bottle and allow them to melt, pouring away the first two or three ounces of water after rinsing the glass. The bottle full of air to be examined is then shaken well, not to say violently, for a few minutes; another quantity of air may be then introduced without a fresh addition of water, and the operation repeated as often as may be desirable. The liquid will soon become turbid, and the number of times the air has to be renewed to give visible turbidity varies very much with different localities. Mr. Stodder has raised objections to this plan of collecting dust from the air, but, as will be seen, it has yielded useful results.

In an experiment made by Dr. Smith with a bottle holding 5 litres, which was refilled 500 times with Manchester air, remarkable results were yielded by Mr. J. B. Dancer's examinations. The magnifying powers employed varied from 120 to 1,600 diameters. The bodies seen were—

1st. Particles of vegetable tissue, many of them partially burnt and quite brown in colour, exhibit-

ing the pitted structure of fir-wood and other *Coniferae*, probably wood used in lighting fires.

2nd. Fragments of vegetation resembling in structure hay, straw, and hay seeds.

3rd. Hairs of plants, and fibres resembling flax.

4th. Cotton-fibres, both white and coloured.

5th. Starch granules.

6th. Wool, white and coloured.

7th. In greatest abundance fungoid matter, spores and sporidia, varying in size from $\frac{1}{10,000}$ to $\frac{1}{50,000}$ of an inch in diameter.

Many of the spores were living and developed forms resembling rust or mildew. A calculation was made as to their number in the following manner.

Thus under each field of the microscope there were more than 100 spores. In each drop of liquid there were over 250,000; the whole quantity consisting of 150 drops; there were then in this water *no fewer than 37½ millions of spores visible to a magnifying power of 1,600 diameters.* This quantity of air is the amount respired by an average-sized man actively employed during a period of 10 hours in Manchester.¹

The researches² of Dr. Cunningham are of

¹ 'Air and Rain, the Beginnings of a Chemical Climatology.' Dr. R. Angus Smith, F.R.S.

² 'Microscopic Examinations of Air.' D. Douglas Cunningham, M.B., Surgeon to H.M. Indian Medical Service. Published by Government, 1874.

equal interest and the most extensive on the subject yet published. They were made at Calcutta for sanitary purposes. By means of a vane attached to a tube a stream of air passed over a microscope slide moistened with glycerine, and the solid contents of the air were retained by this viscid liquid. The arrangement is really a weather-cock, which always presents the glass slide to the wind. The microscopic examination was made with powers varying from 400 to 1,000 diameters. At a height of about 5 ft. from the ground the following objects were collected :

1. Particles of siliceous matter.
2. Particles of carbonaceous matter.
3. Hair and animal substances.
4. Fragments of the cellular tissues of plants.
5. Pollen grain, amongst which were those of many common grasses, besides some of plants belonging to other natural orders.
6. Algæ. In comparison few in number. There were, however, undoubted fragments of *Oscillatoria*, *Desmidiaceæ* *Closterium*, and apparently *Diatomaceæ*, though these were very few.
7. Sporidiæ of lichens frequent.
8. Far the greater part of the bodies are sporidiæ of fungi, often referable at once to their proper genera.¹ Spores of *Macrosporium* and one

¹ See Rev. M. J. Berkeley's article in 'Quarterly Journal of Microscopical Science,' April 1874.

or two other allied genera are extremely common. *Cladiosporium herbarum*, one of the most universally diffused fungi, appears in one case with a spore *in situ*. *Helminthosporium* is represented; *Sporidesmium* is not unfrequent. The yeast fungus, a particular condition of common species of *Penicillium*, *Aspergillus*, and *Mucor* frequently occur, either in scattered particles or branched. A young *Mucor* with sporangia is seen amongst the other objects drawn. Spores of *Uredineæ Puccinia* and *Sphaeriaceæ*, with many others.

‘The existence of *bacteria* in half of the specimens is also very worthy of consideration when the extreme rarity of such organisms in a recognisable form as a constituent of common atmospheric dust is recollected.’ The author concludes his work with the following remark:—‘No connection can be traced between the numbers of *bacteria* spores, &c., present in the air and the occurrence of diarrhœa, dysentery, cholera, ague, or dengue, nor between the presence or abundance of any special form or forms of cells and the prevalence of any of these diseases. Not less interesting are the researches of Mr. Blackley on *Catarrhus æstivus*, or hay fever. They were made in the summer of 1873 during the months including April and July. Slips of glass covered with a liquid not likely to dry were exposed horizontally a height of about 4 ft. 9 inches from the ground, and these were found

to become covered with large quantities of pollen grains. The greatest number obtained on a square centimètre of surface ($\frac{3}{8}$ -inch square) in 24 hours, was 880, on June 28th. Sudden diminutions in quantity were observed occasionally, and these were due to either a fall in the temperature or a shower of rain, or both causes combined. A most ingenious method was employed to obtain some knowledge as to the height at which this pollen dust could be detected floating in the air. The glass slips were attached to kites, and by fixing the string of one kite to the body of another an elevation of 1,000 ft. was attained. By this means the discovery was made that at the higher strata of the air pollen was much more largely present than at the 'breathing level,' in so large a proportion as 19 to 1. Abundant proof was obtained of the presence of fungoid spores in large quantities of air. On one occasion the spores of a cryptogam at the height of 1,000 ft. were found to be so numerous that they could not be counted, but they were estimated to be not fewer in number than 40,000 to the square inch. That these minute organised particles travel through the air to great distances was proved by a series of experiments made near Manchester, within the boundary of one of the most populous parts and with no grass land within a third of a mile. The quantity of pollen in this

locality was about $\frac{1}{10}$ that collected in the country. Mr. Blackley considers he has proved that hay-fever is caused by the inhalation of air containing pollen in considerable quantity, and that the pollen adhering to the membranous lining of the larynx and air passages and nasal lachrymal membrane causes irritation, and the excessive secretion from these parts. A solution of quinine, which is destructive to minute forms of life, has been shown by Helmholtz to be an effective application in cases of this disagreeable malady. The liquid is syringed into the nostrils. The Rev. W. H. Dallinger has shown that even the microscopical powers employed by Dr. Cunningham are insufficient, that it is only by employing object glasses of the highest powers, such as the $\frac{1}{28}$ and $\frac{1}{30}$, in conjunction with the most powerful eyepieces, that anything like the true number of organisms collected from the air can be seen. Dr. Lionel Beale, in his work on 'Disease Germs,' gives a drawing of *bacteria* germs magnified 5,000 diameters, and also minute fungus germs of much less size than the $\frac{1}{100,000}$ of an inch. If then these organised bodies are so extremely minute, how much more emphatic would Pasteur's evidence have been made had he been able in 1859 to have employed such enormous magnifying power; how many more living corpuscles would he not have discovered! It is a matter almost of

surprise that he did so much with a microscope magnifying only 350 diameters.

In the 'Monthly Microscopical Journal' for August 1, 1873, there is an account of experiments by W. H. Dallinger and J. Drysdale, M.D., entitled 'Researches on the Life History of a Cercomonad; a Lesson in Biogenesis.' It is a most instructive and suggestive paper, containing in a very short space the results of a great deal of very laborious research. They remark that 'the appearance or non-appearance of organic forms in certain infusions placed in sealed flasks or tubes, or otherwise conditioned, is held to be decisive of their production *de novo* or otherwise, but in point of fact we know *nothing*—absolutely nothing—of the life history of the greater number of forms produced.' 'To attempt to decide therefore from the experiments as yet published, that their production in gross masses in inorganic infusions proves that inorganic elements produced them, may be to beg the whole question.' They speak of the advantage secured by the combined investigation of two observers, and allude to the fact that objects have been under constant observation for nine consecutive hours, while the same drop of liquid has been watched for three weeks. A peculiar flagellated monad¹ noticed in the water in which the head of

¹ A somewhat oval-shaped organism to which is attached a long whip-like appendage.

a cod-fish was macerated, was seen to multiply by fission for a period extending over from two to eight days. It then became amœboid ; two individuals coalesce, slowly increase in size, and become a tightly distended cyst. This cyst bursts, and an innumerable host of immeasurably small sporules are poured out as if in a viscid fluid and densely packed, barely recognisable with a $\frac{1}{80}$ -inch objective magnifying 2,500 diameters. These sporules are scattered, they slowly enlarge, acquire a sort of motion like the fly-wheel of a watch, flagella are next formed, and they become active, attain rapidly the parent form, and once more multiply by fission. The average of forty observations on this fission shows that it occupied 4 minutes 40 seconds for completion. It was found that when adult forms and sporules together in a drop of liquid were evaporated to dryness and heated to 250° F., on moistening again all the adult organisms were found dead ; but in two out of seven instances gelatinous points were seen like the early stage of developing sporules, and these after some hours' watching attained the flagellate state. Although in water at a temperature of 150° F. the adult monads are killed, *young monads appear and develope* in an infusion which has been raised to 260° F. ; suggesting that the sporule is uninjured at a temperature much higher than that wholly destructive to the adult.

Cohn confirms the fact that Bacteria are not always killed by *boiling* in flasks, and that *Bacillus subtilis*, the lactic ferment, survives the boiling of the solution in which it is contained, and that in every case the boiling should be continued an hour. This establishes the accuracy of Pasteur's observations on the fermentation of milk.

Although we generally associate the growth of minute vegetation with the destruction of food, as, for instance, when bread or fruit becomes mouldy, and milk curdled and sour, yet in some cases the growth of the microscopic fungi effecting such changes converts plain viands into delicacies. For instance, the mellowing of a Stilton cheese is due to the spreading of a fungus, the *Aspergillus glaucus*, throughout the mass. The chemical change produced in the ageing of Roquefort cheese, which is owing to a like cause, has been made the subject of a careful investigation by a French chemist, M. Blondeau. The Roquefort cellars of Messrs. Rigal are teeming with the peculiar minute cryptogamic vegetation which we commonly call mould; and into these places are the new cheeses brought to remain a year. At first they consist of pure caseine or curd free from fat, odourless and tasteless. After a year's time the cheese acquires a greyish-brown colour, a strong and peculiar odour, and piquant flavour. Subjoined

in a tabular form are shown the constituent substances of the old and new cheese, the results of a great many analyses :—

Composition of New Cheese and Old Cheese kept 1 year.

	Per cent.		Per cent.
Caseine . . .	85.43	Caseine . . .	40.23
Fatty matter . . .	1.85	Margarine . . .	16.85
Lactic acid88	Oleine . . .	1.48
Water . . .	11.84	Butyrate of ammonia . . .	5.62
	100.00	Caproate of ammonia . . .	7.31
		Caprylate of ammonia . . .	4.18
		Caprate of ammonia . . .	4.21
		Common salt . . .	4.45
		Water . . .	15.16
			99.49

Salt not estimated.

One-half the caseine by keeping becomes converted into other substances, many of these being compounds of ammonia with what are called fatty acids. These acids can be obtained by the oxidation of oleine, and are, in fact, produced when butter becomes rancid. The flavour of rancid butter differs from that of Roquefort cheese, because the acids in the latter substance are neutralised by ammonia. Caseine is a highly nitrogenous substance: the nitrogen it contains gets converted into ammonia or some compound thereof, while portions of its carbon and hydrogen are combined in such a way as to form oleine. This oleine undergoes oxidation, producing the fatty acids which, instead

of being in the free state as in rancid butter, are neutralised by the formation of ammonia salts.

It is more particularly to caprate of ammonia that Roquefort cheese owes its peculiar sharp flavour.

As was first shown by Cagniard de Latour and Schwann, the growth of a fungus at a prodigious rate, and the alterations it causes in the chemical composition of saccharine fluids, constitutes fermentation, the most essential process in all breweries and distilleries. In this case the yeast plant, a particular fungus of the torula form called *Saccharomyces cerevisiæ*, *Mycoderma cerevisiæ*, also *Hormiscium cerevisiæ*, is put into the liquid to be fermented. Its chemical action on the sugar is so extraordinary, and its own growth at the expense of the nitrogenous matter and phosphates in the wort so rapid, that were it not for microscopical research it would be impossible to believe the changes due to vital processes. The change consists in the conversion of sugar into alcohol and carbonic acid: there are other products, such as glycerine and succinic acid, but they occur only in small quantities. When a saccharine fluid containing nitrogenous matter and phosphates is exposed to the air it is always found to undergo a similar change.

The growth of other fungi often takes place in liquids which contain phosphates and ammonia salts or nitrogenous organic matter. Being exces-

sively minute, these organisms require only exceedingly small quantities of matter for their nutrition, and as a large mass of evidence points to the conclusion that a most deadly class of diseases—for instance, most notably cholera—is due to the action of such organisms, it is of the highest importance that our drinking water should be so free from nitrogenous organic matter as to be incapable of affording sufficient nutriment to germ-life. Practical experience has shown that wherever water used for domestic purposes is contaminated even to a very small extent by organic matter of a nitrogenous nature (which is generally found to be derived from sewage), outbreaks of typhoid fever are common. Cholera has been distinctly traced to drinking water so contaminated, and the sewage spreading the mischief to houses where the disease had been raging. It seems, however, that typhoid fever, which is always present more or less in this country, is propagated by germs present in all kinds of sewage matter, while the cholera germs are imported from the East and developed and transmitted in water to which they gain access. One of the most important applications of science to sanitary purposes is the chemical analysis of drinking-water, not that we can identify those special kinds of organic pollution in the water which are injurious, but we gain information as to whether it contains sufficient food

for the nutrition of minute forms of living and possibly hurtful matter.

The Destruction of Organic Substances.

Broadly expressed, all nature consists of three divisions, the mineral, vegetable, and animal. Mineral matter in one form or another is the food and substance of plants. On the other hand, animals would starve on mineral food ; the body is unable to convert it into substances which would replenish wasted tissues, neither has chemistry so far advanced that such a process could be devised. But animal life, in consuming vegetation while accomplishing its own nutrition, helps to return to earth and air, by the process of excretion, those simpler substances which form the sustenance of plants. From recently expressed views regarding the origin of life, the first living cell must have belonged to either a bacterium or one of the fungi, for it is from the supposed evolution of these forms of life from lifeless materials that deductions have been drawn. But neither bacteria nor fungi can exist on matter which, setting aside all life operations, is derived from the earth's crust and atmosphere, neither has it been attempted to be shown that they can be evolved from such materials. In truth, bacteria and fungi occupy a

remarkable position, for while they may augment their substance by deriving nitrogen from ammonia salts or the ammonia of the air, carbon must be presented in an unoxidised form, or perhaps, to speak more correctly, in an oxidisable form.

Animals and even plants themselves are by a process of decay made into food for plant life. Such a process as putrefaction or decay cannot be without a cause, and such cause we cannot trace to the very sluggish action of air or water at ordinary temperatures. It is easy to obtain evidence exactly corroborating the views of Pasteur published in 1863:¹ 'Fermentation, putrefaction, and slow combustion are the three natural phenomena which concur in the grand operation of the destruction of organised matter, a necessary condition of the perpetuity of life on the earth's surface.'

'Above all things, life, manifested in the production of the lowest organisms, appears to me one of the essential conditions of these phenomena, but life of a kind unknown until now, that is to say, independent of air or free oxygen.'

'Dead matter which ferments or putrefies does not yield, alone at least, to forces of a nature purely physical or chemical.'

With regard to the statement concerning organisms capable of living without oxygen, in a

¹ 'Comptes Rendus,' lvi. p. 734.

previous communication a particular fermentation of tartrate of lime mixed with water had been described, and an experiment of this kind was made. Tartrate of lime boiled with water to get rid of the dissolved oxygen and kill the germs was protected from the air by a layer of oil ; a minute quantity of the organised ferment, a species of vibrio, was then introduced under the oil, and this multiplied in the deposit of the lime salt. Similar facts were noticed in butyric fermentation and in the conversion of lactate into butyrate of lime. Another method of making the experiment consisted in filling a flask, having a curved tube-neck, with water and the tartrate ; the liquid was not boiled, but the tube-neck dipped into mercury to exclude air. *Monas* and *Bacterium termo* were developed in the liquid by means of their germs, and the oxygen dissolved in the water. In twenty-four to thirty-six hours all the oxygen was consumed to the last trace, and replaced by a somewhat smaller volume of carbonic acid : only then did the ferment, a species of vibrio, probably *Bacillus subtilis*, make its appearance. Strange to say, oxygen is not only unnecessary, but is poison to these living ferments. On passing air through the liquids containing them they sink to the bottom and the fermentation ceases. Carbonic acid, on the other hand, does not affect them. With regard to the part which oxygen plays in the de-

struction of organised matter, Pasteur discovered, by an analysis of the air in certain flasks (already alluded to) containing liquids, which had been opened and resealed, that only a trifling absorption of oxygen had taken place, notwithstanding the duration of these experiments was three years, for eighteen months of which time the temperature was maintained at 86° F. The exact quantity of oxygen absorbed was found in the carbonic acid produced, allowing for the co-efficients of solubility of the two gases in the liquids experimented on. In the case of milk the oxidation of the fat had caused the oxygen to disappear and be replaced by carbonic acid ; it is usually so with other oily matters. The liquids were otherwise unchanged, being perfectly transparent as regards the sugar solution and urine, while the milk was uncurdled and possessed of its naturally alkaline reaction. Sawdust from oak-wood boiled with a little water absorbed from pure air only a few cubic centimètres of oxygen in a month, while the same quantity not heated and exposed to ordinary air absorbed nearly 140 cubic centimètres of oxygen in fourteen days, and was found to be coated with mycelium filaments and spores of *Mucedines*.

Perhaps of all matter the most putrescible is blood. With the assistance of M. Claude Bernard, Pasteur took specimens of arterial and venous blood

from a healthy dog, and kept it in flasks of pure air at a temperature of 86° F. It never showed any signs of putrefaction and retained its fresh smell. After preservation for a month or six weeks the air of the flasks was analysed and the loss of oxygen was only 2 or 3 per cent. In one case of a flask containing urine, the air was so little altered as to give on analysis the following percentage composition :—

Oxygen	19·2
Carbonic acid	0·8
Nitrogen	80·0
							<hr/> 100·0

This is a proof of the almost indestructible nature of organic substances of whatever kind when protected from the ravages of living matter.

Could any one doubt the accuracy of these observations and conclusions, it is only necessary to turn to the recent very important experiments of Drs. Sanderson and Ferrier,¹ 'on the origin and distribution of microzymes (bacteria) in water.' They are of the greatest interest and in addition to Pasteur's researches contribute greatly to our knowledge of what is constantly going on around us. No special method calling for description was em-

¹ 'Thirteenth Report of the Medical Officer of the Privy Council,' and 'Quarterly Journal of Microscopical Science,' vol. xi. p. 323.

ployed. To clean the vessels and tubes, after very careful washing they as usual in such cases heated to 390° F., or thereabouts. Certain of their sealed tubes containing liquids were heated to near that temperature, while, to protect liquids from the action of the air, plugs of cotton wool were used: the formation of microscopic fungi was thus prevented. In not a single instance did they find reason to suspect the evolution of organisms *de novo*, while they distinctly prove that bacteria arise from pre-existing germinal matter which has hitherto proved itself to be beyond the search of the microscope. This germinal matter is attached to all ordinary surfaces and is distributed in distilled water, being derived from contact with unclean vessels. Its presence was demonstrated in water apparently perfectly pure chemically, microscopically, and almost so to a beam of light, the scattered rays being of a blue colour.

Vessels washed and dried in the ordinary way with a towel contain this germinal matter. Now as it is impossible to admit of living matter being soluble in water, these living particles must be of excessive minuteness. Blood, urine, albumen, and Pasteur's sugar solution which never came in contact with any but previously heated vessels, when protected from atmospheric germs by plugs of cotton wool remained perfectly clear and limpid.

When, however, a drop of cold distilled water was placed in the liquids, the contagion charged them with bacteria. One very remarkable experiment was the cutting of a piece of muscle from a freshly slaughtered rabbit by means of a previously heated knife, it being hung up by glass hooks, also cleansed by heat, under a bell jar. No putrefaction took place: a growth of mould simply formed on the surface.

Regarding the action of bacteria on the liquids in which they live, the most important facts observed are: (1.) That their growth is attended with absorption of oxygen and discharge of carbonic acid. (2.) That they are remarkably independent of the chemical constitution of the medium, provided they are supplied with oxygen; and (3.) That they take nitrogen from almost any source which contains it, and use it for building up their own protoplasm.

‘It is this last power which specially indicates their place in nature as the universal destroyers of nitrogenous substances, acting as the pioneers if not the producers of putrefaction.’¹

According to Pasteur, both bacteria and the organisms he calls vibriones co-operate in the process of putrefaction.

¹ ‘Quarterly Journal Microscop. Soc.’ pp. 326–7. Cohn quotes and confirms these results.

When a substance putrefies, for instance, a liquid, it is first deprived of all its dissolved oxygen by the bacteria. They form a layer on the surface (zoogloea), the result being that no oxygen can penetrate to the liquid. The vibriones are then developed, being protected from the too direct action of the air. The putrescible liquid then becomes the seat of two kinds of very distinct chemical actions related to the physiological functions of two different varieties of organisms which it nourishes. The vibriones, on the one part, living without the aid of atmospheric oxygen in the interior of the liquid, determine processes of fermentation, that is to say, they transform nitrogenous matters into more simple but nevertheless complex substances. The bacteria, on the other part, burn up those same products and resolve them into the most simple binary compounds, water, ammonia, and carbonic acid. Such are the results of putrefaction effected by means of free contact with the air. On the contrary, in the case of putrefaction out of contact with the atmosphere, the products of the splitting up of the putrescible matter remain unaltered. This explains why putrefaction in contact with air is a phenomenon which, if not always more rapid, is at least more complete, more destructive of organic matter than putrefaction out of contact with air. For example, if calcium lactate be putrefied out of

contact with air, the vibrio ferment transforms the lactate into various products, amongst which calcium butyrate is always found. This new compound, undecomposable by the vibrio which has provoked its formation, will remain indefinitely in the liquid without any alteration. But repeat the operation in contact with air, and it will be found that in proportion as the vibrio ferment acts on the inner portion of the liquid, the pellicle on the surface consumes the butyrate gradually and completely. If the fermentation is very active, the combustion at the surface is arrested, but only because the carbonic acid disengaged prevents the access of atmospheric air. The phenomenon recommences as soon as the fermentation is completed or slackened. This is, indeed, why, if a natural saccharine liquid is fermented out of contact with air, the liquid is charged with alcohol, whilst if the operation takes place with access of air the alcohol becomes acetified, burnt up, and transformed entirely into water and carbonic acid; then the vibriones appear, and following them putrefaction, when the liquid contains no more than water and nitrogenous matter. Finally in their turn the vibriones and the products of putrefaction are burnt up by the bacteria, the last surviving of which cause the destruction of those preceding them and thus accomplish the complete return of organised matter

to the atmosphere and to the mineral kingdom. Cohn's researches here again corroborate Pasteur's.

The advantage of this action on organic matter being a life process, a function of living things, is, that inorganic matter is not oxidised; it is only that kind of material which affords carbonic acid, water, and ammonia which is liable to attack.

The fetid smell evolved during putrefaction is due to certain sulphur compounds, the offensiveness increasing with the proportion of sulphur contained in the putrefying matter. The smell is scarcely perceptible, if the substance contains no sulphur. Every chemist is acquainted with the disgusting nature of most organic sulphur compounds and of sulphuretted hydrogen, one of the most constant products of the decomposition of albuminous matters, and therefore the bad odour of sulphurised products of putrefaction is easily intelligible.

In the consideration of solid substances under putrefaction Pasteur showed, if all possibility of living germs gaining access to the interior of an animal were prevented, that after death putrefaction commenced on the surface. But if an animal be kept under usual conditions putrefaction commences in the intestinal canal, for there are found in this part, as Leuwenhoeck has noticed, fully developed vibriones. Being out of contact with air

and bathed in liquid, their condition is that most favourable to multiplication and performance of their functions, and the body which has been preserved by the life and nutrition of its organs will then succumb to their action.

The living body is, as regards moisture, temperature, and material at hand, a most fitting place for the development of life, from the multitudes of germs that cannot escape being taken in with the air breathed and the food consumed; but the natural functions interfere and the gastric juice, as shown by Severi,¹ has the power of destroying the agents of putrefaction and disease. As Dr. Lionel Beale has pointed out, 'so long as the higher living matter lives and grows, the vegetable germs are passive and dormant, but when changes occur and the normal condition departs, they become active and multiply.'

Such are mainly the views of Pasteur, and it must be admitted that they are the result of the most extensive and elaborate research allied to the most brilliant reasoning.

The liquid and solid constituents of meat react on each other to produce to some extent a sort of change which is entirely different from fermentation or putrefaction. If the meat be enveloped in a cloth steeped in alcohol and placed in a vessel so

¹ 'Zeitschrift für Chemie,' (2) iv. 285.

that the vapours do not escape, there will be no putrefaction, neither on the outside, because of the antiseptic power of the alcohol vapour, nor within, because vibriones are absent. If the quantity of flesh thus preserved be small, it acquires a gamey flavour, and if large it presents the appearance of gangrene. There is no resemblance either in nature or in origin between putrefaction and gangrene; indeed, gangrene may be likened to the ripening of fruit after gathering.

As a necessary step to an exposition of the changes that occur in fruit when ripening, it is necessary to return to the action of ferments. Ferments such as yeast are rendered less active or produce a smaller amount of alcohol in exact proportion to the quantity of free oxygen they absorb. Under ordinary circumstances common mould when growing in a solution of sugar absorbs oxygen from the air, feeds upon the organic matter in the solution, and discharges carbonic acid without the production of any alcohol. If, however, the mould be immersed in the liquid, and so deprived of atmospheric oxygen, carbonic acid is slowly liberated and alcohol produced, in other words, the fungus under these altered conditions acts as a ferment. In the former case, the free oxygen absorbed from the air enters into combination with the carbon of the organic matter, giving out

heat, and yielding carbonic acid ; it resembles, in fact, ordinary respiratory combustion. In the latter example, the life of the fungus continues to be deriving from the decomposition of the sugar the heat required for its existence. Alcohol and carbonic acid are produced precisely in proportion to the duration of the acts of nutrition of the plant under its changed existence. It is from the small amount of oxygen in the fermentable matter that the carbonic acid is derived. Evidently then under these latter conditions, in order to produce the same amount of heat, or of carbonic acid, a much greater quantity of saccharine matter must be decomposed, much larger, in fact, than the weight of the original organisms. These facts have the following bearing upon the ripening of fruit. When connection with the parent tree is severed by, for instance, gathering, the cells are still living, and a process similar to that of fermentation goes on in the interior, from the saccharine matter and the organic acids, alcohol and various ethers are produced, giving the distinctive flavour to the fruit and reducing its acidity. The exposure to the air leads to an absorption of the oxygen ; the alcohol, ethers, and sugar are changed to water and carbonic acid—this latter gas being equal in bulk to the oxygen absorbed—the fruit loses flavour, softens, and spoils, and is finally attacked by those organ-

isms whose life-processes are characteristic of putrefaction. If, however, as Bérard has shown, the fruit be immersed in carbonic acid, or some other inactive gas, and so protected from oxygen, a large amount of alcohol is produced. A very perceptible amount of carbonic acid is still formed but no softening of the fruit occurs. This is seen with bunches of grapes when taken out of a wine tub; their flavour differs entirely from that of freshly gathered fruit, because they are immersed in carbonic acid. Fresh grapes enclosed in a vessel of this gas acquire the same flavour. Pasteur made an experiment with plums taken from a tree when nearly ripe; twenty-four of these being placed in carbonic acid, yielded after a few days 6·5 grammes of alcohol, while the fruit had not undergone any change, but was quite sound. A corresponding quantity of sugar had of course disappeared. The remaining twenty-four plums left in contact with air had become soft, watery, and very sweet. Grapes, melons, and all fruits containing acids behave thus. The living cells of the plums containing fermentable matters, act upon these as ferments do. Something similar to what occurs with the ripening of fruit may reasonably be considered as taking place with newly-mown grass; a process of fermentation is commenced, developing a large amount of heat, and one of the products is

the sweet scent of hay. Surely no better conditions for ripening fruit could be devised than those which exist inside a hay-stack, where, protected from oxygen, it would be surrounded by carbonic acid and kept warm ; and here theory accords with practice, for it is an old custom to ripen hard green pears by burying them an arm's length in a hay-rick. By organised ferments are meant organisms which can directly assimilate oxygenated matters capable of supplying heat by their decomposition, such, for instance, as sugar. Viewed in this light, fermentation appears to be a peculiar case of a very general phenomenon ; and all living things might be regarded as ferments in certain conditions of their lives, because none exist in which the action of free oxygen could not be momentarily suspended. Now to generalise from this, let any living being be killed by asphyxia, by section of nerves, &c., or let any organ belonging to such a living being be deprived of connection with the neighbouring parts, the consequence is that since the chemical and physical operations of life cannot be instantly extinguished, they will continue, and if this happens with privation of oxygen, usually gaining access from within or without, then the being, the organ, or the cell will derive the heat required for its modified process of nutrition, or for change in its tissues, from the

immediately surrounding materials. From that moment it will decompose those materials, and the peculiar character of fermentation will be manifest if the quantity of heat developed corresponds to the decomposition of a weight of fermentable matter, perceptibly greater than the weight of materials usually set in action by the living being, by the organism, or by the cell. Pasteur's published researches have not yet been extended to the study of these points of view in animal organisms. It is probable that the phenomena may differ in certain particulars from those exhibited by vegetable cells. Doubtless further investigations by this great chemist will throw much light upon these obscure points so intimately connected with chemistry, physiology, and pathology, the phenomena of putrefaction and gangrene. However, as regards dead organic matter, whether of animal or vegetable origin, it would be well-nigh indestructible, if all those minute and apparently useless organisms which are the cause of putrefaction were themselves destroyed ;¹ 'life would be impossible, because the return to earth and air of all that had ceased to live would be suddenly suspended.'

¹ 'Comptes Rendus,' lvi. p. 738. Pasteur.

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